

**PATENT APPLICATION**

**FOR**

**UNITED STATES LETTERS PATENT**

**TITLE: CROWN ETHER DERIVATIVES**

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VLADIMIR MARTIN**

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# CROWN ETHER DERIVATIVES

## INTRODUCTION

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### **Field of the Invention**

The invention relates to derivatives of crown ether chelators, including chromophoric and fluorescent derivatives that are useful for chelating metal cations. Where the chelator is labeled with a fluorophore, it is an indicator useful for the detection, discrimination and  
10 quantification of metal cations. The chelators are optionally substituted one or more times with a chemically reactive group or a conjugated substance, such as a biological or nonbiological polymer, or a lipid.

### **Background of the invention**

15 Metal ions play an important role in biological systems. Cells utilize metal ions for a wide variety of functions, such as regulating enzyme activity, protein structure, cellular signaling, as catalysts, as templates for polymer formation and as regulatory elements for gene transcription. Metal ions can also have a deleterious effect when present in excess of bodily requirements or capacity to excrete. A large number of natural and synthetic materials are  
20 known to selectively or non-selectively bind to or chelate metal ions. Ion chelators are commonly used in solution for in vivo control of ionic concentrations and detoxification of excess metals, and as in vitro buffers. When bound to a fluorophore, ion chelators are typically used as optical indicators of ions and are useful in the analysis of cellular microenvironments or dynamic properties of proteins, membranes and nucleic acids.

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Such indicators are also useful for measuring ions in extracellular spaces; in vesicles; in vascular tissue of plants and animals; biological fluids such as blood and urine; in fermentation media; in environmental samples such as water, soil, waste water and seawater; and in chemical reactors. Optical indicators for ions are important for qualitative and quantitative  
30 determination of ions, particularly in living cells. Fluorescent indicators for metal cations also permit the continuous or intermittent optical determination of these ions in living cells, and in solutions containing the ions.

A variety of fluorescent indicators that are useful for the detection of biologically relevant  
35 soluble free metal ions (such as  $\text{Ca}^{2+}$ ,  $\text{Mg}^{+}$  and  $\text{Zn}^{2+}$ ) have been described that utilize oxygen-containing anionic or polyanionic chelators to bind to metal ions. In particular,

fluorescent indicators utilizing a polycarboxylate BAPTA chelator have been previously described (U.S. Patent No.: 4,603,209 to Tsien et al. (1986); U.S. Patent No. 5,049,673 to Tsien et al. (1991); U.S. Patent No. 4,849,362 to DeMarinis et al. (1989); U.S. Patent No. 5,453,517 to Kuhn et al. (1995); U.S. Patent No. 5,501,980 to Malekzadeh et al. (1996); U.S. Patent No. 5,459,276 to Kuhn et al. (1995); U.S. Patent No. 5,501,980 to Katerinopoulos et al. (1996); U.S. Patent No. 5,459,276 to Kuhn et al. (1995). Some fluorescent indicators selective for  $\text{Li}^+$ ,  $\text{Na}^+$  and  $\text{K}^+$  in aqueous or organic solution have also been described, based on the chemical modification of crown ethers (U.S. Patent No. 5,134,232; and U.S. Patent No. 5,405,975; Gromov et al, Russian Chemical Bulletin (1999) 48:6 p. 1190-1192; Lockhart et al, J.C.S. Perkin I (1977) p 202-204).

In general, a useful property for metal ion indicators is the ability to detect and/or quantify a selected metal ion in the presence of other metal ions. Discrimination of  $\text{Ca}^{2+}$ ,  $\text{Na}^+$  and  $\text{K}^+$  ions in the presence of other metal ions is particularly useful for certain biological or environmental samples. For most biological applications, it is essential that the indicators be effective in aqueous solutions. It is also useful that indicators for biological applications be relatively insensitive to pH changes over the physiological range (pH 6-8) and sensitive to ion concentrations in the physiological range (for sodium, a  $K_d$  of about 5 mM to about 20 mM). It is also beneficial if the indicator absorbs and emits light in the visible spectrum where biological materials have low intrinsic absorbance or fluorescence.

Also useful are chelators that possess a chemically reactive functional group, so that the chelating group can be attached to polymers for use in remote sensing of ions or enhancing the solubility or localization of the optical sensor. Many chelators bind to intracellular proteins, altering the chelator's metal binding properties. In addition, due to their relatively small size, they are readily sequestered non-selectively in intracellular vesicles, further limiting their effectiveness. One means of circumventing these problems is to attach the desired crown ether to a large, water-soluble polysaccharide, such as dextran or FICOL, by means of modification of the polysaccharide to allow covalent attachment of the indicator. Dextrans and FICOLs are especially suitable for this application, as they are low cost, optically transparent above about 250 nm and available in multiple ranges of molecular weights. Furthermore, polysaccharides and their conjugates are reasonably compatible with most biological materials and do not interact significantly with intracellular components. Although fluorescent polysaccharides have been previously described, as have indicator conjugates of dextrans, none possess the advantageous properties of the indicator conjugates of the current invention.

The crown ether chelators of the invention show significant ability to discriminate between metal ions under physiological conditions, particularly  $\text{Ca}^{2+}$ ,  $\text{Na}^+$  and  $\text{K}^+$  ions. This selectivity can be tailored by careful selection of crown ether substituents. The compounds of the invention are typically soluble in aqueous solutions.

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The compounds of the invention that act as indicators for target ions absorb and emit light in the visible spectrum and possess significant utility as a means of detecting and quantifying certain metal ion levels in living cells, biological fluids or aqueous solutions. Upon binding the target ion in the chelating moiety of the indicator, the optical properties of the attached fluorophore are generally affected in a detectable way, and this change is correlated with the presence of the ion according to a defined standard. Compounds having relatively long wavelength excitation and emission bands can be used with a variety of optical devices and require no specialized (quartz) optics, such as are required by indicators that are excited or that emit at shorter wavelengths. These indicators are suitable for use in fluorescence microscopy, flow cytometry, fluoroscopy, or any other application that currently utilize fluorescent metal ion indicators.

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## **SUMMARY OF THE INVENTION**

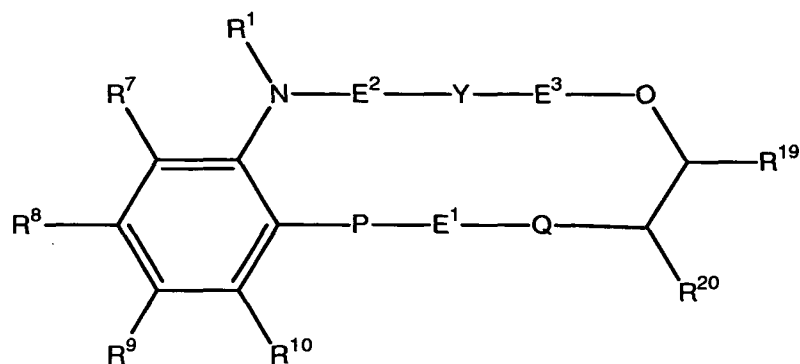
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The present invention provides metal chelating compounds that are derivatives of crown ether compounds that bind many metal cations including physiological relevant levels of metal cations such as sodium. These metal chelating compounds find utility in detecting, quantitating and monitoring cations such as  $\text{Na}^+$ ,  $\text{Li}^+$ ,  $\text{K}^+$ ,  $\text{Ca}^{2+}$ ,  $\text{Zn}^{2+}$  and  $\text{Rb}^+$ . A particular useful application is the binding of physiological levels of sodium ions in living cells wherein the compounds of the present invention provide for the detection, quantitation and monitoring of the intracellular sodium ions.

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The metal chelating compounds of the present invention are derivatives of crown ether compounds and have the following formula:



Formula (II)

wherein the compound contains at least one oxygen atom and preferably three to five oxygen atoms in the crown of the compound. The oxygen atoms are preferably separated by  $-(CH_2)_2-$ . Crown ether compounds that contain an oxygen and nitrogen atom ortho to the benzo moiety find particular use in binding sodium ions and in generating a detectable signal when bound by an  $-L-DYE$  moiety at one of the benzo substituents.

Thus, P and Q are independently O, S or  $NR^3$ , wherein each  $R^3$  is independently H or  $C_1-C_6$  alkyl. Typically P and Q are O.

More specifically, Y is O, S,  $NR^4$  or is absent.  $R^4$  is selected from the group consisting of H,  $-L-R_X$ ,  $-L-S_C$ ,  $-L-DYE$ ,  $C_1-C_{18}$  alkyl, aryl and heteroaryl ring system, which alkyl or ring system is optionally substituted by halogen, azido, nitro, nitroso, amino,  $C_1-C_6$  alkylamino,  $C_2-C_{12}$  dialkylamino, cyano,  $-L-R_X$ ,  $-L-S_C$ ,  $-L-DYE$ ,  $C_1-C_6$  alkyl or  $C_1-C_6$  alkoxy that is itself optionally substituted by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ .  $R^{15}$  is selected from the group consisting of H,  $C_1-C_6$  alkyl,  $-L-R_X$ ,  $-L-S_C$  and  $-L-DYE$  and  $R^{16}$  is selected from the group consisting of H,  $C_1-C_6$  alkyl, benzyl, a biologically compatible esterifying group, a biologically compatible salt,  $-L-R_X$ ,  $-L-S_C$  and  $-L-DYE$ .  $R^{17}$  and  $R^{18}$  are independently selected from the group consisting of H,  $C_1-C_6$  alkyl,  $C_1-C_6$  carboxyalkyl, alpha-acyloxyalkyl, trialkylsilyl, a biologically compatible salt,  $-L-R_X$ ,  $-L-S_C$  and  $-L-DYE$ ; or  $R^{17}$  and  $R^{18}$  taken in combination form a 5- or 6-membered aliphatic ring that optionally incorporates an oxygen atom;

$R_X$  is a reactive group that is capable of forming a covalent bond with another substance containing an appropriate reactive group to form a conjugated substance ( $S_C$ ). Particularly useful reactive groups of the metal chelating compounds include carboxylic acid and activated esters of carboxylic acid for labeling amines and alcohols of biomolecules. Conjugated substances are intended to mean any biomolecule or non-biomolecule that

contains a moiety capable of forming a covalent linkage with another moiety or is modified to contain such a reactive group. Particularly useful conjugated substances include proteins, peptides and non-biomolecule polymers. The covalent linkage (L) can be a single covalent bond or a series of stable bonds containing 1-20 non-hydrogen atoms including P, C, N, O and S.

$E^1$ ,  $E^2$ , and  $E^3$  are independently  $-(CR^5)_n-$ ,  $-(C(O)CH_2)_n-$ ,  $-(CR^5)_nO(CR^5)_n-$  or  $E^2$  is absent, where  $n = 2, 3$  or  $4$ , and each  $R^5$  is independently H or  $CH_3$ , or two  $R^5$  moieties on adjacent carbons of one or more of  $E^1$ ,  $E^2$  or  $E^3$ , when taken in combination, form a 5- or 6-membered aliphatic ring.

$R^1$  is selected from the group consisting of  $-L-R_X$ ,  $-L-S_C$ ,  $-L-DYE$ ,  $C_1-C_{18}$  alkyl and  $C_7-C_{18}$  arylalkyl, each of which is optionally substituted by halogen, azido, nitro, nitroso, amino, hydroxy, cyano, an aryl or heteroaryl ring system,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ ,  $-(C=O)-NR^{17}R^{18}$ ,  $C_1-C_6$  alkylamino,  $C_2-C_{12}$  dialkylamino,  $C_1-C_6$  alkyl or  $C_1-C_6$  alkoxy and each of which are optionally further substituted by halogen, amino ( $-NR^{17}R^{18}$ ), hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$  or  $-(C=O)-NR^{17}R^{18}$ .

A particularly preferred  $R^1$  is represented by methyl or ethyl that is substituted by  $-(C=O)-O-R^{16}$  or  $-(C=O)-R^{15}$  wherein  $R^{16}$  or  $R^{15}$  is a methyl group. These particular substituents appear to play a role in stabilizing the metal ion in the chelating ring of the present compounds.

$R^{19}$  and  $R^{20}$  are independently selected from the group consisting of H, halogen, azido, nitro, nitroso, amino, cyano,  $-L-R_X$ ,  $-L-S_C$ ,  $-L-DYE$ ,  $C_1-C_6$  alkyl and  $C_1-C_6$  alkoxy, each of which is itself optionally substituted by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ . Alternatively,  $R^{19}$  and  $R^{20}$  taken in combination form a fused six-membered benzo moiety that is optionally substituted by halogen, azido, nitro, nitroso, amino, cyano,  $-L-R_X$ ,  $-L-S_C$ ,  $-L-DYE$ ,  $C_1-C_6$  alkyl or  $C_1-C_6$  alkoxy, each of which is itself optionally substituted by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ .

$R^7$ ,  $R^8$ ,  $R^9$  and  $R^{10}$  are independently selected from the group consisting of H, halogen, azido, nitro, nitroso, amino, cyano,  $-L-R_X$ ,  $-L-S_C$ ,  $-L-DYE$ ,  $C_1-C_6$  alkyl or  $C_1-C_6$  alkoxy, each of which is itself optionally substituted by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ . Alternatively, any two adjacent substituents  $R^7-R^{10}$ , taken in combination, form a fused six-membered benzo moiety, which is optionally substituted by halogen, azido, nitro, nitroso, amino, cyano,  $-L-R_X$ ,  $-L-S_C$ ,  $-L-DYE$ ,  $C_1-C_6$  alkyl

or C<sub>1</sub>-C<sub>6</sub> alkoxy, each of which is optionally substituted by halogen, amino, hydroxy, -(C=O)-R<sup>15</sup>, -(C=O)-O-R<sup>16</sup>, or -(C=O)-NR<sup>17</sup>R<sup>18</sup>. In addition, any two adjacent substituents R<sup>7</sup>-R<sup>10</sup>, or R<sup>19</sup> and R<sup>20</sup>, taken in combination with each other, form a fused DYE wherein the dye shares the benzo moiety of the present compounds.

- 5 A particular useful compound of the present invention is when R<sup>8</sup> or R<sup>9</sup> is represented by a -L-DYE or R<sup>8</sup> and R<sup>9</sup> taken in combination form a fused DYE. In addition, when a lipophilic group such as an AM ester substitutes the DYE moiety the present compounds find use in binding *in vivo* metal cations. This is a particular useful aspect of the present invention wherein certain embodiments of the compounds, such as compounds containing only one  
10 nitrogen atom and one benzo moiety, enter cells and bind target ions with better results compared to similar compounds.

- The compounds of the present invention are useful for binding physiological relevant levels of cations, particularly sodium ions. When the present compounds comprise a DYE moiety the compounds find use to monitor, detect and quantitate such cations. Furthermore,  
15 compounds comprising a lipophilic group such as an AM or acetate ester group provide compounds that are cell permeable but are well retained in the cell after the lipophilic group is cleaved by nonspecific esterases resulting in a charged molecule.

## 20 BRIEF DESCRIPTION OF THE FIGURES

- Figure 1:** Shows the Na<sup>+</sup>-dependent fluorescence excitation spectra of Compound **51** in a series of solutions containing 0 to 1000 mM free Na<sup>+</sup>, with fluorescence emission monitored at 510 nm (as described in Example 71).

- 25 **Figure 2:** Shows the Na<sup>+</sup>-dependent fluorescence emission spectra of Compound **22** in a series of solutions containing 0 to 500 mM free Na<sup>+</sup>, with excitation at 488 nm (as described in Example 71).

- 30 **Figure 3:** Shows a comparison of the intracellular sodium response of SODIUM GREEN tetraacetate indicator (Molecular Probes, Inc.) or Compound **25**. At every intracellular Na<sup>+</sup> concentration, Compound **25** demonstrates stronger fluorescence intensity than SODIUM GREEN sodium indicator at the same concentration (as described in Example 72).

- 35 **Figure 4:** Shows the Na<sup>+</sup> (Fig. 4A) and K<sup>+</sup> (Fig. 4B)-dependent fluorescent emission spectra

of Compound **86** in a series of solutions containing 0 to 1000mM free Na<sup>+</sup> and 0 to 5M of free K<sup>+</sup> ions, with excitation at 488nm. See, Example 71.

**Figure 5:** Shows intracellular detection of Na<sup>+</sup> ions with Compound **87**. See, Example 72.

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## DETAILED DESCRIPTION OF THE INVENTION

### I. DEFINITIONS

10 Before describing the present invention in detail, it is to be understood that this invention is not limited to specific compositions or process steps, as such may vary. It must be noted that, as used in this specification and the appended claims, the singular form “a”, “an” and “the” include plural referents unless the context clearly dictates otherwise. Thus, for example, reference to “a fusion protein” includes a plurality of proteins and reference to “a  
15 fluorescent compound” includes a plurality of compounds and the like.

Unless defined otherwise, all technical and scientific terms used herein have the same meaning as commonly understood by one of ordinary skill in the art to which this invention is related. The following terms are defined for purposes of the invention as described herein.

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The term “affinity” as used herein refers to the strength of the binding interaction of two molecules, such as a metal chelating compound and a metal ion or a positively charged moiety and a negatively charged moiety.

25 The term “alkyl” as used herein refers to a straight, branched or cyclic hydrocarbon chain fragment containing between about one and about twenty five carbon atoms (e.g. methyl, ethyl and the like). Straight, branched or cyclic hydrocarbon chains having eight or fewer carbon atoms will also be referred to herein as “lower alkyl”. In addition, the term “alkyl” as used herein further includes one or more substitutions at one or more carbon atoms of the  
30 hydrocarbon chain fragment. Such substitutions include, but are not limited to: aryl; heteroaryl; halogen; alkoxy; amine (-NR'R''); carboxy and thio.

The term “amino” or “amine group” refers to the group -NR'R' (or NRR'R'') where R, R' and R'' are independently selected from the group consisting of hydrogen, alkyl, substituted alkyl,  
35 aryl, substituted aryl, aryl alkyl, substituted aryl alkyl, heteroaryl, and substituted heteroaryl. A substituted amine being an amine group wherein R' or R'' is other than hydrogen. In a primary amino group, both R' and R'' are hydrogen, whereas in a secondary amino group,



either, but not both, R' or R" is hydrogen. In addition, the terms "amine" and "amino" can include protonated and quaternized versions of nitrogen, comprising the group –NRR'R" and its biologically compatible anionic counterions.

- 5 The term "aryl" as used herein refers to cyclic aromatic carbon chain having twenty or fewer carbon atoms, e.g., phenyl, naphthyl, biphenyl, and anthracenyl. One or more carbon atoms of the aryl group may also be substituted with, e.g., alkyl; aryl; heteroaryl; a halogen; nitro; cyano; hydroxyl, alkoxyl or aryloxy; thio or mercapto, alkyl-, or arylthio; amino, alkylamino, arylamino, dialkyl-, diaryl-, or arylalkylamino; aminocarbonyl, alkylaminocarbonyl, arylaminocarbonyl, dialkylaminocarbonyl, diarylaminocarbonyl, or arylalkylaminocarbonyl; carboxyl, or alkyl- or aryloxy carbonyl; aldehyde; aryl- or alkylcarbonyl; iminyl, or aryl- or alkyliminyl; sulfo; alkyl- or alkylcarbonyl; iminyl, or aryl- or alkyliminyl; sulfo; alkyl- or arylsulfonyl; hydroximiny, or aryl- or alkoximiny. In addition, two or more alkyl or heteroalkyl substituents of an aryl group may be combined to form fused aryl-alkyl or aryl-heteroalkyl ring systems (e.g., tetrahydronaphthyl). Substituents including heterocyclic groups (e.g., heteroaryloxy, and heteroaralkylthio) are defined by analogy to the above-described terms.
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The term "arylalkyl" as used herein refers to an aryl group that is joined to a parent structure by an alkyl group as described above, e.g., benzyl,  $\alpha$ -methylbenzyl, phenethyl, and the like.

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The term "attachment site" as used herein refers to a site on a moiety or a molecule, e. g. a metal chelating compound, a fluorescent dye, a peptide or a protein, to which is covalently attached, or capable of being covalently attached, to a linker or another moiety.

- 25 The term "aqueous solution" as used herein refers to a solution that is predominantly water and retains the solution characteristics of water. Where the aqueous solution contains solvents in addition to water, water is typically the predominant solvent.

- The term "cell permeable" as used herein refers to compounds of the present invention that are able to cross the cell membrane of live cells. Lipophilic groups that are covalently attached to the present compounds, typically on the DYE moiety, facilitate this permeability and live cell entry. Once inside the cells, the lipophilic groups are hydrolyzed resulting in charged molecules that are well retained in living cells. Particularly useful lipophilic groups include acetoxymethyl (AM) ester and acetate esters wherein once inside the cells the groups are cleaved by nonspecific esterases resulting in charged molecules.
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The term “complex” as used herein refers to the association of two or more molecules, usually by non-covalent bonding.

The term “detectable response” as used herein refers to a change in or an occurrence of, a signal that is directly or indirectly detectable either by observation or by instrumentation and the presence or magnitude of which is a function of the presence of a target metal ion in the test sample. Typically, the detectable response is an optical response resulting in a change in the wavelength distribution patterns or intensity of absorbance or fluorescence or a change in light scatter, fluorescence quantum yield, fluorescence lifetime, fluorescence polarization, a shift in excitation or emission wavelength or a combination of the above parameters. The detectable change in a given spectral property is generally an increase or a decrease. However, spectral changes that result in an enhancement of fluorescence intensity and/or a shift in the wavelength of fluorescence emission or excitation are also useful. The change in fluorescence on ion binding is usually due to conformational or electronic changes in the indicator that may occur in either the excited or ground state of the fluorophore, due to changes in electron density at the ion binding site, due to quenching of fluorescence by the bound target metal ion, or due to any combination of these or other effects. Alternatively, the detectable response is an occurrence of a signal wherein the fluorophore is inherently fluorescent and does not produce a change in signal upon binding to a metal ion or biological compound.

The term “DYE” as used herein refers to a reporter molecule that is inherently fluorescent or demonstrates a change in fluorescence upon binding to a biological compound or metal ion, i.e., fluorogenic. Numerous fluorophores are known to those skilled in the art and include, but are not limited to, coumarin, acridine, furan, indole, borapolyazaindacene, cyanine, benzofuran, quinazolinone, benzazole, oxazine and xanthenes including fluorescein, rhodamine, rosamine and rhodol, as well as other fluorophores described in RICHARD P. HAUGLAND, MOLECULAR PROBES HANDBOOK OF FLUORESCENT PROBES AND RESEARCH PRODUCTS (9<sup>th</sup> edition, CD-ROM, 2002). The DYE moiety may be substituted by substituents that enhance solubility, live cell permeability and alter spectra absorption and emission.

The term “heteroaryl” or “heteroaromatic ring system” as used herein refers to a 5- or 6-membered unsaturated ring that is optionally fused to an additional six-membered aromatic ring(s), or is fused to one 5- or 6-membered unsaturated ring containing one or more heteroatoms, and is optionally substituted as defined below. Each heteroaromatic ring contains at least 1 and as many as 3 heteroatoms that are selected from the group

consisting of O, N or S in any combination. Specific examples of a heteroaryl ring system include, but are not limited to, substituted or unsubstituted derivatives of 2- or 3-furanyl; 2- or 3-thienyl; N-, 2- or 3-pyrrolyl; 2- or 3-benzofuranyl; 2- or 3-benzothieryl; N-, 2- or 3-indolyl; 2-, 3- or 4-pyridyl; 2-, 3- or 4-quinolyl; 1-, 3-, or 4-isoquinolyl; 2-, 4-, or 5-(1,3-oxazolyl); 2-benzoxazolyl; 2-, 4-, or 5-(1,3-thiazolyl); 2-benzothiazolyl; 3-, 4-, or 5-isoxazolyl; N-, 2-, or 4-imidazolyl; N-, or 2-benzimidazolyl; 1- or 2-naphthofuranyl; 1- or 2-naphthothieryl; N-, 2- or 3-benzindolyl; 2-, 3-, or 4-benzoquinolyl; 1-, 2-, 3-, or 4-acridinyl. Preferably the heteroaryl substituent is substituted or unsubstituted 4-pyridyl, 2-thienyl, 2-pyrrolyl, 2-indolyl, 2-oxazolyl, 2-benzothiazolyl or 2-benzoxazolyl. More preferably, the heteroaryl substituent is 2-thienyl or 2-pyrrolyl.

The term "kit" as used refers to a packaged set of related components, typically one or more compounds or compositions.

The term "Linker" or "L" as used herein refers to a single covalent bond or a series of stable covalent bonds incorporating 1-20 nonhydrogen atoms selected from the group consisting of C, N, O, S and P that covalently attach the crown ether metal chelating compounds to another moiety such as a chemically reactive group, a DYE or a conjugated substance including biological and non-biological substances. A "cleavable linker" is a linker that has one or more covalent bonds that may be broken by the result of a reaction or condition. For example, an ester in a molecule is a linker that may be cleaved by a reagent, e.g. sodium hydroxide, resulting in a carboxylate-containing fragment and a hydroxyl-containing product.

The term "metal chelator" or "metal chelating compound" as used herein refers to a chemical compound that combines with a metal ion to form a chelate ring structure.

The term "metal ion" or "target metal ion" as used herein refers to any metal cation that is capable of being chelated by the present crown ether chelate compounds. Typically, these metal ions are physiological and or nutritional relevant metal ion such as  $\text{Na}^+$ ,  $\text{K}^+$ ,  $\text{Zn}^{2+}$  and  $\text{Ca}^{2+}$ . The term metal ion used herein also refers to the metal ions  $\text{Li}^+$  and  $\text{Rb}^+$ .

The terms "protein" and "polypeptide" are used herein in a generic sense to include polymers of amino acid residues of any length. The term "peptide" is used herein to refer to polypeptides having less than 250 amino acid residues, typically less than 100 amino acid residues. The terms apply to amino acid polymers in which one or more amino acid residues are an artificial chemical analogue of a corresponding naturally occurring amino acid, as well as to naturally occurring amino acid polymers.

The term "Rx" or "reactive group" as used herein refers to a group that is capable of reacting with another chemical group to form a covalent bond, i.e. is covalently reactive under suitable reaction conditions, and generally represents a point of attachment for another substance. The reactive group is a moiety, such as carboxylic acid or succinimidyl ester, on the compounds of the present invention that is capable of chemically reacting with a functional group on a different compound to form a covalent linkage. Reactive groups generally include nucleophiles, electrophiles and photoactivatable groups.

The term "reporter molecule" as used herein refers to a chemical moiety that when covalently attached to the present crown ether compound is capable of generating a detectable response. Typically the reporter molecule is a DYE moiety.

The term "sample" as used herein refers to any material that may contain target metal ions, as defined above. Typically, the sample is a live cell or a biological fluid that comprises endogenous host cell proteins. Alternatively, the sample may be a buffer solution or an environmental sample containing target metal ions. The sample may be in an aqueous solution, a viable cell culture or immobilized on a solid or semi solid surface such as a polyacrylamide gel, membrane blot or on a microarray.

## II. COMPOSITIONS AND METHODS OF USE

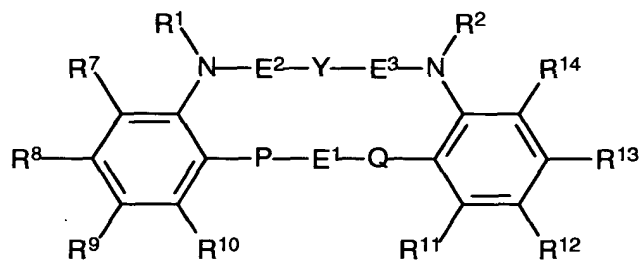
### A. Components of the Crown Ether Chelate Compounds

The present invention provides derivatives of crown ether compounds that bind a wide range of metal cations including physiological relevant levels of metal cations such as sodium. These metal chelating compounds comprise a crown ether moiety, at least one benzo moiety, substituents well known in the art including linkers, chemically reactive groups and DYE moieties that function as reporter groups. The crown ether moiety contains at least four heteroatoms, one of which is required to be a nitrogen atom and is located ortho to the benzo moiety. The remaining heteroatoms may be selected from the group consisting of nitrogen, oxygen and sulfur and are selected based on their ability to bind different metal cations with different affinity. Typically, in addition to the nitrogen atom there is also an oxygen atom ortho to the benzo moiety that facilitates binding of target metal ions. Furthermore, the substituents on the nitrogen atom are further used to alter the affinity for particular metal ions under different environmental conditions.

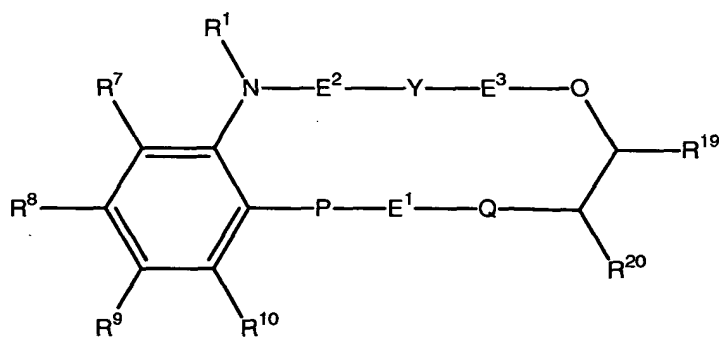
The present compounds find utility in binding target metal ions in a sample. The sample includes live cells or a biological fluid that comprises endogenous host cell proteins, buffer solutions and environmental samples. Therefore, when the present crown ether compounds comprise a DYE moiety they find utility in quantitating, monitoring and detecting target metal ions. Typically, the DYE moiety is directly attached to the benzo moiety or two of the benzo substituents when taken in combination form a fused DYE moiety. Detection of target metal ions can also be accomplished in live cells wherein the DYE moiety comprises a lipophilic group such as an AM or acetate ester that allows for entry across the live cell membrane. Once inside the cells nonspecific esterases cleave the AM or acetate ester resulting a charged molecule that is well retained in the cell. These present compounds are particularly useful for binding physiological relevant levels of sodium, potassium or calcium cations.

### 1. Chelating Moiety

The compounds of the invention are represented by the following two formulas:



Formula (I) or



Formula (II).

The heteroatom Y is O, S, NR<sup>4</sup> or is absent where R<sup>4</sup> is H, a C<sub>1</sub>-C<sub>18</sub> alkyl, or an aryl or heteroaryl ring system. The R<sup>4</sup> alkyl or ring system substituent is optionally substituted one or more times by halogen, azido, nitro, nitroso, amino, alkylamino having 1-6 carbons,

dialkylamino having 2-12 carbons, cyano, or  $R^4$  is substituted one or more times by a  $C_1$ - $C_6$  alkyl or  $C_1$ - $C_6$  alkoxy that is itself optionally substituted one or more times by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ .

Alternatively,  $R^4$  is, or is substituted by  $-L-R_X$ ,  $-L-S_C$ , or  $-L-DYE$ .

5

In one aspect of the invention Y is O and in another aspect Y is absent provided that  $E^2$  is also absent.

10

$R^{15}$  is independently H or  $C_1$ - $C_6$  alkyl. Each  $R^{16}$  is H, a  $C_1$ - $C_6$  alkyl, a benzyl, or forms an ester (e.g.,  $R^{16}$  is an alpha-acyloxyalkyl, a trialkylsilyl, or any other biologically compatible esterifying group). Additionally, any  $R^{16}$  is a biologically compatible salt.  $R^{17}$  and  $R^{18}$  are independently H or  $C_1$ - $C_6$  alkyl or  $C_1$ - $C_6$  carboxyalkyl, or an alpha-acyloxyalkyl, trialkylsilyl, or any other biologically compatible esterifying group, or a biologically compatible salt.

15

Alternatively,  $R^{17}$  and  $R^{18}$  when taken in combination form a 5- or 6-membered aliphatic ring that optionally incorporates an oxygen atom. In addition, one or more of a  $R^{15}$ ,  $R^{16}$ ,  $R^{17}$ , or  $R^{18}$  is permitted to be  $-L-R_X$ ,  $-L-S_C$ , or  $-L-DYE$ .

20

Each L is independently a covalent linkage. Each  $R_X$  is independently a chemically reactive group. Each  $S_C$  is independently a conjugated substance. Each DYE is independently a reporter molecule that is a chromophore that maximally absorbs light at a wavelength greater than 320 nm.

25

The heteroatoms P and Q are independently selected from O, S, or  $NR^3$ , where each  $R^3$  is independently H or an alkyl having 1-6 carbons. In one aspect of the invention P and Q are both O.

30

In another aspect of the invention, P and Q are O, and Y is  $NR^4$ . Typically, P, Q, and Y are each O or Y is absent and P and Q are O. Careful selection of the nature of the P, Q, and Y heteroatoms permits the moderation of the selectivity and binding affinity of the resulting crown ether compound.

35

$E^1$ ,  $E^2$ , and  $E^3$  each independently have the formula  $-(CR^5_2)_n-$ , or  $-[C(O)CH_2]_n-$ ,  $-(CR^5_2)_nO(CR^5_2)_n-$  or  $E^2$  is absent where n is 2, 3 or 4. Each  $R^5$  is independently H or methyl, or the  $R^5$  moieties on adjacent carbon atoms of each chain, when taken in combination, forms a 5- or 6-membered aliphatic ring. For a given E moiety, each  $R^5$  is typically H and n is 2. Where n is 2 for each E moiety, the resulting compound is known as a 15-crown-5 crown ether, having 15 atoms in the chelating ring itself, of which 5 are heteroatoms.

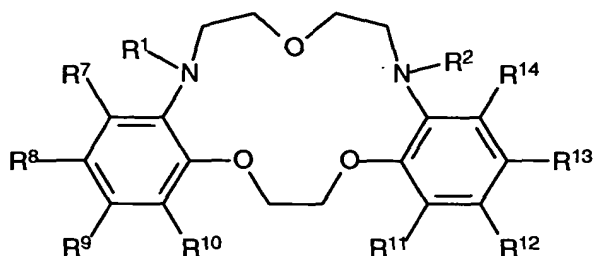
Alternatively, E<sub>2</sub> is absent or a E moiety is  $-(CR^5_2)_nO(CR^5_2)_n-$  resulting in a chelating ring with a different number of total atoms and possible heteroatoms such as oxygen.

In one aspect of the invention E<sup>1</sup>, E<sup>2</sup>, and E<sup>3</sup> are each  $-(CH_2)_2-$  and Y, P and Q are each oxygen. In another aspect, E<sup>2</sup> is absent and P and Q are oxygen wherein Y is absent and E<sup>1</sup> and E<sup>3</sup> are each  $-(CH_2)_2-$ . In yet another aspect of the invention at least one of E<sup>1</sup>, E<sup>2</sup>, or E<sup>3</sup> is  $-(CR^5_2)_nO(CR^5_2)_n-$  wherein Y, P and Q are each oxygen and the remaining E moieties are  $-(CH_2)_2-$ .

Formula (I) contains the amine substituents R<sup>1</sup> and R<sup>2</sup>, however Formula (II) contains only R<sup>1</sup>, the second nitrogen having been replaced by a divalent oxygen atom. Thus, it is understood that while the amine substituents, R<sup>1</sup> and R<sup>2</sup>, are referred to in plural only one is intended for compounds represented by formula (II). These amine substituents are independently H, C<sub>1</sub>-C<sub>18</sub> alkyl, or C<sub>7</sub>-C<sub>18</sub> arylalkyl. Where R<sup>1</sup> or R<sup>2</sup> is alkyl or arylalkyl, it is optionally substituted one or more times by halogen, azido, nitro, nitroso, amino, hydroxy, C<sub>1</sub>-C<sub>6</sub> alkylamino, C<sub>2</sub>-C<sub>12</sub> dialkylamino, cyano, or by an aryl or heteroaryl ring system, or by  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ , or by a C<sub>1</sub>-C<sub>6</sub> alkyl or C<sub>1</sub>-C<sub>6</sub> alkoxy that is itself optionally substituted one or more times by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ .  
Alternatively, R<sup>1</sup> and R<sup>2</sup> are optionally  $-L-R_x$ ,  $-L-S_c$ , or  $-L-DYE$ .

Where the R<sup>1</sup> and R<sup>2</sup> substituents are not  $-L-R_x$ ,  $-L-S_c$ , or  $-L-DYE$ , they are typically both a lower alkyl that is substituted one or more times by carboxylic acids, by carboxylic acid esters, by carboxylic acid amides, or by cyano. Where R<sup>1</sup> and R<sup>2</sup> incorporate carboxylic acid esters they are typically not cleaved by esterase but instead function to stabilize the bound metal ion in the chelate ring. Thus R<sup>1</sup>, and R<sup>2</sup> when present, are typically a methyl or ethyl that is substituted by a carboxylic acid containing group such as  $-(C=O)-R^{15}$  or  $-(C=O)-O-R^{16}$  wherein R<sup>15</sup> and R<sup>16</sup> are each a methyl or ethyl. Selection of the precise nature of R<sup>1</sup> and R<sup>2</sup> can greatly affect the binding selectivity and affinity of the resulting compound for a target ion (see, Table 1 and 2).

The nature of the R<sup>1</sup> and R<sup>2</sup> substituents in a large part determines the response of the indicators to particular target metal ions. For example, where the crown ether derivatives have the formula



where  $R^1$  and  $R^2$  are each methoxycarbonylmethyl selectively bind sodium ions with a dissociation constant ( $K_d$ ) of approximately 20-100 mM, and are relatively insensitive to the presence of potassium ions. In particular, the sodium ion  $K_d$  values typically rise less than about 10% when measured in the presence of 100 mM potassium ion.

The Formula (II) substituents  $R^{19}$  and  $R^{20}$  are represented by H, halogen, azido, nitro, nitroso, amino, cyano,  $-L-R_x$ ,  $-L-S_c$ ,  $-L-DYE$ ,  $C_1-C_6$  alkyl and  $C_1-C_6$  alkoxy, each of which is itself optionally substituted by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ . Alternatively,  $R^{19}$  and  $R^{20}$  taken in combination form a fused six-membered benzo moiety that is optionally substituted by halogen, azido, nitro, nitroso, amino, cyano,  $-L-R_x$ ,  $-L-S_c$ ,  $-L-DYE$ ,  $C_1-C_6$  alkyl or  $C_1-C_6$  alkoxy, each of which is itself optionally substituted by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ . In one aspect of the invention,  $R^{19}$  and  $R^{20}$  are both hydrogen. In another aspect,  $R^{19}$  and  $R^{20}$  form a benzo moiety that is optionally substituted resulting in a crown ether chelate compound with two benzo moieties wherein only one benzo moiety has a nitrogen atom that is ortho to the benzo moiety.

The benzo substituents  $R^7-R^{10}$ , and  $R^{11}-R^{14}$  when present, are independently H, halogen, azido, nitro, nitroso, amino, cyano; or  $-L-R_x$ ,  $-L-S_c$ , or  $-L-DYE$ ; or  $C_1-C_6$  alkyl or  $C_1-C_6$  alkoxy that is itself optionally substituted one or more times by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ .

Alternatively, two adjacent substituents of  $R^7-R^{14}$ , when taken in combination, form a fused six-membered benzo moiety, which is optionally substituted one or more times by a halogen, azido, nitro, nitroso, amino, cyano,  $-L-R_x$ ,  $-L-S_c$ , or  $-L-DYE$ ; or by a  $C_1-C_6$  alkyl or  $C_1-C_6$  alkoxy, which is itself optionally substituted one or more times by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ .

Furthermore, the present compounds comprise a fused DYE wherein two adjacent substituents of  $R^7-R^{14}$ , when taken in combination with each other, and with the aromatic ring



they are bound to, form a fused DYE. In this way the benzo moiety of the present compounds is also part of a DYE moiety.

The compounds of the invention represented by Formula (I) are substituted by at least one -L-DYE, -L-R<sub>x</sub> or -L-S<sub>c</sub> at one or more of R<sup>1</sup>, R<sup>2</sup>, R<sup>4</sup>, and R<sup>7</sup>-R<sup>14</sup>; or two of R<sup>7</sup>-R<sup>14</sup>, taken in combination, form a fused DYE. In one embodiment, the compounds of the invention are substituted by exactly one -L-DYE moiety, which is bound at R<sup>1</sup>, R<sup>2</sup>, an R<sup>4</sup>, or one of R<sup>7</sup>-R<sup>14</sup>, or is a fused DYE moiety at two adjacent substituents of R<sup>7</sup>-R<sup>14</sup>. The DYE moiety is typically bound at one of R<sup>7</sup>-R<sup>11</sup>, preferably at R<sup>9</sup> or R<sup>8</sup>, or is bound at R<sup>4</sup> where Y is NR<sup>4</sup>. In one embodiment, compounds that are substituted by exactly one -L-DYE moiety are optionally further substituted by -L-R<sub>x</sub> or -L-S<sub>c</sub>, typically at R<sup>1</sup>, R<sup>2</sup>, R<sup>4</sup>, or one of R<sup>7</sup>-R<sup>14</sup>.

In another embodiment, the compound of the invention is substituted by exactly two DYE moieties, which may be the same or different, and may be bound by a covalent linkage L or fused to the crown ether chelate. In one embodiment of the invention, a first -L-DYE moiety is bound at one of R<sup>7</sup>-R<sup>10</sup>, while the second -L-DYE moiety is bound at one of R<sup>11</sup>-R<sup>14</sup>. Typically, the first -L-DYE moiety is bound at R<sup>9</sup>, while the second -L-DYE moiety is bound at R<sup>12</sup>, or a DYE moiety is fused at R<sup>8</sup> and R<sup>9</sup> and additionally at R<sup>12</sup> and R<sup>13</sup>.

## **2. Linkers of the crown ether chelate compounds**

The crown ether chelate compounds of the present invention typically comprise a linker that is used to covalently attach a DYE moiety, conjugated substance or reactive group to the compound. When present, the linker is a single covalent bond or a series of stable bonds. Thus, the reporter molecule, conjugated substance or reactive group may be directly attached (where Linker is a single bond) to the crown ether chelate or attached through a series of stable bonds. When the linker is a series of stable covalent bonds the linker typically incorporates 1-20 nonhydrogen atoms selected from the group consisting of C, N, O, S and P. In addition, the covalent linkage can incorporate a platinum atom, such as described in U.S. Patent No. 5,714,327. When the linker is not a single covalent bond, the linker may be any combination of stable chemical bonds, optionally including, single, double, triple or aromatic carbon-carbon bonds, as well as carbon-nitrogen bonds, nitrogen-nitrogen bonds, carbon-oxygen bonds, sulfur-sulfur bonds, carbon-sulfur bonds, phosphorus-oxygen bonds, phosphorus-nitrogen bonds, and nitrogen-platinum bonds. Typically the linker incorporates less than 15 nonhydrogen atoms and are composed of any combination of ether, thioether, thiourea, amine, ester, carboxamide, sulfonamide, hydrazide bonds and aromatic or heteroaromatic bonds. Typically the linker is a single covalent bond or a combination of single carbon-carbon bonds and carboxamide, sulfonamide or thioether

bonds. The bonds of the linker typically result in the following moieties that can be found in the linker: ether, thioether, carboxamide, thiourea, sulfonamide, urea, urethane, hydrazine, alkyl, aryl, heteroaryl, alkoxyl, cycloalkyl and amine moieties. Examples of L include substituted or unsubstituted polymethylene, arylene, alkylarylene, arylenealkyl, or arylthio.

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In one embodiment, L contains 1-6 carbon atoms; in another, L comprises a thioether linkage. In another embodiment, L is or incorporates the formula  $-(CH_2)_d(CONH(CH_2)_e)_z-$  or  $-O(CH_2)_d(CONH(CH_2)_e)_z-$ , where d is an integer from 0-5, e is an integer from 1-5 and z is 0 or 1. In a further embodiment, L is or incorporates the formula  $-O-(CH_2)-$ . In yet another embodiment, L is or incorporates a phenylene or a 2-carboxy-substituted phenylene.

10

Any combination of linkers may be used to attach the DYE, Rx or Sc and the crown ether chelate together, typically a compound of the present invention when attached to more than one DYE, Rx or Sc will have one or two linkers attached that may be the same or different.

15

The linker may also be substituted to alter the physical properties of the crown ether chelate compound, such as binding affinity of the chelating moiety and spectral properties of the dye.

Another important feature of the linker is to provide an adequate space between the crown ether chelate moiety and the DYE, Rx or Sc so as to prevent these substituents from providing a steric hinderance to the binding of the target metal ion for the chelating moiety of the present compounds. Therefore, the linker of the present compounds is important for (1) attaching DYE, Rx or Sc to the metal chelating moiety, (2) providing an adequate space between DYE, Rx or Sc and the metal chelating moiety so as not to sterically hinder the affinity of the chelating moiety and the zinc ions and (3) for altering the affinity of the chelating moiety for the target ions either by the choice of the atoms of the linker or indirectly by addition of substituents to the linker.

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However, it is important to understand that a linker is not an essential component of the present compounds. Depending on the reporter molecule, a linker may not be necessary wherein the reporter molecule shares atoms with the metal chelating moiety, i.e. a fused DYE. A reporter molecule that exemplifies this is the dye benzofuran wherein one of the benzene rings of the dye is also one of the benzene rings of the metal chelating moiety. In addition, compounds represented for Formula (II) may not be substituted by a moiety that incorporates a linker.

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### 3. DYE moiety of the crown ether chelate compounds

The DYE moiety of the present invention functions as a reporter molecule to confer a

detectable signal, directly or indirectly, to the target metal ions. This results in the ability to detect, monitor and quantitate target metal ions in a sample.

5 The DYE moiety includes without limit a fluorophore, a chromophore, a fluorescent protein and an energy transfer pair. When the DYE moiety is a chromophore the crown ether chelate compounds are chromogenic indicators, or more preferably, the DYE moiety is a fluorophore, resulting in a compound that is a fluorogenic indicator for target ions, preferably sodium ions. Therefore, binding a sodium ion with a crown ether compound of the present invention results in a detectable optical response that can be correlated to the presence of  
10 sodium ions.

Where the detectable response is a fluorescence response, it is typically a change in fluorescence, such as a change in the intensity, excitation or emission wavelength distribution of fluorescence, fluorescence lifetime, fluorescence polarization, or a  
15 combination thereof. Preferably, the detectable optical response upon binding a target sodium ion is a change in fluorescence intensity that is greater than approximately 10-fold, more preferably greater than 50-fold, and most preferably more than 100-fold. This large increase in fluorescent signal over baseline has not been previously observed with other sodium indicators that comprise a different metal chelating moiety. In another aspect, the  
20 detectable optical response upon binding the target metal ion is a shift in maximal excitation or emission wavelength that is greater than about 20 nm, more preferably greater than about 30 nm. Sodium and potassium indicators of the type that exhibit significant excitation and/or emission shifts have not been previously described.

25 The DYE moiety is any chemical moiety that exhibits an absorption maximum beyond 320 nm, that is bound to the crown ether chelate by a covalent linkage L, or that is fused to the crown ether chelate. A preferred embodiment for detecting sodium ions in live cells is a fluorogenic crown ether chelate compound wherein the DYE moiety is substituted with a lipophilic group. As described above, the covalent linkage can be a single covalent bond or  
30 a combination of stable chemical bonds. The covalent linkage binding the DYE moiety to the crown ether chelator is typically a single bond, but optionally incorporates 1-20 nonhydrogen atoms selected from the group consisting of C, N, O, P, and S.

A wide variety of chemically reactive fluorescent dyes that may be suitable for incorporation  
35 into the compounds of the invention are already known in the art (RICHARD P. HAUGLAND, MOLECULAR PROBES HANDBOOK OF FLUORESCENT PROBES AND RESEARCH PRODUCTS (*Supra*); BIOPROBES 32 (December 1999); BIOPROBES 33

(February 2000); BIOPROBES 34 (May 2000); and BIOPROBES 35 (November 2000)).

The spectral properties of candidate dyes in solution or when conjugated to proteins such as IgG are known or are readily measured using an absorption spectrometer or a spectrofluorometer.

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Thus, the DYE moiety of the present invention include, without limitation; a pyrene, an anthracene, a naphthalene, an acridine, a stilbene, an indole or benzindole, an oxazole or benzoxazole, a thiazole or benzothiazole, a 4-amino-7-nitrobenz-2-oxa-1, 3-diazole (NBD), a carbocyanine (including any corresponding compounds in US Serial No. 09/557,275; US  
10 Publication Nos. 2002/0077487 and 2002/0064794 and US patent Nos. 6,403,807; 6,348,599; 5,486,616; 5,268,486; 5,569,587; 5,569,766; 5,627,027 and 6,048,982), a carbostyryl, a porphyrin, a salicylate, an anthranilate, an azulene, a perylene, a pyridine, a quinoline, a borapolyazaindacene (including any corresponding compounds disclosed in US Patent Nos. 4,774,339; 5,187,288; 5,248,782; 5,274,113; and 5,433,896), a xanthene  
15 (including any corresponding compounds disclosed in U.S. Patent No. 6,162,931; 6,130,101; 6,229,055; 6,339,392; 5,451,343 and US Publication No. 2002/0059684), an oxazine or a benzoxazine, a carbazine (including any corresponding compounds disclosed in US Patent No. 4,810,636), a phenalenone, a coumarin (including an corresponding compounds disclosed in US Patent Nos. 5,696,157; 5,459,276; 5,501,980 and 5,830,912), a benzofuran  
20 (including an corresponding compounds disclosed in US Patent Nos. 4,603,209 and 4,849,362) and benzphenalenone (including any corresponding compounds disclosed in US Patent No. 4,812,409) and derivatives thereof. As used herein, oxazines include resorufins (including any corresponding compounds disclosed in 5,242,805), aminooxazinones, diaminoxazines, and their benzo-substituted analogs.

25

Where the DYE moiety is a xanthene, the dye is optionally a fluorescein, a rhodol (including any corresponding compounds disclosed in US Patent Nos. 5,227,487 and 5,442,045), or a rhodamine (including any corresponding compounds in US Patent Nos. 5,798,276 and 5,846,737). As used herein, fluorescein includes benzo- or dibenzofluoresceins,  
30 seminaphthofluoresceins, or naphthofluoresceins. Similarly, as used herein rhodol includes seminaphthorhodafluors (including any corresponding compounds disclosed in U.S. Patent No. 4,945,171). Fluorinated xanthene dyes have been described previously as possessing particularly useful fluorescence properties (Int. Publ. No. WO 97/39064 and U.S. Patent No. 6,162,931).

35

Alternatively, the DYE moiety is a xanthene that is bound via an L that is a single covalent bond at the 9-position of the xanthene. Preferred xanthenes include derivatives of 3H-

xanthen-6-ol-3-one bound at the 9-position, derivatives of 6-amino-3*H*-xanthen-3-one bound at the 9-position, or derivatives of 6-amino-3*H*-xanthen-3-imine bound at the 9-position.

Preferred DYE moieties of the present invention include xanthene (including rhodol, fluorescein, rhodamine), benzofuran, indole, carbocyanine, quinazolinone, a benzazole, oxazine, coumarin and borapolyazaindacene. The xanthene dyes of this invention comprise both compounds substituted and unsubstituted on the carbon atom of the central ring of the xanthene by substituents typically found in the xanthene-based dyes such as phenyl and substituted-phenyl moieties. In addition, a preferred DYE moiety includes a xanthene-based moiety such as fluorescein that has a lipophilic group substituted on the oxygen atom. Most preferred dyes are rhodamine, fluorescein, borapolyazaindacene, indole and benzofuran. The choice of the dye attached to the chelating moiety will determine the crown ether chelate compound's absorption and fluorescence emission properties as well as its live cell properties, i.e. substituted lipophilic groups

Typically the DYE moiety contains one or more aromatic or heteroaromatic rings, that are optionally substituted one or more times by a variety of substituents, including without limitation, halogen, nitro, sulfo, cyano, alkyl, perfluoroalkyl, alkoxy, alkenyl, alkynyl, cycloalkyl, arylalkyl, acyl, aryl or heteroaryl ring system, benzo, or other substituents typically present on chromophores or fluorophores known in the art.

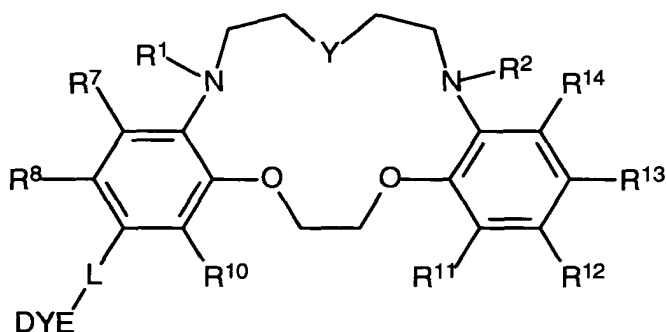
In one aspect of the invention, the DYE moiety has an absorption maximum beyond 480 nm. In a particularly useful embodiment, the DYE moiety absorbs at or near 488 nm to 514 nm (particularly suitable for excitation by the output of the argon-ion laser excitation source) or near 546 nm (particularly suitable for excitation by a mercury arc lamp). The DYE moiety may be a chromophore, resulting in a compound that acts as a chromogenic indicator, or more preferably, DYE is additionally a fluorophore, resulting in a compound that is a fluorescent indicator.

Selected sulfonated DYE moieties also exhibit advantageous properties, and include sulfonated pyrenes, coumarins, carbocyanines, and xanthenes (as described in U.S. Patent Nos. 5,132,432; 5,696,157; 5,268,486; 6,130,101). Sulfonated pyrenes and coumarins are typically excited at wavelengths below about 450 nm (US Patent Nos. 5,132,432 and 5,696,157).

Fluorescent proteins also find use as DYE moieties for the crown ether chelate compounds of the present invention. Examples of fluorescent proteins include green fluorescent protein

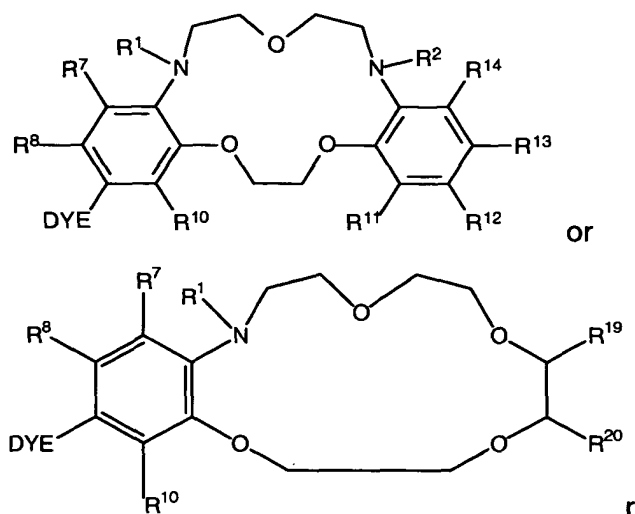
(GFP) and the phycobiliproteins and the derivatives thereof. The fluorescent proteins, especially phycobiliproteins, are particularly useful for creating tandem dye-reporter molecules. These tandem dyes comprise a fluorescent protein and a fluorophore for the purposes of obtaining a larger Stokes shift, wherein the emission spectra are farther shifted from the wavelength of the fluorescent protein's absorption spectra. This property is particularly advantageous for detecting a low quantity of a target sodium ion in a sample wherein the emitted fluorescent light is maximally optimized; in other words, little to none of the emitted light is reabsorbed by the fluorescent protein. For this to work, the fluorescent protein and fluorophore function as an energy transfer pair wherein the fluorescent protein emits at the wavelength that the acceptor fluorophore absorbs and the fluorophore then emits at a wavelength farther from the fluorescent proteins than could have been obtained with only the fluorescent protein. Alternatively, the fluorophore functions as the energy donor and the fluorescent protein is the energy acceptor. Particularly useful fluorescent proteins are the phycobiliproteins disclosed in US Patents 4,520,110; 4,859,582; 5,055,556 and the fluorophore bilin protein combinations disclosed in US Patent 4,542,104. Alternatively, two or more fluorophore dyes can function as an energy transfer pair wherein one fluorophore is a donor dye and the other is the acceptor dye including any dye compounds disclosed in US Patent Nos. 6,358,684; 5,863,727; 6,372,445; 6,221,606; 6,008,379; 5,945,526; 5,863,727; 5,800,996; 6,335,440; 6,008,373; 6,184,379; 6,140,494 and 5,656,554.

In one aspect of the invention, the compound of the invention has the formula



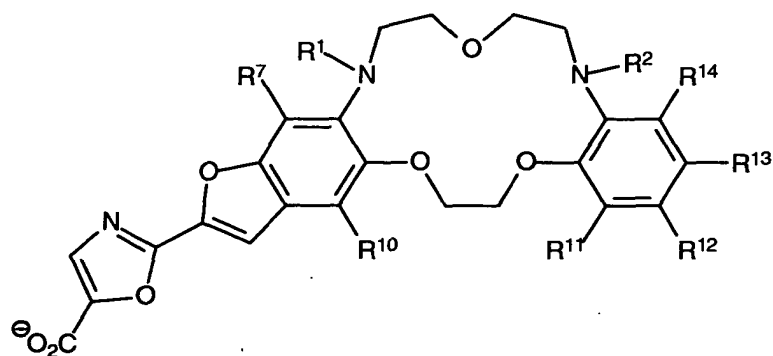
where Y is either O or NR<sup>4</sup>; and the DYE moiety is an indole, a coumarin, a stilbene, a xanthene, or a polyazaindacene. In this embodiment, preferably the DYE moiety is a xanthene, a polyazaindacene, or an oxazine.

In one aspect of the invention, the compounds of the invention are fluorescent indicators having the following structure:



- 5 wherein the linker is a single covalent bond. These indicators typically exhibit a low fluorescence quantum efficiency in the absence of metal ions. However, in the presence of increasing metal ion concentration the fluorescence quantum efficiency rises dramatically. For example, selected indicators of this family exhibit a fluorescence signal increase of over 100-times between zero and a saturating sodium concentration. Other selected indicators of
- 10 the invention exhibit a shift of the wavelength of the absorption (excitation) maximum, emission maximum, or both, upon binding the target ion. It appears that having the DYE moiety covalently attached or fused to the benzo moiety of the crown ether chelate compounds provides an additional channel, in addition to the chelate ring, to conduct electron density changes that occur upon metal binding, resulting in larger optical change.
- 15 This is true for compounds represented by Formula (I) or (II)

In another aspect of the invention, the compounds of the invention are fluorescent indicators having the following structure:



This class of indicators, in which at least one of the aromatic rings of the crown ether portion of the indicator is also incorporated in the DYE moiety, typically exhibit ratiometric fluorescence excitation changes in response to changing metal ion concentration. That is, there is a shift in the excitation maximum wavelength in the presence of increasing metal ion concentrations.

In one embodiment of this aspect,  $R^1$  and  $R^2$  are  $C_1$ - $C_6$  alkyl that are substituted one or more times by cyano, an aryl or heteroaryl ring system, or by  $-(C=O)-O-R^{16}$  or  $-(C=O)-NR^{17}R^{18}$ , where  $R^{16}$  is H, a  $C_1$ - $C_6$  alkyl, a benzyl, a biologically compatible esterifying group, or a biologically compatible salt; and  $R^{17}$  and  $R^{18}$  are independently H,  $C_1$ - $C_6$  alkyl,  $C_1$ - $C_6$  carboxyalkyl, an alpha-acyloxymethyl, or a biologically compatible salt. The substituents  $R^7$ ,  $R^{10}$ , and  $R^{11}$ - $R^{14}$ , are independently H, chloro, bromo, fluoro, nitro, amino, or cyano; or  $C_1$ - $C_6$  alkyl or  $C_1$ - $C_6$  alkoxy that is itself optionally substituted by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ .

In one embodiment of this aspect,  $R^1$  and  $R^2$  are  $C_1$ - $C_6$  alkyl that are substituted one or more times by cyano, an aryl or heteroaryl ring system, or by  $-(C=O)-O-R^{16}$  or  $-(C=O)-NR^{17}R^{18}$ , where  $R^{16}$  is H, a  $C_1$ - $C_6$  alkyl, a benzyl, a biologically compatible esterifying group, or a biologically compatible salt; and  $R^{17}$  and  $R^{18}$  are independently H,  $C_1$ - $C_6$  alkyl,  $C_1$ - $C_6$  carboxyalkyl, an alpha-acyloxymethyl, or a biologically compatible salt. The substituents  $R^7$ ,  $R^8$ ,  $R^{10}$ , and  $R^{11}$ - $R^{14}$ , are independently H, chloro, bromo, fluoro, nitro, amino, or cyano; or  $C_1$ - $C_6$  alkyl or  $C_1$ - $C_6$  alkoxy that is itself optionally substituted by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ . The DYE moiety is a polyazaindacene, an oxazine, or a xanthene, which is optionally substituted one or more times by halogen, nitro, sulfo, cyano, an aryl or heteroaryl ring system, or benzo, or alkyl, perfluoroalkyl, alkoxy, alkenyl, alkynyl, cycloalkyl, arylalkyl, acyl, or carboxylic acids or carboxylic acid esters, the alkyl portions of which contain fewer than 20 carbons.

An additional selected embodiment of the invention has the formula



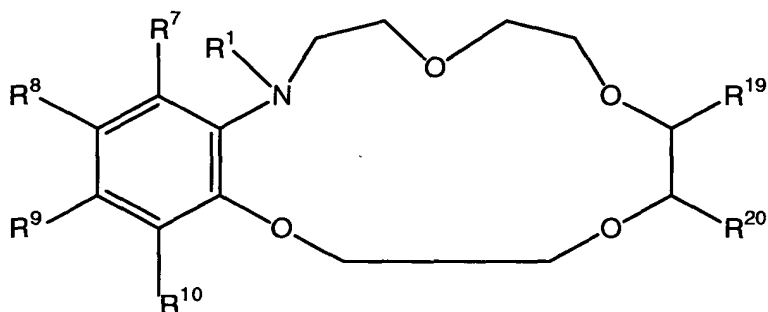


Chemical structure of a bis-oxazoline derivative. The molecule features two oxazoline rings connected by a central ethoxy chain. Each oxazoline ring is substituted with a carboxylate group ( $R^{16}-O_2C-$  and  $-CO_2-R^{16}$ ), a phenyl ring, and a 1H-indole-3-carboxylate group ( $CO_2-R^{26}$ ). The phenyl rings are further substituted with  $R^8$  and  $CH_3$  groups.



cycloalkyl, alkylthio, alkylamido, amino, monoalkylamino or dialkylamino wherein the alkyl portions of each contain fewer than 20 carbons.

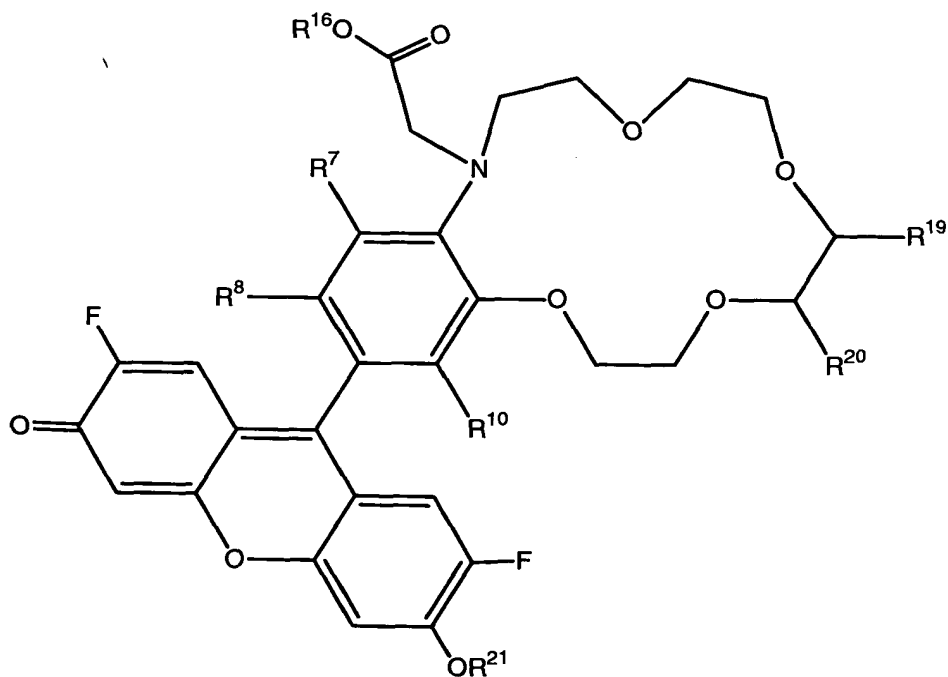
In another aspect of the invention, the crown ether chelate compounds are fluorescent indicators represented by the following structure:



Formula (II)(a and b)

wherein Y, P and Q are oxygen and R<sup>9</sup> or R<sup>8</sup> is represented by a -L-DYE. Preferably R<sup>9</sup> is -L-DYE wherein the linker is typically a single covalent bond and the DYE moiety is selected from the group consisting of borapolyazaindacene, xanthene and indole. Most preferred are xanthene and indole DYE moieties.

A particuary preferred embodiment of this aspect is represent by the following formula:



wherein  $R^{16}$  is selected from the group consisting of H,  $C_1$ - $C_6$  alkyl, benzyl, a biologically compatible esterifying group, and a biologically compatible salt. Preferably  $R^{16}$  is methyl, a biologically compatible esterifying group or a biologically compatible salt.

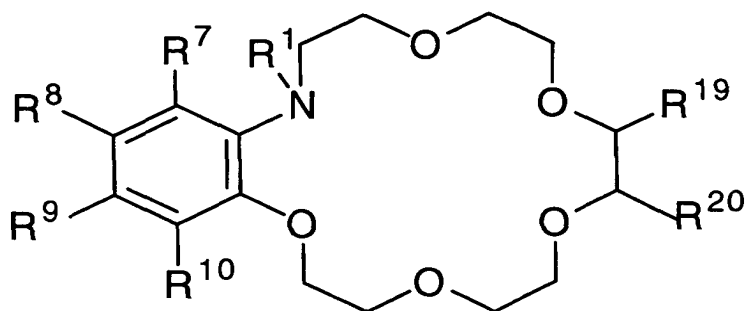
- 5  $R^{19}$  and  $R^{20}$  are selected from the group consisting of H, halogen, azido, nitro, nitroso, amino, cyano,  $-L-R_X$ ,  $-L-S_C$ ,  $-L-DYE$ ,  $C_1$ - $C_6$  alkyl and  $C_1$ - $C_6$  alkoxy, each of which is itself optionally substituted by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$  and  $-(C=O)-NR^{17}R^{18}$ . Alternatively,  $R^{19}$  and  $R^{20}$  taken in combination form a fused six-membered benzo moiety that is optionally substituted by halogen, azido, nitro, nitroso, amino, cyano,  $-L-R_X$ ,  $-L-S_C$ ,  $-L-DYE$ ,  $C_1$ - $C_6$  alkyl or  $C_1$ - $C_6$  alkoxy, each of which is itself optionally substituted by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ . Typically,  $R^{19}$  and  $R^{20}$  are hydrogen or taken together form a fused benzo moiety.

- 15  $R^7$ ,  $R^8$ , and  $R^{10}$  are independently selected from the group consisting of H, halogen, azido, nitro, nitroso, amino, cyano,  $-L-R_X$ ,  $-L-S_C$ ,  $-L-DYE$ ,  $C_1$ - $C_6$  alkyl and  $C_1$ - $C_6$  alkoxy, each of which is optionally substituted by halogen, amino, hydroxy,  $-(SO_2)-R^{15}$ ,  $-(SO_2)-O-R^{15}$ ,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ . Alternatively,  $R^7$  taken in combination with  $R^8$  form a fused six-membered benzo moiety, which is optionally substituted by halogen, azido, nitro, nitroso, amino, cyano,  $-L-R_X$ ,  $-L-S_C$ ,  $-L-DYE$ ,  $C_1$ - $C_6$  alkyl or  $C_1$ - $C_6$  alkoxy, each of which is optionally substituted by halogen, amino, hydroxy,  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $-(C=O)-NR^{17}R^{18}$ . Typically,  $R^7$ ,  $R^8$  and  $R^{10}$  are hydrogen.

$R^{21}$  is selected from the group consisting of H,  $C_1$ - $C_{18}$  alkyl,  $C_7$ - $C_{18}$  arylalkyl and lipophilic group each alkyl is optionally substituted by  $-(C=O)-R^{15}$ ,  $-(C=O)-O-R^{16}$ , or  $C_1$ - $C_6$  alkoxy.

- 25 A particular preferred embodiment is compound **86** and **115** and the corresponding cell-permeant versions, compounds **87** and **116** (See, Examples 79, 80, 105 and 106).

In another aspect of the invention, the crown ether chelate compounds are represented by the formula:

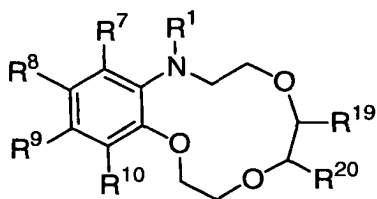


Formula (II)(c and d)

wherein  $R^9$  or  $R^8$  are  $-L-DYE$ , preferably  $R^9$  is  $-L-DYE$  wherein  $L$  is typically a single covalent bond and the DYE moiety is borapolyazaindacene, xanthene or indole. Most preferred are xanthene and indole DYE moieties.  $Y$ ,  $P$  and  $Q$  are oxygen.  $E^2$  and  $E^3$  are  $-(CH_2)_2-$  and  $E^1$  is  $-(CH_2)O(CH_2)-$ .

A preferred embodiment is represented by compound **109** and the cell-permeant version; compound **110** (See, Examples 99 and 100).

In another aspect of the invention, the crown ether chelate compounds are represented by the formula:



Formula (II)(e and f)

wherein  $R^9$  or  $R^8$  are  $-L-DYE$ , preferably  $R^9$  is  $-L-DYE$ , wherein  $L$  is typically a single covalent bond and the DYE moiety is borapolyazaindacene, xanthene or indole. Most preferred is a xanthene DYE moiety.  $P$  and  $Q$  are oxygen and  $Y$  is absent.  $E^1$  and  $E^3$  are  $-(CH_2)_2-$  and  $E^2$  is absent.

Typically  $R^7$ ,  $R^8$  or  $R^9$ ,  $R^{10}$ ,  $R^{19}$  and  $R^{20}$  are hydrogen and  $R^1$  is a methyl or ethyl group that is substituted by  $-(C=O)-R^{15}$  or  $-(C=O)-O-R^{16}$  wherein  $R^{15}$  and  $R^{16}$  are methyl or ethyl.

A preferred embodiment is represented by compound **104** (See, Example 94).

The family of crown ether chelate compounds represented by Formula (II) are particularly useful for binding physiological relevant levels of metal ions *in vivo* wherein the DYE moiety is substituted by an AM or acetate ester. In addition, this family of compounds unexpectedly

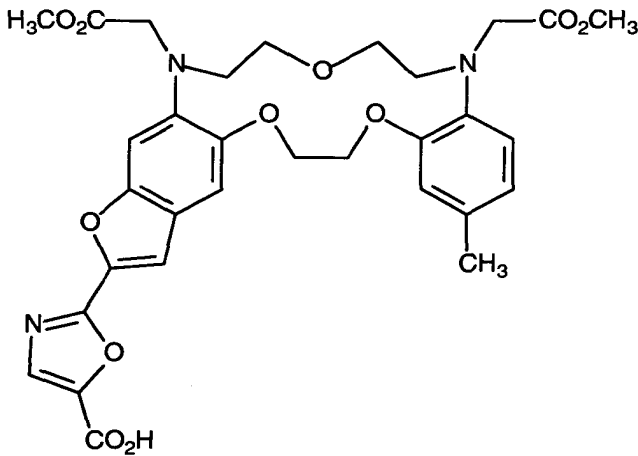
cross live cell membranes easier than compounds represented by Formula (I) and are thus preferred for binding of target metal ions in live cells. Furthermore, compounds represented by Formula (II), wherein the second nitrogen atom has been replaced by an oxygen atom, unexpectedly resulted in a higher affinity binding of target metal ions such as sodium ions.

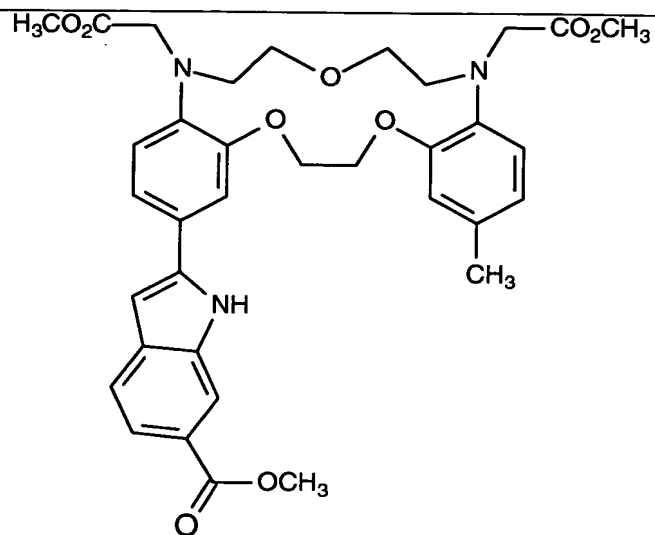
- 5 This, in combination with their ability to effectively load into live cells, provides for unexpected advantages over compounds represented by Formula (I) for *in vivo* binding of target ions.

- 10 The family of crown ether chelate compounds represented by Formula (II) are also very useful for the binding and detection of target metal ions *in vitro* (See, Table 2). These compounds demonstrate a significant change in fluorescent signal after binding the target metal ions.

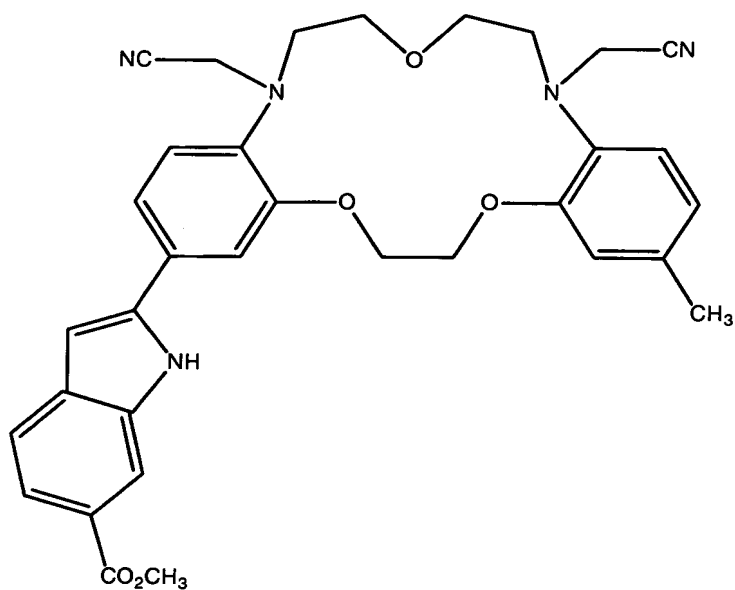
- 15 Selected embodiments of the invention are given in Table 1, showing a variety of distinct DYE moieties and crown ether substituents useful in the invention.

Table 1: Selected embodiments of the invention for Formula (I) compounds.

Compound	Dissociation Constant (target ion)
 <p>Compound 51</p>	$K_d(\text{Na}^+) = \sim 52 \text{ mM}$ $K_d(\text{K}^+) = \sim 330 \text{ mM}$
	$K_d(\text{Na}^+) = \sim 30 \text{ mM}$ $K_d(\text{K}^+) = \sim 115 \text{ mM}$

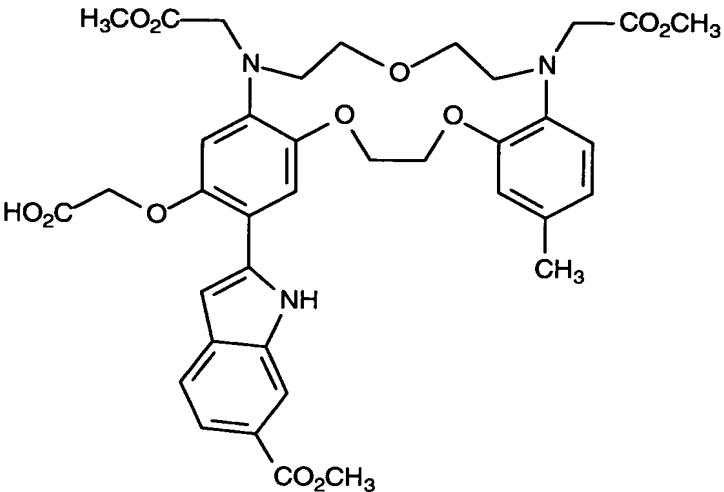
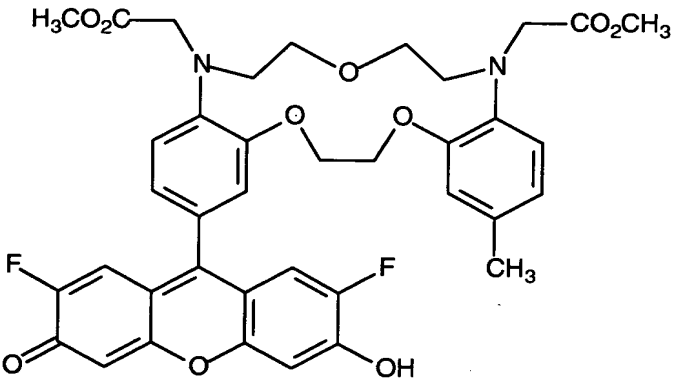
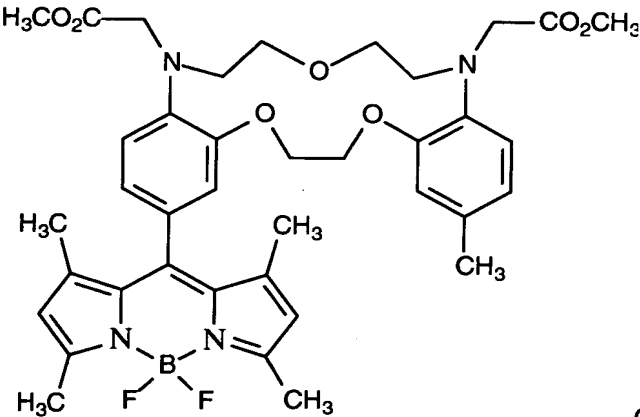


Compound 34

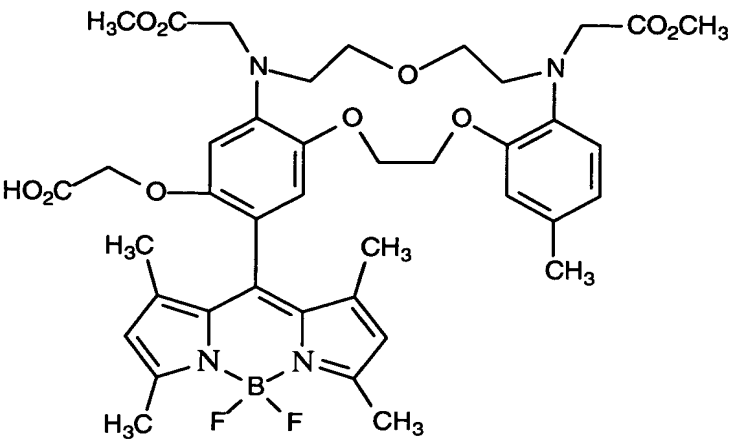
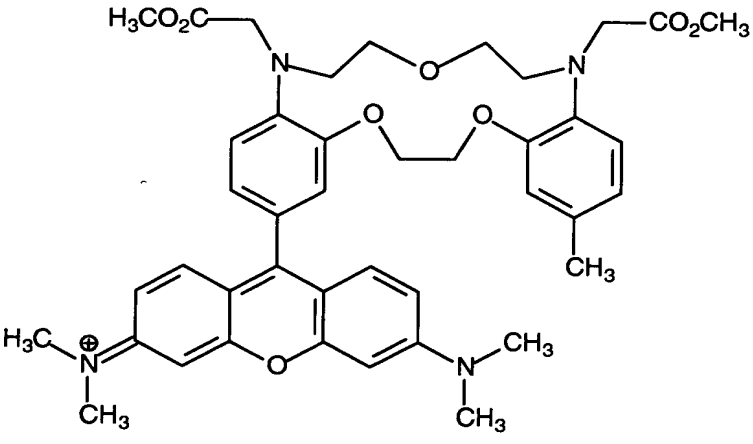
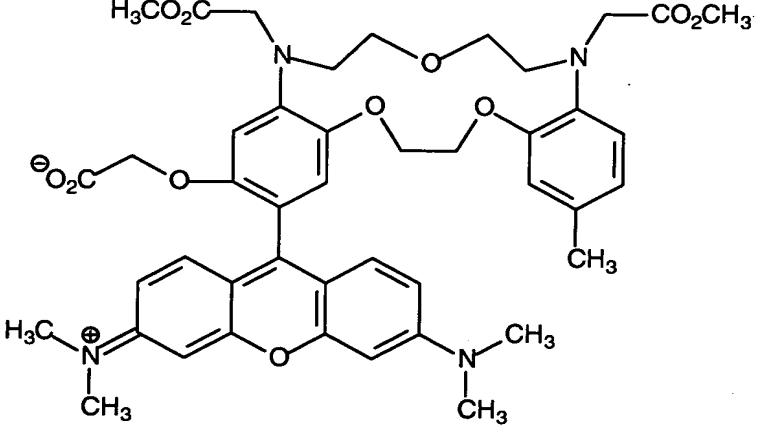


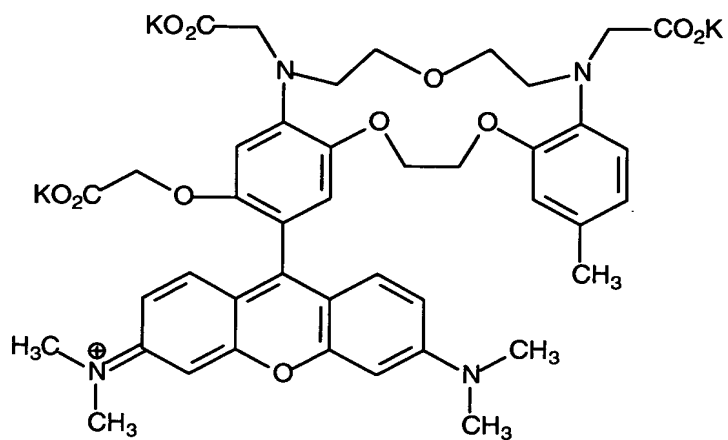
Compound 63

$K_d(\text{Na}^+) = \sim 220 \text{ mM}$

 <p style="text-align: right;"><b>Compound 57</b></p>	$K_d(\text{Na}^+) = \sim 52 \text{ mM}$ $K_d(\text{K}^+) = \sim 250 \text{ mM}$
 <p style="text-align: right;"><b>Compound 22</b></p>	$K_d(\text{Na}^+) = \sim 60 \text{ mM}$ $K_d(\text{K}^+) = \sim 205 \text{ mM}$
 <p style="text-align: right;"><b>Compound 32</b></p>	$K_d(\text{Na}^+) = \sim 42 \text{ mM}$



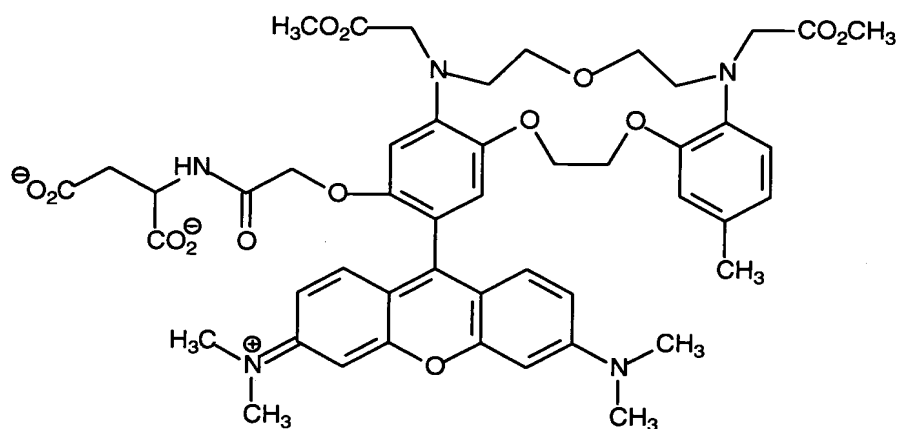
 <p>Chemical structure of Compound 79: A boronate ester derivative. The central boron atom is bonded to two fluorine atoms and two nitrogen atoms. Each nitrogen is part of a 2,4,6-trimethylphenyl group. The boronate ester is linked via a 1,3-bis(methoxycarbonylmethyl)oxy chain to a 4-methylphenyl group. The other end of the chain is linked to a 2,4,6-trimethylphenyl group via an ether linkage.</p>	<p><math>K_d (\text{Na}^+) = \sim 28 \text{ mM}</math>  <math>K_d (\text{K}^+) = \sim 130 \text{ mM}</math></p>
 <p>Chemical structure of Compound 27: A xanthine derivative. The central xanthine core is substituted with two dimethylamino groups at the 2 and 6 positions. The xanthine is linked via a 1,3-bis(methoxycarbonylmethyl)oxy chain to a 4-methylphenyl group. The other end of the chain is linked to a 2,4,6-trimethylphenyl group via an ether linkage.</p>	<p><math>K_d (\text{Na}^+) = \sim 95 \text{ mM}</math>  <math>K_d (\text{K}^+) = \sim 300 \text{ mM}</math></p>
 <p>Chemical structure of Compound 38: A xanthine derivative. The central xanthine core is substituted with two dimethylamino groups at the 2 and 6 positions. The xanthine is linked via a 1,3-bis(methoxycarbonylmethyl)oxy chain to a 4-methylphenyl group. The other end of the chain is linked to a 2,4,6-trimethylphenyl group via an ether linkage.</p>	<p><math>K_d (\text{Na}^+) = \sim 92 \text{ mM}</math>  <math>K_d (\text{K}^+) = \sim 705 \text{ mM}</math></p>



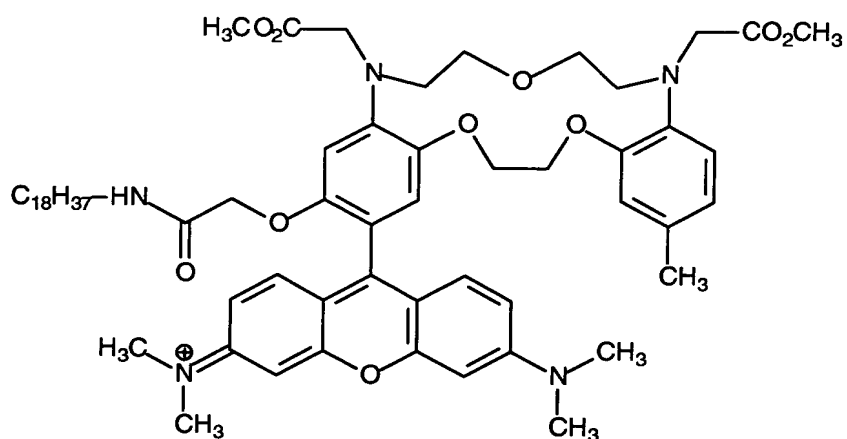
### Compound 60

$$K_d(\text{Na}^+) = \sim 70 \text{ mM}$$

$$K_d(\text{Zn}^{2+}) = \sim 100 \text{ nM}$$

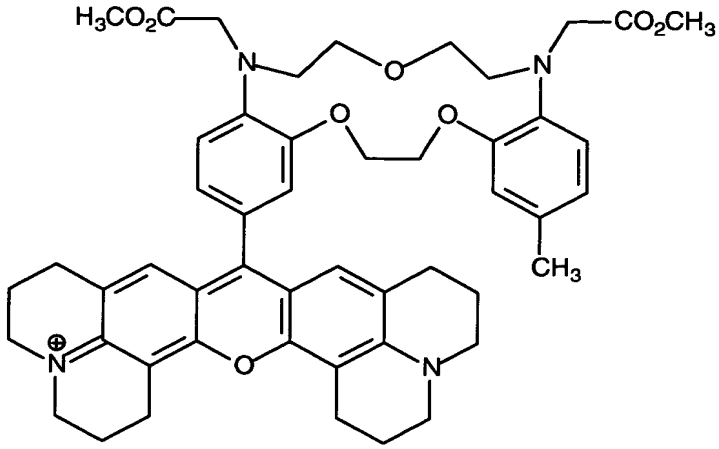
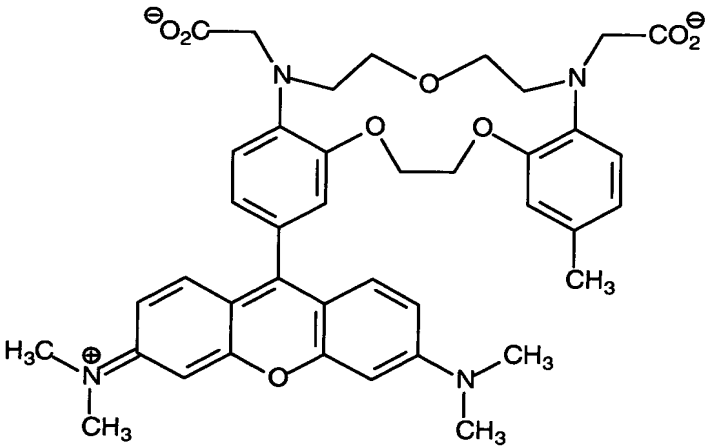
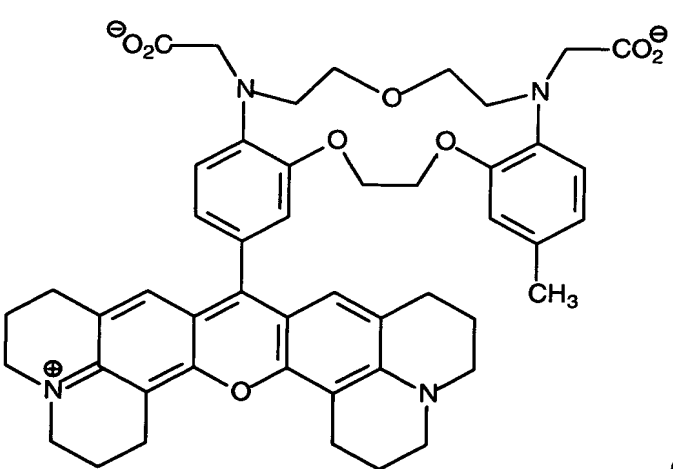


### Compound 78

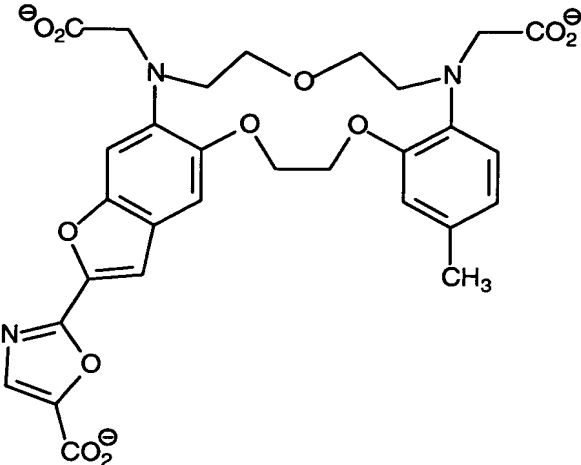
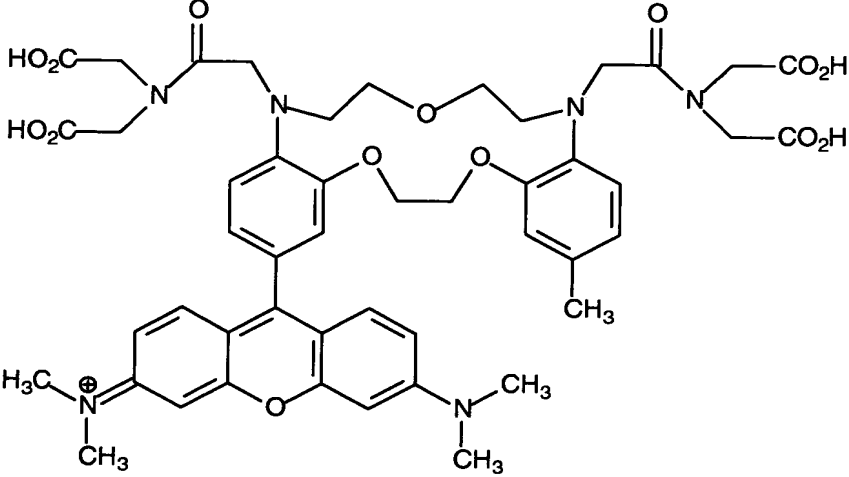
$$K_d(\text{Na}^+) = \sim 85 \text{ mM}$$
$$K_d(\text{K}^+) = \sim 255 \text{ mM}$$


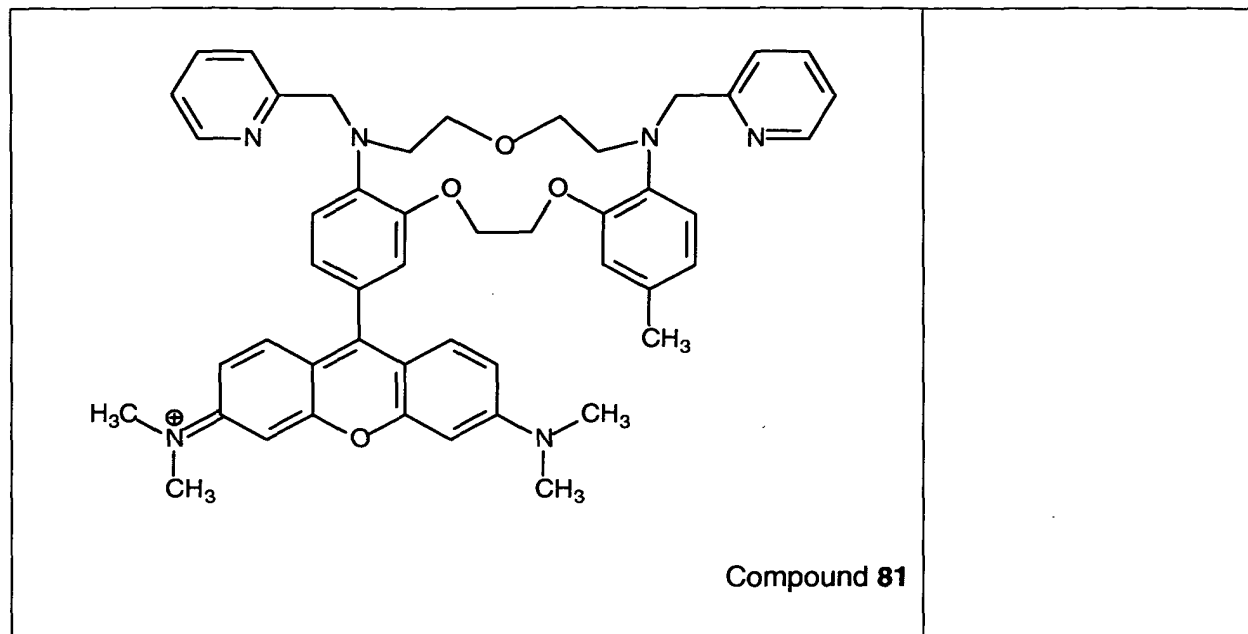
### Compound 42

$$K_d (\text{Na}^+) = \sim 14 \text{ mM}$$

 <p style="text-align: right;"><b>Compound 30</b></p>	$K_d (\text{Na}^+) = \sim 103 \text{ mM}$ $K_d (\text{K}^+) = \sim 205 \text{ mM}$
 <p style="text-align: right;"><b>Compound 28</b></p>	$K_d (\text{Zn}^{2+}) = \sim 300 \text{ nM}$ $K_d (\text{Ca}^{2+}) = \sim 4 \text{ }\mu\text{M}$ $K_d (\text{Na}^+) = \sim 38 \text{ mM}$ $K_d (\text{K}^+) = \sim 270 \text{ mM}$
 <p style="text-align: right;"><b>Compound 31</b></p>	$K_d (\text{Zn}^{2+}) = \sim 8 \text{ }\mu\text{M}$ $K_d (\text{Ca}^{2+}) = \sim 7 \text{ }\mu\text{M}$



 <p>Chemical structure of Compound 54, a bis-phenol derivative with a central ether linkage. The left phenol is substituted with a 2-(4-carboxyphenyl)-5-isoxazol-3-yl group and a 2-carboxymethyl group. The right phenol is substituted with a 4-methyl group and a 2-carboxymethyl group.</p>	<p><math>K_d</math> (<math>\text{Ca}^{2+}</math>) = ~ 400 nM  <math>K_d</math> (<math>\text{Na}^+</math>) = ~ 160 mM</p>
 <p>Chemical structure of Compound 80, a complex molecule featuring a central ether linkage between two phenols. The left phenol is part of a larger system including a 4-methyl-2-(dimethylamino)phenyl group and a 2-carboxymethyl group. The right phenol is substituted with a 4-methyl group and a 2-carboxymethyl group.</p>	<p>Compound 80</p>



Selected preferred embodiments for in solution binding of target ions with compounds represented by Formula (II):

5 Table 2:

Compound #	Dye moiety	R <sup>1</sup> =	Dissociation constant (K <sub>d</sub> ) and Response to complexation (R = F/F <sub>0</sub> ) for Target metal ions
<b>104</b>	xanthene	-CH <sub>2</sub> COOCH <sub>3</sub>	K <sub>d</sub> (Li <sup>+</sup> ) = 438 mM; R (Li <sup>+</sup> ) = 2.8 K <sub>d</sub> (Na <sup>+</sup> ) = 226 mM; R (Na <sup>+</sup> ) = 19.7 K <sub>d</sub> (K <sup>+</sup> ) = 978 mM; R (K <sup>+</sup> ) = 9.7 K <sub>d</sub> (Rb <sup>+</sup> ) = 376 mM; R (Rb <sup>+</sup> ) = 2.9
<b>86</b>	xanthene	-CH <sub>2</sub> COOCH <sub>3</sub>	K <sub>d</sub> (Li <sup>+</sup> ) = 142 mM; R (Li <sup>+</sup> ) = 6.8 K <sub>d</sub> (Na <sup>+</sup> ) = 82 mM; R (Na <sup>+</sup> ) = 28.5 K <sub>d</sub> (K <sup>+</sup> ) = 291 mM; R (K <sup>+</sup> ) = 6.9 K <sub>d</sub> (Rb <sup>+</sup> ) = 319 mM; R (Rb <sup>+</sup> ) = 2.7
<b>110</b>	xanthene	-CH <sub>2</sub> COOCH <sub>3</sub>	K <sub>d</sub> (Li <sup>+</sup> ) = no sensitivity K <sub>d</sub> (Na <sup>+</sup> ) = 15 mM; R (Na <sup>+</sup> ) = 1.5 K <sub>d</sub> (K <sup>+</sup> ) = 11 mM; R (K <sup>+</sup> ) = 2.2

			$K_d (\text{Rb}^+) = 25 \text{ mM}; R (\text{Rb}^+) = 1.5$
<b>93</b>	indole	$-\text{CH}_2\text{COOCH}_3$	$K_d (\text{Na}^+) = 89 \text{ mM}; R (\text{Na}^+) = 25$ $K_d (\text{K}^+) = \text{no sensitivity}$
<b>89</b>	xanthene	$-\text{CH}_2\text{COOCH}_3$	$K_d (\text{Na}^+) = 163 \text{ mM}; R (\text{Na}^+) = 41$ $K_d (\text{K}^+) = 254 \text{ mM}; R (\text{K}^+) = 10.6$
<b>90</b>	borapolyazaindacene	$-\text{CH}_2\text{COOCH}_3$	$K_d (\text{Na}^+) = 100 \text{ mM}; R (\text{Na}^+) = 7.3$ $K_d (\text{K}^+) = 170 \text{ mM}; R (\text{K}^+) = 2.9$
<b>127</b>	xanthene	$-\text{CH}_2\text{COOH}$	$K_d (\text{Na}^+) = 20 \text{ mM}; R (\text{Na}^+) = 3.6$
<b>115</b>	xanthene	$-(\text{CH}_2)_2\text{COCH}_3$	$K_d (\text{Na}^+) = 78.8 \text{ mM}; R (\text{Na}^+) = 5.9$ $K_d (\text{K}^+) = 269 \text{ mM}; R (\text{K}^+) = 2.8$
<b>118</b>	xanthene	$-(\text{CH}_2)_2\text{COCH}_3$	$K_d (\text{Na}^+) = 150 \text{ mM}; R (\text{Na}^+) = 12$
<b>120</b>	indole	$-(\text{CH}_2)_2\text{COCH}_3$	$K_d (\text{Na}^+) = 89 \text{ mM}; R (\text{Na}^+) = 2$
<b>126</b>	indole	$-(\text{CH}_2)_2\text{CN}(\text{CH}_3)_2$	$K_d (\text{Na}^+) = 379 \text{ mM}; R (\text{Na}^+) = -1.2$
<b>99</b>	xanthene	$-\text{CH}_2\text{COOCH}_3$	$K_d (\text{Li}^+) = 65 \text{ mM}; R (\text{Li}^+) = 3.1$ $K_d (\text{Na}^+) = 59 \text{ mM}; R (\text{Na}^+) = 3.9$ $K_d (\text{K}^+) = 144 \text{ mM}; R (\text{K}^+) = 2.7$ $K_d (\text{Rb}^+) = 157 \text{ mM}; R (\text{Rb}^+) = 1.9$

#### 4. Reactive Functional Groups and Conjugated Substances of the crown ether chelate compounds.

As described above, the compounds of the invention may be substituted one or more times by a-L-R<sub>x</sub> moiety or -L-S<sub>c</sub> moiety. L is a covalent linkage that is a single covalent bond or a series of stable bonds comprising 1-20 nonhydrogen atoms selected from the group consisting of C, O, N, P and S. R<sub>x</sub> is a reactive group that functions as the site of attachment for another moiety wherein the reactive group chemically reacts with an appropriate reactive or functional group on another substance or moiety. These reactive groups are synthesized during the formation of the present compounds providing present crown ether chelate compounds that can be covalently attached to another substance, conjugated substance, facilitated by the reactive group. In this way, compounds incorporating a reactive group (R<sub>x</sub>) can be covalently attached to a wide variety of biomolecules or non-biomolecules that contain or are modified to contain functional groups with suitable reactivity, resulting in chemical attachment of the conjugated substance (S<sub>c</sub>), represented by -L-S<sub>c</sub>. The reactive group and functional group are typically an electrophile and a nucleophile that can generate a covalent linkage. Alternatively, the reactive group is a photoactivatable group, and becomes chemically reactive only after illumination with light of an appropriate wavelength. Typically, the conjugation reaction between the reactive group and the substance to be conjugated results in one or more atoms of the reactive group R<sub>x</sub> to be incorporated into a new linkage attaching the compound of the invention to the conjugated substance S<sub>c</sub>. Selected examples of functional groups and linkages are shown in Table 3, where the reaction of an electrophilic group and a nucleophilic group yields a covalent linkage.

Table 3: Examples of some routes to useful covalent linkages

Electrophilic Group	Nucleophilic Group	Resulting Covalent Linkage
activated esters*	amines/anilines	carboxamides
acrylamides	thiols	thioethers
acyl azides**	amines/anilines	carboxamides
acyl halides	amines/anilines	carboxamides
acyl halides	alcohols/phenols	esters
acyl nitriles	alcohols/phenols	esters
acyl nitriles	amines/anilines	carboxamides
aldehydes	amines/anilines	imines
aldehydes or ketones	hydrazines	hydrazones
aldehydes or ketones	hydroxylamines	oximes



alkyl halides	amines/anilines	alkyl amines
alkyl halides	carboxylic acids	esters
alkyl halides	thiols	thioethers
alkyl halides	alcohols/phenols	ethers
alkyl sulfonates	thiols	thioethers
alkyl sulfonates	carboxylic acids	esters
alkyl sulfonates	alcohols/phenols	ethers
anhydrides	alcohols/phenols	esters
anhydrides	amines/anilines	carboxamides
aryl halides	thiols	thiophenols
aryl halides	amines	aryl amines
aziridines	thiols	thioethers
boronates	glycols	boronate esters
carbodiimides	carboxylic acids	N-acylureas or anhydrides
diazoalkanes	carboxylic acids	esters
epoxides	thiols	thioethers
haloacetamides	thiols	thioethers
haloplatinate	amino	platinum complex
haloplatinate	heterocycle	platinum complex
haloplatinate	thiol	platinum complex
halotriazines	amines/anilines	aminotriazines
halotriazines	alcohols/phenols	triazinyl ethers
halotriazines	thiols	triazinyl thioethers
imido esters	amines/anilines	amidines
isocyanates	amines/anilines	ureas
isocyanates	alcohols/phenols	urethanes
isothiocyanates	amines/anilines	thioureas
maleimides	thiols	thioethers
phosphoramidites	alcohols	phosphite esters
silyl halides	alcohols	silyl ethers
sulfonate esters	amines/anilines	alkyl amines
sulfonate esters	thiols	thioethers
sulfonate esters	carboxylic acids	esters
sulfonate esters	alcohols	ethers
sulfonyl halides	amines/anilines	sulfonamides

sulfonyl halides	phenols/alcohols	sulfonate esters
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\* Activated esters, as understood in the art, generally have the formula -CO $\Omega$ , where  $\Omega$  is a good leaving group (e.g., succinimidyl (-OC<sub>4</sub>H<sub>4</sub>O<sub>2</sub>) sulfosuccinimidyl (-OC<sub>4</sub>H<sub>3</sub>O<sub>2</sub>-SO<sub>3</sub>H), -1-oxybenzotriazolyl (-OC<sub>6</sub>H<sub>4</sub>N<sub>3</sub>); or an aryloxy group or aryloxy substituted one or more times by electron withdrawing substituents such as nitro, fluoro, chloro, cyano, or trifluoromethyl, or combinations thereof, used to form activated aryl esters; or a carboxylic acid activated by a carbodiimide to form an anhydride or mixed anhydride -OCOR<sup>a</sup> or -OCNR<sup>a</sup>NHR<sup>b</sup>, where R<sup>a</sup> and R<sup>b</sup>, which may be the same or different, are C<sub>1</sub>-C<sub>6</sub> alkyl, C<sub>1</sub>-C<sub>6</sub> perfluoroalkyl, or C<sub>1</sub>-C<sub>6</sub> alkoxy; or cyclohexyl, 3-dimethylaminopropyl, or N-morpholinoethyl).

\*\* Acyl azides can also rearrange to isocyanates

Choice of the reactive group used to attach the compound of the invention to the substance to be conjugated typically depends on the reactive or functional group on the substance to be conjugated and the type or length of covalent linkage desired. The types of functional groups typically present on the organic or inorganic substances (biomolecule or non-biomolecule) include, but are not limited to, amines, amides, thiols, alcohols, phenols, aldehydes, ketones, phosphates, imidazoles, hydrazines, hydroxylamines, disubstituted amines, halides, epoxides, silyl halides, carboxylate esters, sulfonate esters, purines, pyrimidines, carboxylic acids, olefinic bonds, or a combination of these groups. A single type of reactive site may be available on the substance (typical for polysaccharides or silica), or a variety of sites may occur (e.g., amines, thiols, alcohols, phenols), as is typical for proteins. A conjugated substance may be conjugated to more than one crown ether chelate compound, which may be the same or different, or to a substance that is additionally modified by a hapten, such as biotin. Although some selectivity can be obtained by careful control of the reaction conditions, selectivity of labeling is best obtained by selection of an appropriate reactive functional group.

Typically, R<sub>x</sub> will react with an amine, an alcohol, an aldehyde, a ketone, or with silica. Preferably R<sub>x</sub> reacts with an amine or a thiol functional group, or with silica. In one embodiment, R<sub>x</sub> is an acrylamide, an activated ester of a carboxylic acid, an acyl azide, an acyl nitrile, an aldehyde, an alkyl halide, a silyl halide, an anhydride, an aniline, an aryl halide, an azide, an aziridine, a boronate, a diazoalkane, a haloacetamide, a halotriazine, a hydrazine (including hydrazides), an imido ester, an isocyanate, an isothiocyanate, a maleimide, a phosphoramidite, a reactive platinum complex, a sulfonyl halide, or a thiol group. By "reactive platinum complex" is particularly meant chemically reactive platinum complexes such as described in U.S. Patent No. 5,714,327.

Where  $R_x$  is an activated ester of a carboxylic acid, the resulting compound is particularly useful for preparing conjugates of proteins, nucleotides, oligonucleotides, or haptens.

Where  $R_x$  is a maleimide or haloacetamide the resulting compound is particularly useful for conjugation to thiol-containing substances. Where  $R_x$  is a hydrazide, the resulting

compound is particularly useful for conjugation to periodate-oxidized carbohydrates and glycoproteins, and in addition is an aldehyde-fixable polar tracer for cell microinjection.

Where  $R_x$  is a silyl halide, the resulting compound is particularly useful for conjugation to silica surfaces, particularly where the silica surface is incorporated into a fiber optic probe subsequently used for remote ion detection or quantitation.

Preferably,  $R_x$  is a succinimidyl ester of a carboxylic acid, a haloacetamide, a hydrazine, an isothiocyanate, a maleimide group, an aliphatic amine, a silyl halide, or a psoralen. More preferably,  $R_x$  is a succinimidyl ester of a carboxylic acid, a maleimide, an iodoacetamide, or a silyl halide. In a particular embodiment  $R_x$  is a silyl halide or an isothiocyanate.

The compounds of the invention that possess a reactive functional group are useful for conjugation to any substance that possesses a suitable functional group for covalent attachment of the chelate. Examples of particularly useful conjugates include, among others, conjugates of antigens, steroids, vitamins, drugs, haptens, metabolites, toxins, environmental pollutants, amino acids, peptides, proteins, nucleic acids, nucleic acid polymers, carbohydrates, lipids, and non-biological polymers. Alternatively, these are conjugates of cells, cellular systems, cellular fragments, or subcellular particles. Examples include, among others, virus particles, bacterial particles, virus components, biological cells (such as animal cells, plant cells, bacteria, yeast, or protists), or cellular components.

Preferably the conjugated substance is a protein, polysaccharide, lipid, lipid assembly, non-biological polymer, or polymeric microparticle. Another class of preferred conjugated substances includes particles or fibers composed of silica or other glasses, useful for preparing optical devices for remote sensing.

Preferred nucleic acid polymer conjugates are labeled, single- or multi-stranded, natural or synthetic DNA or RNA, DNA or RNA oligonucleotides, or DNA/RNA hybrids, or incorporate an unusual linker such as morpholine derivatized phosphates (AntiVirals, Inc., Corvallis OR), or peptide nucleic acids such as *N*-(2-aminoethyl)glycine units.

In a preferred embodiment, the conjugated substance ( $S_c$ ) is a carbohydrate that is typically a polysaccharide, such as a dextran, FICOLL, heparin, glycogen, amylopectin, mannan,

inulin, starch, agarose and cellulose. Alternatively, the carbohydrate is a polysaccharide that is a lipopolysaccharide. Preferred polysaccharide conjugates are dextran, FICOLL, or lipopolysaccharide conjugates.

- 5 In another embodiment, the conjugated substance ( $S_c$ ), is a lipid moiety (typically having 6–60 carbons), including glycolipids, phospholipids, sphingolipids, and steroids. Alternatively, the conjugated substance is a lipid assembly, such as a liposome. The lipid moiety may be used to retain the conjugated substances in cells, as described in U.S. Pat. 5,208,148.
- 10 Other conjugates of non-biological materials include conjugates of organic or inorganic polymers, polymeric films, polymeric wafers, polymeric membranes, polymeric particles, or polymeric microparticles, including magnetic and non-magnetic microspheres, conducting and non-conducting metals and non-metals, and glass and plastic surfaces and particles. Conjugates are optionally prepared by copolymerization of a compound of the invention that
- 15 contains an appropriate functionality while preparing the polymer, or by chemical modification of a polymer that contains functional groups with suitable chemical reactivity. Other types of reactions that are useful for preparing conjugates of polymers include catalyzed polymerizations or copolymerizations of alkenes and reactions of dienes with dienophiles, transesterifications or transaminations. In another embodiment, the conjugated
- 20 substance is a glass or silica, which may be formed into an optical fiber or other structure.

Conjugates typically result from mixing appropriate reactive dyes and the substance to be conjugated in a suitable solvent in which both are soluble. Labeling of insoluble polymers or silica can be performed in a suspension of the insoluble polymer in a suitable solvent. For

25 those reactive groups that are photoactivated, conjugation requires illumination of the reaction mixture to activate the reactive group.

### Synthesis

- 30 There are typically three components to the methodology used to prepare the compounds of the invention. The first involves the formation of the crown ether chelate itself, the second involves the appropriate derivatization of the secondary amine nitrogen atom(s) of the crown ether chelate, and the third, when needed, involves modification of the crown ether chelate by forming a reactive functional group, covalently attaching a conjugate, or covalently
- 35 attaching a DYE moiety to form an indicator. It should be understood that the DYE moiety is typically not attached to the crown ether chelate compound but that the conjugated substance is attached in this way. Although these synthetic components are typically

performed in the order given, they may be carried out in any other suitable sequence. For example, a portion of the chelate may be derivatized with a fluorescent dye prior to formation of the complete chelate ring.

5 Where the P and Q moieties are both oxygen, and E<sup>1</sup> is ethylene, the crown ether chelate is typically prepared by acylation of a bis-(2-aminophenoxy)ethane with a bis-(acid chloride), such as diglycolyl chloride, followed by reduction of the resulting bis-amide to the corresponding bis-secondary amine. Selection of the appropriate bis-acid chloride results in the particular desired crown ether (as in Examples 64-69 and 70).

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The secondary amine nitrogen atom(s) present in the crown ether are typically derivatized with an alkylating agent. As the metal binding ability of the resulting crown ether is significantly influenced by the nature of the amine substituents, careful selection of the alkylating agent may be necessary to prepare a reporter for a particular target ion. Where  
15 the crown nitrogens are alkylated by methyl bromoacetate, the resulting bis-aza-crown ether is typically selective for sodium ions. If the alkylating agent is 2-picolyl chloride, the resulting crown ether is typically selective for zinc ions. As discussed above, the presence of esters vs. carboxylic acids on the amine nitrogen substituents may influence the relative binding affinity of selected target ions. Selection of an alkylating agent that incorporates a precursor  
20 to a reactive functional group is useful for producing chemically reactive compounds of the invention, as well as acting as a useful intermediate for preparing conjugates, as described above. Additionally, an alkylating agent that incorporates a reporter DYE results in a crown ether compound that functions as an indicator for selected target ions.

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25 More typically, the crown ether chelate is derivatized at the benzo ring of the crown ether. As described above, typically a suitable crown ether is prepared, and then bound to a DYE moiety. In one aspect of the invention, an ortho-hydroxy aromatic aldehyde is treated with a chloromethyl heterocycle to yield a fused reporter (as described in Example 46). In another aspect of the invention, derivatization with a DYE moiety is carried out by modifying a crown  
30 ether that possesses an aldehyde or ketone functional group.

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In one aspect of the invention, the crown ether is substituted by an aldehyde and the fluorophore precursors and the crown ether are combined under anaerobic or non-oxidative conditions (e.g., under nitrogen), and subsequently oxidized using a mild oxidant (e.g., a  
35 quinone oxidant, preferably DDQ or chloranil). Where xanthene fluorophore precursors are condensed under anaerobic conditions, the resulting fluorophore is the non-fluorescent dihydro species, which may be utilized without prior oxidation as a sensor for oxidative

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subenvironments, e.g., in cells.

In yet another aspect of the invention, the crown ether is substituted by a carboxylic acid or by an aldehyde that is converted to the carboxylic acid in the course of synthesis of the crown ether. The chelating moiety is then condensed with the fluorophore precursors to yield the resulting indicator directly.

Synthesis of conventional xanthene dyes such as fluoresceins, rhodamines and rhodols typically involve the condensation of two equivalents of resorcinol (for fluoresceins), aminophenol (for rhodamines) or a mixture of a resorcinol and an aminophenol (for rhodols) with a carbonyl-containing moiety such as a phthalic acid derivative or benzaldehyde. In the synthesis of the xanthene indicators of the invention, the desired resorcinol or aminophenol is condensed with the substituted crown ether, yielding either the reduced xanthene (where the crown ether contains an aldehyde) or the oxidized xanthene (where the crown ether contains a carboxylic acid or acyl halide) bound directly to the chelating moiety.

An oxidation step is typically required after condensation of a formyl-substituted crown ether with the fluorophore precursors. Optionally, the dihydro condensation product is isolated and subsequently oxidized with air or by standard chemical oxidants, such as chloranil. For some fluorophores, the oxidation reaction is enhanced by acidic reaction conditions. These mild oxidation reaction conditions tolerate a wide variety of substituents on the fluorophore and/or crown ether of the resulting indicators.

Unsymmetrical xanthene dyes are typically constructed by statistical methods, using a 1:1 mixture of the desired resorcinols or aminophenols in the condensation reaction, and purifying the desired product from the statistical mix of products using methods known in the art.

The synthesis of polyazaindacene dyes, particularly dipyrrometheneboron difluoride dyes, has been well documented (U.S. Patent Nos. 4,774,339; 5,187,288; 5,248,782; 5,274,113; 5,338,854 and 5,433,896). The procedure typically consists of an acid-catalyzed condensation of a benzaldehyde with a pyrrole that has a hydrogen at the 2-position, followed by *in situ* oxidation of the condensed intermediate by air, oxygen or a chemical oxidant such as 2,3-dichloro-5, 6-dicyano-1, 4-benzoquinone (DDQ). The condensation of two appropriately substituted pyrroles, each having a hydrogen at the 2-position, with a formyl-substituted crown ether, followed by *in situ* oxidation of the condensed intermediate and treatment with a boron trifluoride etherate yields the dipyrrometheneboron difluoride

indicators of the invention. Alternatively, the indicators are formed via the direct condensation of a carboxyl- or chlorocarbonyl-substituted crown ether with two equivalents of appropriately substituted pyrroles, which may be the same or different, provided each has a hydrogen at the 2-position. The latter procedure does not require oxidation.

5

Post-condensation modifications of both the crown ether and the fluorophore moiety are typically strictly analogous to known methods of indicator modification; for example, the reduction of nitro substituents to amino groups, the conversion of carboxy substituents to cyano groups, and the preparation of esters of carboxylic acids, including acetoxymethyl esters. Additionally, salts and counterions of the indicators of the invention are readily converted to other salts by treatment with ion-exchange resins, selective precipitation, and basification, as is well-known in the art.

10

Post-condensation modifications of xanthylium dyes are well known. For instance, the xanthenone portion of the dye can be halogenated by treatment with the appropriate halogenating agent, such as liquid bromine. Xanthenes containing unsaturated fused rings can be hydrogenated to the saturated derivatives.

15

The reduced and oxidized versions of the xanthene indicators are freely interconvertible by well-known oxidation or reduction reagents, including borohydrides, aluminum hydrides, hydrogen/catalyst, and dithionites. Care must be exercised to select an oxidation or reducing agent that is compatible with the crown ether chelator. A variety of oxidizing agents mediate the oxidation of dihydroxanthenes, including molecular oxygen in the presence or absence of a catalyst, nitric oxide, peroxynitrite, dichromate, triphenylcarbenium and chloranil. The dihydroxanthenes are also oxidized electrochemically, or by enzyme action, including horseradish peroxidase in combination with peroxides or by nitric oxide.

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Rather than condensing the DYE moiety precursors directly with a substituted crown ether, the preformed DYE moiety may be covalently bound to the crown ether via a conventional cross-linking reaction. A wide variety of chemically reactive or potentially chemically reactive and fluorescent fluorescein, rhodamine, rhodol, benzoxanthenes, dibenzoxanthene and other xanthene oxygen heterocycles that absorb maximally beyond about 490 nm are commercially available as described by Haugland, HANDBOOK OF FLUORESCENT PROBES AND RESEARCH PRODUCTS (*Supra*), as described above, or in other literature references. The nature of the bond that links the DYE moiety to the crown ether chelate appears to have an effect on the optical response of the DYE moiety to ion binding, sometimes a significant effect. Acceptability of the linking chemistry can be determined by

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titration of the resultant indicator with the ion of interest over the target range of response (as described in Example 71).

## **B. Method of Use**

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The crown ether compounds of the invention are useful for any application where it is desirable to complex a target metal ion. Selected crown ether compounds of the invention may be useful as ionophores, that is, they facilitate the transport of selected target ions across cell membranes. Where the crown ether compound is bound to a conjugated substance that is a polymeric matrix, such as a microparticle, or agarose, the compounds are useful for depleting a sample solution of a selected target ion, particularly where the polymeric matrix is used to pack a chromatography column. Other crown ether compounds (those bound to a DYE moiety) are useful as colorimetric or fluorescent indicators for a selected target ion.

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In order for a particular indicator of the present invention to be useful for detection purposes, it must exhibit a detectable change in spectral properties upon complexation of the desired metal ion (target ion) in the chelating moiety. Preferably the change in spectral properties is a change in fluorescence properties. More preferably, the instant indicators display an intensity increase or decrease in emission energy upon the complexation of the desired target ion.

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However, it should be appreciated that at least the compounds represented by Formula (II) do not require a DYE moiety. These compounds are useful for binding target ions resulting in a complex of the target ion and the present crown ether chelate compounds. Therefore, an additional aspect of the invention includes the compound of the invention further comprising a metal ion that is associated and/or complexed within the crown ether chelate portion of the compound. The metal ion is optionally a monocation (such as  $\text{Li}^+$ ,  $\text{Na}^+$ ,  $\text{K}^+$ ,  $\text{Rb}^+$ , or  $\text{Cs}^+$ ), a dication (such as  $\text{Ca}^{2+}$ ,  $\text{Zn}^{2+}$ , or  $\text{Mg}^{2+}$ ), or a polycation (such as  $\text{Tb}^{3+}$  or  $\text{Eu}^{3+}$ ). Preferably the crown ether chelate metal in complex comprises physiological relevant cations such as sodium, potassium, calcium and zinc.

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Accordingly, a method for binding target metal ions in a sample comprises the following steps:

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- a) contacting said sample with a crown ether chelate compound of the present invention; and,



- b) incubating said sample and said metal chelating compound for sufficient time to allow said compound to chelate said target metal ion whereby said metal ion is bound.

- 5 When the present compounds are used as indicators a DYE moiety is covalently attached to the crown ether chelate. The sample is illuminated with an appropriate wavelength whereby the target ion is detected. In such an assay the target ion can also be quantitated and monitored.
- 10 The specific indicator used in an assay or experiment is selected based on the desired affinity for the target ion as determined by the expected concentration range in the sample, the desired spectral properties, and the desired selectivity. Initially, the suitability of a material as an indicator of ion concentration is commonly tested by mixing a constant amount of the indicating reagent with a measured amount of the target ion under the
- 15 expected experimental conditions.

- Preferred indicators display a high selectivity, that is, they show a sufficient rejection of non-target ions. The interference of a non-target ion is tested by a comparable titration of the indicator with that ion. Although preferred target ions for most indicators of the present
- 20 invention are  $\text{Na}^+$  and  $\text{K}^+$ , any ion that yields a detectable change in absorption wavelengths, emission wavelengths, fluorescence lifetimes or other measurable optical property over the concentration range of interest is potentially measured using one of the indicators of this invention. Modifications to the electronic structure of the crown ether or indicator to produce an indicator having the appropriate combination of binding affinity, ion selectivity and
- 25 spectral response for a wide variety of metal ions.

- In one embodiment of the invention, the target ions for the indicators of the present invention are selected from  $\text{Li}^+$ ,  $\text{Na}^+$ ,  $\text{K}^+$ ,  $\text{Cs}^+$ ,  $\text{Ca}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Rb}^+$ ,  $\text{Tb}^{3+}$  or  $\text{Eu}^{3+}$ . In another embodiment of the invention, the target ions are selected from  $\text{Li}^+$ ,  $\text{Na}^+$ ,  $\text{K}^+$ ,  $\text{Ca}^{2+}$ ,  $\text{Zn}^{2+}$ , and
- 30  $\text{Mg}^{2+}$ . Additional target ions for selected embodiments of the present indicators also include  $\text{Mn}^{2+}$ ,  $\text{Fe}^{2+}$ ,  $\text{Fe}^{3+}$ ,  $\text{Co}^{2+}$ ,  $\text{Ni}^{2+}$ ,  $\text{Cu}^{2+}$ ,  $\text{Cu}^+$ ,  $\text{Zn}^{2+}$ ,  $\text{Al}^{3+}$ ,  $\text{Cd}^{2+}$ ,  $\text{Ag}^+$ ,  $\text{Au}^+$ ,  $\text{Tl}^+$ ,  $\text{Pd}^{2+}$ ,  $\text{Hg}^{2+}$ ,  $\text{Hg}^+$ ,  $\text{Sn}^{2+}$ ,  $\text{Pb}^{2+}$ ,  $\text{Sr}^{2+}$ ,  $\text{Ba}^{2+}$ ,  $\text{Mo}^{3+}$ ,  $\text{Ga}^{3+}$ ,  $\text{In}^{3+}$ ,  $\text{La}^{3+}$ ,  $\text{Eu}^{3+}$ ,  $\text{Tb}^{3+}$ ,  $\text{Dy}^{3+}$ ,  $\text{Ru}^{3+}$ ,  $\text{Sc}^{3+}$ ,  $\text{As}^{3+}$ ,  $\text{Sb}^{3+}$ ,  $\text{Cr}^{3+}$ ,  $\text{Bi}^{3+}$ ,  $\text{Ce}^{3+}$ ,  $\text{Ce}^{4+}$ ,  $\text{Pd}^{2+}$ ,  $\text{Pt}^{2+}$  and  $\text{Pt}^{4+}$  ions. In yet another embodiment of the invention, the target ions of the instant indicators are  $\text{Fe}^{2+}$ ,  $\text{Fe}^{3+}$ ,  $\text{Co}^{2+}$ ,  $\text{Ni}^{2+}$ ,  $\text{Cu}^{2+}$ ,  $\text{Cu}^+$ ,  $\text{Zn}^{2+}$ ,  $\text{Al}^{3+}$ ,  $\text{Cd}^{2+}$ ,  $\text{Hg}^{2+}$ ,  $\text{Pb}^{2+}$ ,
- 35  $\text{Ba}^{2+}$ ,  $\text{La}^{3+}$ ,  $\text{Tb}^{3+}$  and  $\text{Cr}^{3+}$  ions. In yet another embodiment, the target ions are selected from the group consisting of  $\text{Fe}^{3+}$ ,  $\text{Ni}^{2+}$ ,  $\text{Cu}^{2+}$ ,  $\text{Cu}^+$ ,  $\text{Hg}^{2+}$ , or  $\text{Pb}^{2+}$ .

The indicator is generally prepared for use as a detection reagent by dissolving the indicator in solution at a concentration that is optimal for detection of the indicator at the expected concentration of the target ion. Modifications that are designed to enhance permeability of the indicator through the membranes of living cells, such as acetoxymethyl esters and acetates, may require the indicator to be predissolved in an organic solvent such as dimethylsulfoxide (DMSO) before addition to a cell suspension, where the indicators then readily enter the cells. Intracellular enzymes cleave the esters to the more polar acids and phenols that are then well retained inside the cells. For applications where permeability of cell-membranes is required, the indicators of the invention are typically substituted by only one fluorophore.

Therefore, a method for binding and detecting target ions in a live cell comprises the following steps:

- a) contacting a sample of live cells with a crown ether chelate compound of the present invention wherein said compound comprises a DYE moiety and at least one lipophilic group;
- b) incubating said sample and said crown ether chelate compound for sufficient time to allow said compound to chelate said target metal ion; and,
- c) illuminate said sample with an appropriate wavelength whereby said target ion is detected in a live cell.

Typically, the lipophilic group is an AM or acetate ester that is directly attached to the DYE moiety of the crown ether chelate compound.

A specific indicator of the present invention is useful for the detection and/or quantification of a desired target ion, when the binding of the target ion in the metal ion-binding moiety of the indicator results in a detectable change in spectral properties. Preferably, the change in spectral properties is a detectable fluorescence response.

A preferred indicator for a specific target ion is an indicator that shows at least a two-fold change in net fluorescence emission intensity (either higher or lower), or a 1 nanosecond difference in fluorescence lifetime (either shorter or longer), preferably a five-fold or greater change in net fluorescence emission intensity or a 100% change in fluorescence lifetime in response to the target ion. Alternatively, an indicator that exhibits a shift in excitation or emission wavelength of at least 10 nm (either to shorter or longer wavelength) is also preferred, more preferably exhibiting a shift of 25 nm or greater.

The optical response of the indicating reagent is determined by changes in absorbance or fluorescence, preferably fluorescence. If absorbance measurements are used to determine ion concentrations, then it is usually optimal to adjust the optical density of the indicator in the sample over the range of analyte concentration to a value of approximately 0.02 to 2.5 (most preferably 0.1 to 1). For fluorescence measurements, the concentration of the indicator will depend mostly on the sensitivity of the equipment used for its detection.

If the optical response of the indicator will be determined using fluorescence measurements, samples are typically stained with indicator concentrations of  $10^{-9}$  M to  $10^{-2}$  M. The most useful range of analyte concentration is about one log unit above and below the dissociation constant of the ion-indicator complex. This dissociation constant is determined by titration of the indicator with a known concentration of the target ion, usually over the range of virtually zero concentration to approximately 100 millimolar of the target ion, depending on which ion is to be measured and which indicator is being used. The dissociation constant may be affected by the presence of other ions, particularly ions that have similar ionic radii and charge. It may also be affected by other conditions such as ionic strength, pH, temperature, viscosity, presence of organic solvents and incorporation of the sensor in a membrane or polymeric matrix, or conjugation or binding of the sensor to a protein or other biological molecule. Any or all of these effects need to be taken into account when calibrating an indicator.

The indicator is combined with a sample in a way that will facilitate detection of the target ion concentration in the sample. The sample is generally a representative cell population, fluid or liquid suspension that is known or suspected to contain the target ion. Representative samples include intracellular fluids such as in blood cells, cultured cells, muscle tissue, neurons and the like; extracellular fluids in areas immediately outside of cells; in vesicles; in vascular tissue of plants and animals; in biological fluids such as blood, saliva, and urine; in biological fermentation media; in environmental samples such as water, soil, waste water and sea water; in industrial samples such as pharmaceuticals, foodstuffs and beverages; and in chemical reactors. Detection and quantitation of the target ion in a sample can help characterize the identity of an unknown sample, or facilitate quality control of a sample of known origin.

In one embodiment of the invention, the sample contains cells, and the indicator is combined with the sample in such a way that the indicator is present within the sample cells. By selection of the appropriate chelating moiety, fluorophore, and the substituents thereon, indicators are prepared that will selectively localize in desired organelles, and provide

measurements of the target ion in those organelles. Conjugates of the indicators of the invention with organelle-targeting peptides are used to localize the indicator to the selected organelle, facilitating measurement of target ion presence or concentration within the organelle (as described in U.S. Patent No. 5,773,227). Alternatively, selection of a lipophilic fluorophore, or a fluorophore having predominantly lipophilic substituents will result in localization in lipophilic environments in the cell, such as cell membranes. Selection of cationic indicators will typically result in localization of the indicator in mitochondria.

In another aspect of the invention, a composition of matter comprises any of the compounds described above, and optionally includes a metal ion. In one embodiment, the compounds of the invention, in any of the embodiments described above, are associated, either covalently or noncovalently, with a surface such as a microfluidic chip, a silicon chip, a microscope slide, a microplate well, or another solid matrix, and is combined with the sample of interest as it flows over the surface. The detectable optical response is therefore detected on the matrix surface itself, typically by use of an instrumental. This embodiment of the invention is particularly suited to high-throughput screening using automated methods.

Quantification of target ion levels in samples is typically accomplished using the indicators of the present invention by methods known in the art. For example, the ratiometric measurement of ion concentration provides accurate measurement of ion concentrations by the treatment of the fluorescence data as the ratio of excitation or fluorescence intensities at two wavelengths, rather than the absolute intensity at a single wavelength. Using the ratio method, a number of variables that may perturb the ion concentration measurements are eliminated. In particular, ion-dependent factors that affect the signal intensity, such as nonuniform intracellular dye concentrations, probe leakage, dye bleaching and cell thickness, are canceled in the ratio measurements, since these parameters have a similar effect on intensities at both wavelengths. While the ratio method can be used to determine concentrations using observation of either the excitation spectra of the indicator, the emission spectra of the indicator, or both, in the case of the indicators of the present invention, the shift in excitation energy upon binding metal ions makes observation of the excitation spectrum a more useful technique. In either case, to achieve maximal utility, the indicator must be calibrated (to compensate for variance in the dissociation constant of the indicator due to ionic strength, viscosity, or other conditions within the sample). To calibrate the indicator, ionophores such as A-23187, gramicidin, valinomycin, or ionomycin are used. Non-ratiometric analysis can also be accomplished by calibration with a second fluorescent dye present in the sample.

The optical response of the indicator to the ion can be detected by various means that include measuring absorbance or fluorescence changes with an instrument, visually, or by use of a fluorescence sensing device. Several examples of fluorescence sensing devices are known, such as fluorometers, fluorescence microscopes, laser scanners, flow  
5 cytometers, and microfluidic devices, as well as by cameras and other imaging equipment. These measurements may be made remotely by incorporation of the fluorescent ion sensor as part of a fiber optic probe. The indicator is covalently attached to the fiber optic probe material, typically glass or functionalized glass (e.g., aminopropyl glass) or the indicator is attached to the fiber optic probe via an intermediate polymer, such as polyacrylamide. The  
10 indicator solution is alternatively incorporated non-covalently within a fiber optic probe, as long as there is a means whereby the target ion can come into contact with the indicator solution.

### **C. Kits of the Invention**

15 Due to the advantageous properties and the simplicity of use of the instant crown ether compounds, they are particularly useful in the formulation of a kit for the complexation, detection, quantification or monitoring of selected target ions, comprising one or more compounds or compositions of the invention in any of the embodiments described above  
20 (optionally in a stock solution), instructions for the use of the crown ether compound to complex or detect a desired target ion, and optionally comprising additional components. In one aspect, the compounds of the invention are associated with a surface, such as a chip, microplate well, or other solid matrix, and the sample of interest flows over the surface. The detectable optical response is therefore detected on the matrix surface itself.

25 Therefore a kit of the present invention for binding a target metal ion in a sample comprises a compound represented by Formula (I) or Formula (II) and comprising one or more components selected from the group consisting of a calibration standard of a metal ion, an ionophore, a fluorescent standard, an aqueous buffer solution and an organic solvent.

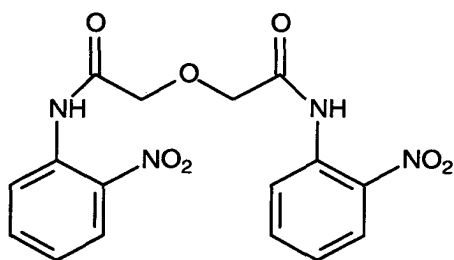
30 The additional kit components may be selected from, without limitation, calibration standards of a target ion, ionophores, fluorescence standards, aqueous buffers, and organic solvents. The additional kit components are present as pure compositions, or as aqueous solutions that incorporate one or more additional kit components. Any or all of the kit components  
35 optionally further comprise buffers.

The examples below are given so as to illustrate the practice of this invention. They are not intended to limit or define the entire scope of this invention.

## 5 EXAMPLES

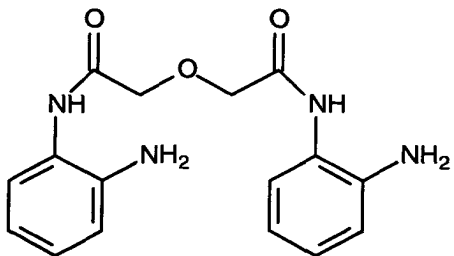
### Example 1. Preparation of a tetraaza-crown ether.

2-Nitroaniline is acylated with diglycolyl chloride in anhydrous THF, using 0.5 equivalents of diglycolyl chloride in dilute solution. A 0.1 M THF solution of diglycolyl chloride is slowly  
10 added to a 0.1 M THF solution of 2-nitroaniline, containing two equivalents of triethylamine. After TLC indicates consumption of all starting 2-nitroaniline, volatiles are removed in vacuo and the residue is partitioned between water and ethyl acetate. The ethyl acetate layer is concentrated to give Compound 1.



15 Compound 1

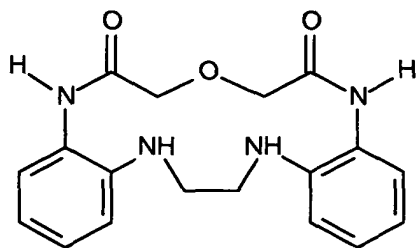
Compound 1 is dissolved in methanol and treated with 20% by weight of 10% palladium on carbon. The resulting mixture is shaken under 40 psi hydrogen until analysis by thin layer chromatography (TLC) indicates the reaction is complete. The reaction mixture is filtered  
20 and the bis-aniline 2 is isolated by concentration *in vacuo* of the filtrate.



Compound 2

Compound 2 is dissolved in DMF to achieve a 0.1 M solution and then treated with 0.9  
25 equivalents of 1,2-dibromoethane, two equivalents of diisopropylethylamine, and catalytic sodium iodide. The resulting solution is heated until most of compound 2 is consumed, as

judged by TLC analysis. The volatiles are removed *in vacuo*, and the residue is partitioned between ethyl acetate and water. The ethyl acetate layer is concentrated *in vacuo* to give compound 3.

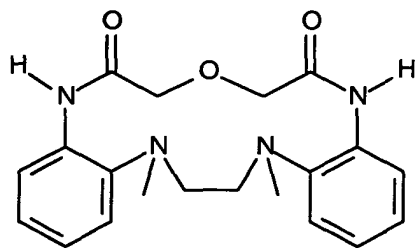


5

Compound 3

Compound 3 is dissolved in acetonitrile to achieve a 0.5 M concentration. Ten equivalents of formaldehyde, as a 37% aqueous solution, are added followed by 4 equivalents of sodium cyanoborohydride. The pH of the reaction mixture is adjusted to 7 with acetic acid, and the resulting solution is stirred until TLC analysis indicates consumption of compound 3 and formation of compound 4. Volatiles are removed *in vacuo*, and the residue is partitioned between ethyl acetate and water. The ethyl acetate layer is concentrated *in vacuo* to give compound 4.

10

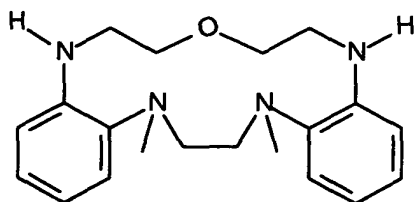


15

Compound 4

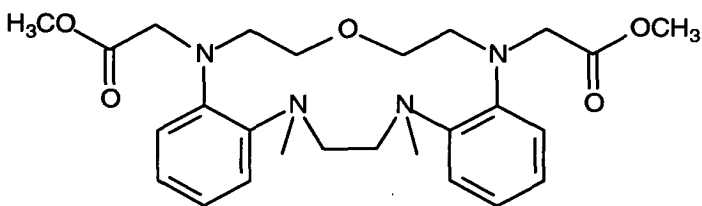
Compound 4 is dissolved in THF at 1 M concentration, and treated with 10 equivalents of borane-THF (1M). The resulting solution is heated at reflux until compound 4 is consumed and compound 5 is formed, as judged by TLC analysis. The reaction is quenched by addition of 5% aqueous sodium hydroxide, followed by extraction with ethyl acetate. The ethyl acetate layer is concentrated *in vacuo* to give compound 5.

20



Compound 5

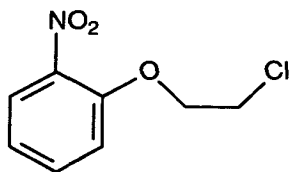
Compound 5 is dissolved in DMF at 0.5 M concentration. Ten equivalents of methyl bromoacetate and 4 equivalents of diisopropylethylamine are added, and the resulting solution is stirred at 100 degrees centigrade until TLC analysis indicates consumption of compound 5 and formation of compound 6. Volatiles are removed in vacuo, and the residue is partitioned between water and ethyl acetate. The ethyl acetate layer is concentrated in vacuo to give compound 6, which can be further purified by flash chromatography on silica gel using methanol in chloroform as eluant.



Compound 6

**Example 2.** Preparation of (2'-Nitrophenoxy)-2-chloroethane (7).

A suspension of 2-nitrophenol (50.0 g, 0.36 mol), 1-bromo-2-chloroethane (45 mL, 0.54 mol), and  $K_2CO_3$  (100.0 g, 0.72 mol) in DMF (200 mL) was stirred at 90°C for 2 h, cooled to room temperature, poured into ice-water, filtered, washed with  $H_2O$  and dried to give Compound 7, 70.3 g (96%) as a yellow solid.

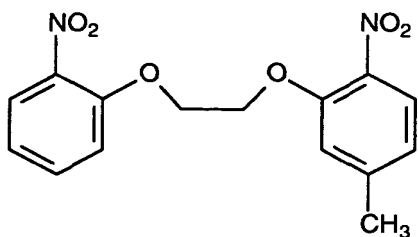


Compound 7

**Example 3:** Preparation of 1-(5'-methyl-2'-nitrophenoxy)-2-(2''-nitrophenoxy)ethane (8).

A suspension of Compound 7 (60.45 g, 0.30 mol), 5-methyl-2-nitrophenol (50.49 g, 0.33 mol), and  $K_2CO_3$  (82.80 g, 0.60 mol) in DMF (450 mL) was stirred at 130 °C for 18 h, cooled to room temperature, poured into ice-water, filtered, washed with  $H_2O$  and dried to give Compound 8, 89.50 g (93.5%) as an orange solid.

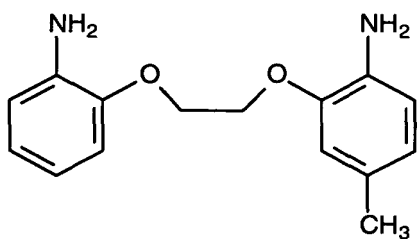




Compound 8

**Example 4.** Preparation of 1-(2'-amino-5'-methylphenoxy)-2-(2''-aminophenoxy)ethane (**9**).

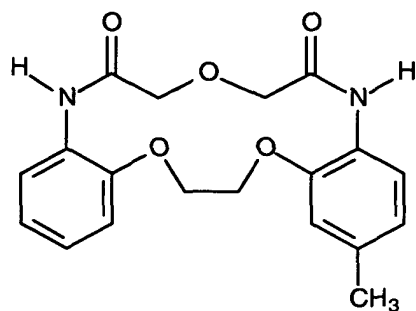
- 5 Compound **8** (24.0 g, 75 mmol) was hydrogenated over 10% Pd/C (1.5 g) in DMF (250 mL) at 40 psi for 18 h. The mixture was filtered from catalyst through a CELITE pad on a fritted glass filter. The filtrate was evaporated and ether (50 mL) was added. The product was filtered, washed with ether and dried to give compound **9**, 18.0 g (95%) as a brown solid.



Compound 9

**Example 5.** Preparation of Compound **10**.

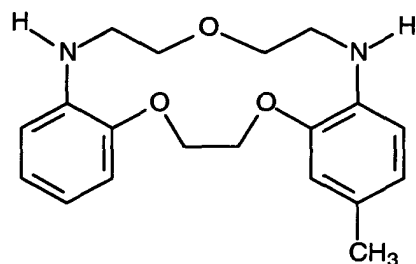
- 15 To a stirred mixture of diamine Compound **9** (13.10 g, 50 mmol) and Et<sub>3</sub>N (20.8 mL, 150 mmol) in dry THF (2.5 L) was slowly added over 12 h a solution of diglycolyl dichloride (6.3 mL, 55 mmol) in dry THF (0.5 L). The reaction mixture was stirred for 6 h, filtered from the precipitated hydrochloride, and washed with THF. The combined organic filtrate was evaporated. The residue was dissolved in CHCl<sub>3</sub> (800 mL) and washed successively with 0.5 M HCl, H<sub>2</sub>O, saturated NaHCO<sub>3</sub> then saturated NaCl. The organic layer was dried over
- 20 MgSO<sub>4</sub> and evaporated. Ether (100 mL) was added and the precipitated product was filtered, washed with ether and dried to give Compound **10**, 16.0 g (90%) as an off-white solid.



Compound **10**

**Example 6.** Preparation of Compound **11**

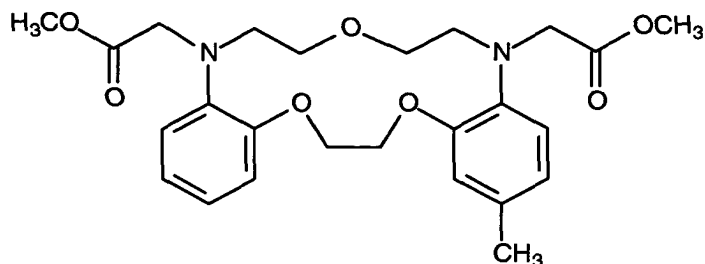
- 5 To a suspension of diamide Compound **10** (15.60 g, 44 mmol) in dry THF (400 mL) was added a 1 M solution of  $\text{BH}_3$ -THF complex in dry THF (390 mL, 390 mmol). The mixture was stirred at 70 °C for 16 h. Dry methanol (300 mL) was added dropwise over 1 h into a boiling mixture with strong gas evaluation. The mixture was heated at reflux for 2 h, cooled to room temperature, evaporated and co-evaporated with MeOH to destroy the borane complex.
- 10 Ether (100 mL) was added. The solid product was filtered, washed with ether and dried to give Compound **11**, 12.3 g (85%) as an off-white solid.



Compound **11**

15 **Example 7.** Preparation of Compound **12**.

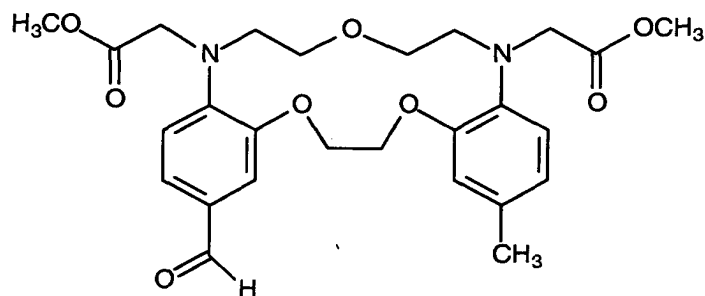
- A mixture of diamine **11** (6.56 g, 20 mmol), methyl bromoacetate (38 mL, 400 mmol), diisopropylethylamine (DIPEA) (104 mL, 600 mmol), and NaI (1.50 g, 10 mmol) in MeCN (300 mL) was heated at reflux for 70 h. After cooling to room temperature, the MeCN was
- 20 evaporated. The residue was dissolved in  $\text{CHCl}_3$ , washed with 1% AcOH then  $\text{H}_2\text{O}$ , dried and evaporated. The residual oily product was cooled to 0 °C and washed with cold hexane to remove most of the alkylating reagent. The crude product, compound **12**, was used immediately in the next step without further purification.



Compound 12

**Example 8.** Preparation of Compound 13

- 5 To a stirred solution of the Vilsmeier reagent prepared from  $\text{POCl}_3$  (4.65 mL, 50 mmol) and 30 mL DMF was added over 5 min a solution of Compound 12 (4.8 g, 10 mmol) in DMF (20 mL). The mixture was stirred for 16 h, cooled by the addition of ice and neutralized with saturated  $\text{K}_2\text{CO}_3$  to pH 7-8. The suspension was extracted with  $\text{CHCl}_3$ , washed successively with 0.1 M HCl, saturated NaCl, saturated  $\text{NaHCO}_3$  and saturated NaCl and then
- 10 evaporated. The residue was dissolved in  $\text{CHCl}_3$  and chromatographed on  $\text{SiO}_2$  with an EtOAc gradient in hexanes (20% to 30%) to give aldehyde 13, 2.60 g (52%) as colorless crystals.

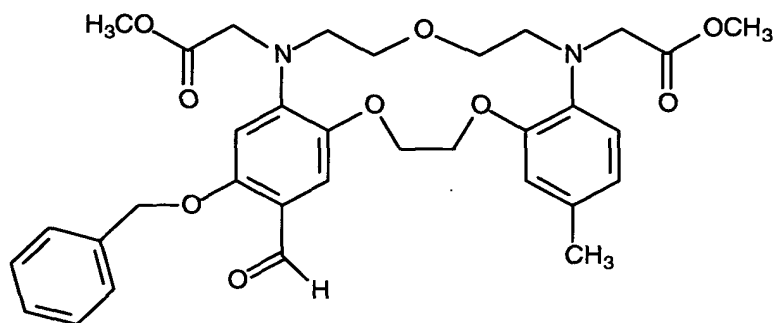


Compound 13

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**Example 9.** Preparation of Compound 14:

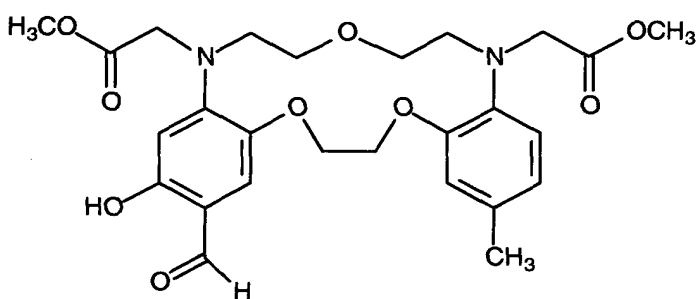
- Compound 14 was prepared analogously with Compound 13, starting with 5-methyl-2-nitrophenol in place of 2-nitrophenol, and utilizing 4-benzyloxy-2-nitrophenol in place of 5-methyl-2-nitrophenol.
- 20



Compound 14

**Example 10. Preparation of Compound 15**

- 5 Compound 14 (3.06 g, 5.0 mmol) was hydrogenated over 10% Pd/C (0.50 g) in AcOH (100 mL) at 50 psi for 5 h. The mixture was filtered through a CELITE pad on a fritted glass filter. The filtrate was evaporated, ether (50 mL) was added, the product was filtered, washed with ether and dried to give Compound 15, 2.13 g (82%) as a colorless solid.



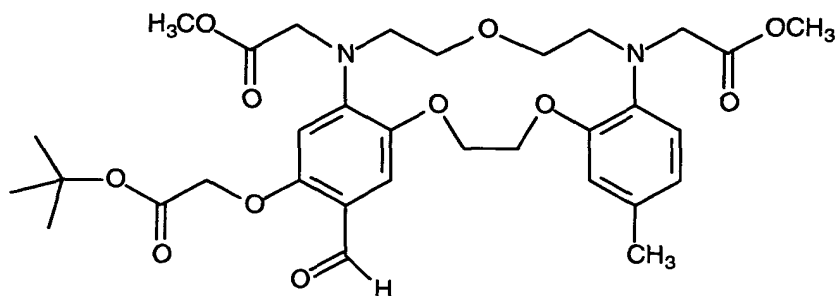
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Compound 15

**Example 11. Preparation of Compound 16**

- 15 To a suspension of compound 15 (0.800 g, 1.55 mmol),  $K_2CO_3$  (1.04 g, 7.5 mmol) and NaI (0.03 g, 0.2 mmol, catalyst) in DMF (10 mL) was added dropwise *t*-butyl bromoacetate (0.45 mL, 3.0 mmol). The mixture was stirred for 16 h, diluted with  $H_2O$ , extracted with  $CHCl_3$ , dried over  $MgSO_4$  and evaporated. The residue was dissolved in hexane: EtOAc (1:1) and chromatographed on silica gel with 40% EtOAc in hexanes to give aldehyde 16, 0.875 g (90%) as colorless crystals.

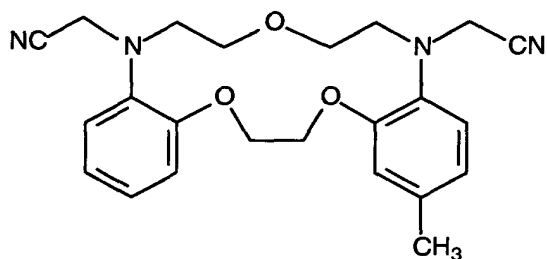
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Compound 16

**Example 12. Preparation of Compound 17**

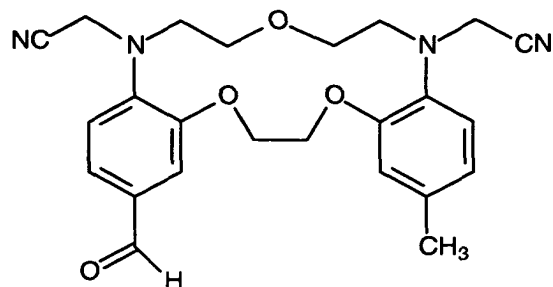
- 5 A stirred mixture of diamine **11** (3.38 g, 10 mmol), bromoacetonitrile (14 mL, 200 mmol), DIPEA (26 mL, 150 mmol), NaI (1.5 g, 10 mmol; catalyst) in MeCN (150 mL) was heated at reflux for 70 h. After cooling to room temperature the MeCN was evaporated. The residue was dissolved in CHCl<sub>3</sub> (800 mL), washed with H<sub>2</sub>O, dried over MgSO<sub>4</sub> and evaporated. The residue was chromatographed on SiO<sub>2</sub> with 40% EtOAc in hexanes to give Compound **17**,  
 10 3.56 g (90%) as a colorless solid.



Compound 17

**Example 13. Preparation of Compound 18**

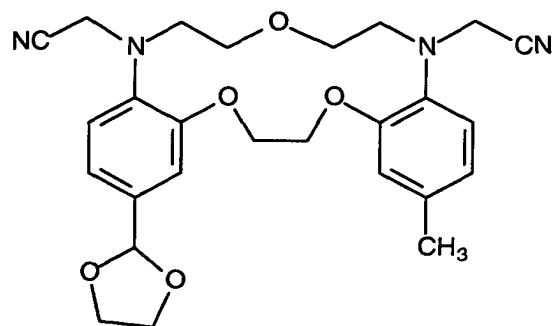
- 15 To a stirred solution of the Vilsmeier reagent prepared from POCl<sub>3</sub> (9.5 mL, 102 mmol) and 60 mL DMF was added over 5 min to a solution of Compound **17**. The mixture was stirred for 100 h at 40 °C, cooled by the addition of ice and neutralized with saturated K<sub>2</sub>CO<sub>3</sub> to pH 8. The suspension was extracted with CHCl<sub>3</sub>, washed with H<sub>2</sub>O, dried over MgSO<sub>4</sub> and  
 20 evaporated. The residue was chromatographed over SiO<sub>2</sub> using CHCl<sub>3</sub> as eluant to give compound **18**, 3.12 g (73%) as colorless crystals.



Compound 18

**Example 14. Preparation of Compound 19**

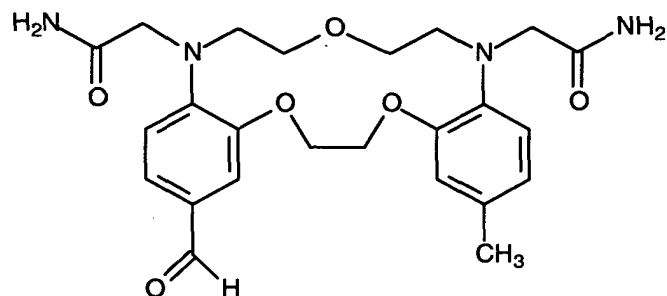
- 5 A mixture of aldehyde 18 (3.08 g, 7.3 mmol), ethylene glycol (8 mL, 140 mmol), *p*-toluenesulphonic acid (0.20 g, catalyst) and benzene (80 mL) was refluxed with a Dean-Stark trap for 4 h. After cooling to room temperature the mixture was evaporated. The residue was dissolved in  $\text{CHCl}_3$  (300 mL), washed with saturated  $\text{NaHCO}_3$ , dried over  $\text{MgSO}_4$ , and evaporated to give compound 19, 2.82 g (85%) as an orange solid. Compound
- 10 19 was pure on TLC and used in the next step without additional purification.



Compound 19

**Example 15. Preparation of Compound 20**

- 15 To a solution of Compound 19 (0.200 g, 0.44 mmol) in MeOH (25 mL) and 1 M NaOH (15 mL, 15 mmol) was added over 5 min 30%  $\text{H}_2\text{O}_2$  (5 mL, catalyst). The mixture was stirred for 1 h, then acidified with 1 M HCl to pH 1. The acidified mixture was stirred for 30 min, diluted with 3 M NaOAc (150 mL), extracted with  $\text{CHCl}_3$ , washed with saturated  $\text{NaHCO}_3$ , dried over
- 20  $\text{MgSO}_4$  and evaporated to give aldehyde 20, 0.191 g (95%) as a colorless crystalline material, pure by TLC.

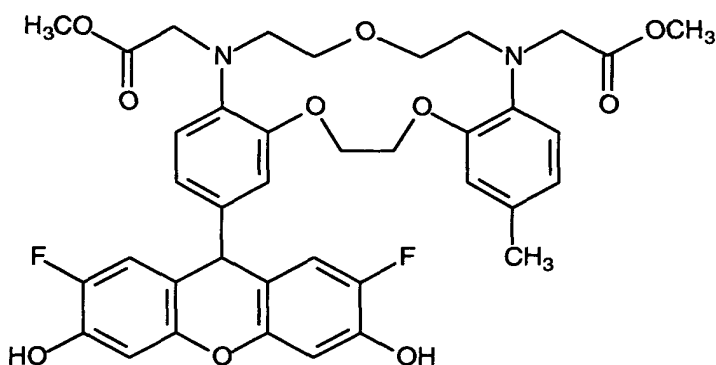


Compound 20

**Example 16. Preparation of Compound 21**

- 5 A mixture of aldehyde **13** (1.150 g, 2.3 mmol) and 4-fluororesorcinol (0.670 g, 5.2 mmol) in  $\text{MeSO}_3\text{H}$  (25 mL) was stirred overnight and then poured into 3 M NaOAc (300 mL). The precipitated solid was filtered, washed with  $\text{H}_2\text{O}$  and dried to give compound **21**, 1.605 g (97%) as an off-white solid. Compound **21** was unstable to oxidation and was used in the next step without additional purification.

10

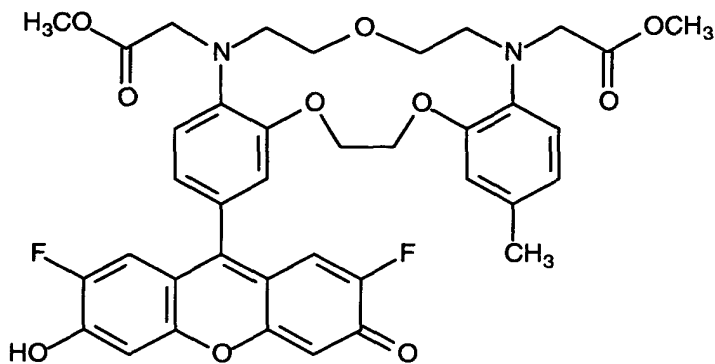


Compound 21

**Example 17. Preparation of Compound 22**

- 15 A mixture of Compound **21** (150 mg, 0.21 mmol) and chloranil (246 mg, 1.0 mmol) in MeOH (5 mL) and  $\text{CHCl}_3$  (5 mL) was refluxed for 50 h, then cooled to room temperature, filtered from excess oxidizer, and evaporated. The residue was purified by preparative TLC on two  $\text{SiO}_2$  plates using  $\text{CHCl}_3$ : MeOH:AcOH (20:2:1) to give compound **22**, 44 mg (20%) as a red-brown solid.

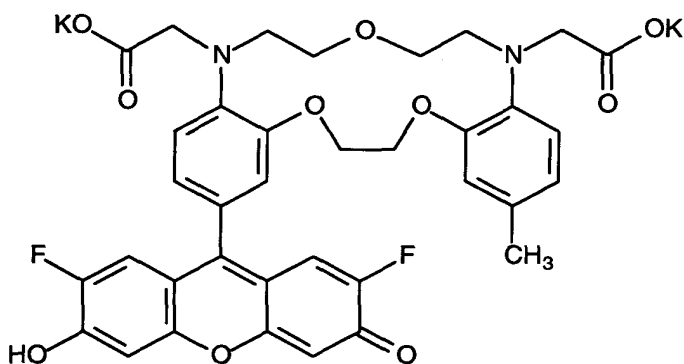
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Compound **22**

**Example 18. Preparation of Compound 23**

- 5 A mixture of Compound **22** (30 mg, 0.04 mmol) in MeOH (5 mL) and 1 M KOH (1 mL, 1.0 mmol) was stirred for 16 h, then neutralized with 1 M HCl to pH 7, and evaporated to dryness. The residue was purified by preparative TLC on two SiO<sub>2</sub> plates using CHCl<sub>3</sub>: MeOH:AcOH (13:3:1) to give compound **23**, 21 mg (74%) as a brown solid.



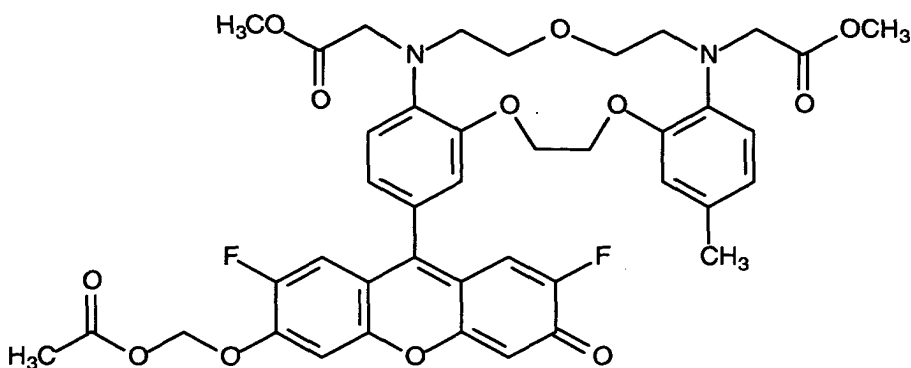
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Compound **23**

**Example 19. Preparation of Compound 24**

- 15 A mixture of Compound **22** (78 mg, 0.1 mmol), bromomethyl acetate (0.05 mL, 0.5 mmol), and DIPEA (0.175 mL, 1 mmol) in DMF (1 mL) was stirred for 2 h, then poured into 1% AcOH (200 mL). The suspension was extracted with CHCl<sub>3</sub>, washed with H<sub>2</sub>O, filtered and evaporated. The residue was purified by preparative TLC on two SiO<sub>2</sub> plates using 5% MeOH in CHCl<sub>3</sub> to give compound **24**, 25 mg (32%) as a brown solid.



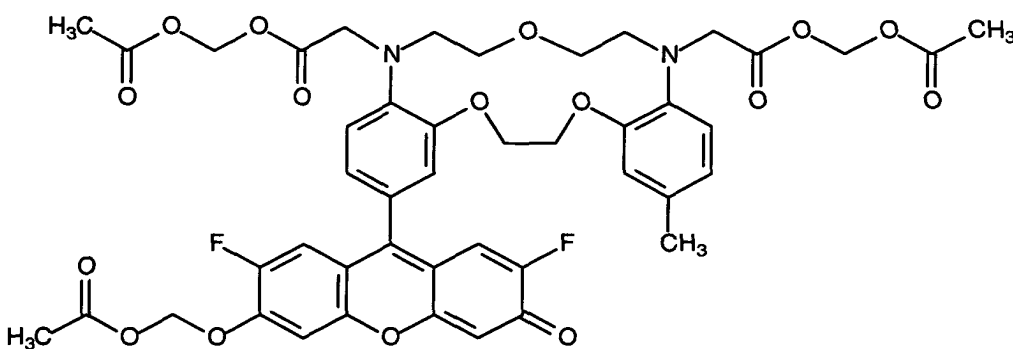


Compound 24

**Example 20. Preparation of Compound 25**

- 5 A mixture of Compound 23 (69 mg, 0.1 mmol), bromomethyl acetate (0.05 mL, 0.5 mmol), and DIPEA (0.175 mL, 1 mmol) in DMF (1 mL) was stirred for 2 h, then poured into 1% AcOH (200 mL). The suspension was extracted with CHCl<sub>3</sub>, washed with H<sub>2</sub>O, filtered and evaporated. The residue was purified by preparative TLC on two SiO<sub>2</sub> plates using 5% MeOH in CHCl<sub>3</sub> to give Compound 25 as a brown solid.

10



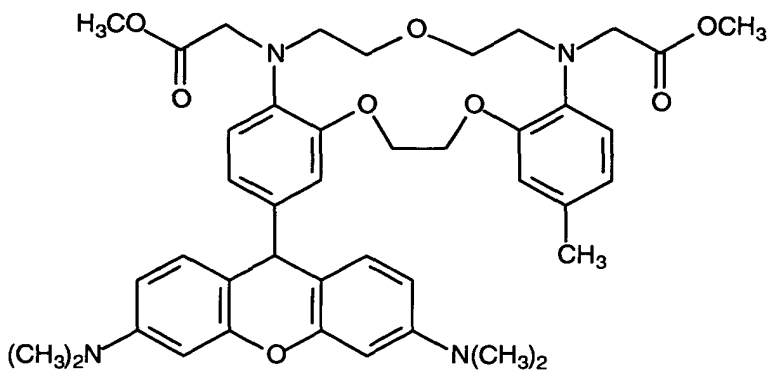
Compound 25

**Example 21. Preparation of Compound 26**

15

A mixture of aldehyde 13 (250 mg, 0.5 mmol), 3-dimethylaminophenol (157 mg, 1.1 mmol), and p-toluenesulphonic acid (10 mg, catalyst) in propionic acid (5 mL) was stirred overnight at 60 °C, then cooled to room temperature and poured into 3 M NaOAc (100 mL). The precipitated solid was filtered, washed with H<sub>2</sub>O and dried to give compound 26, 360 mg (97%) as a rose-colored solid. Compound 26 was unstable to oxidation and was used in subsequent reactions without additional purification.

20

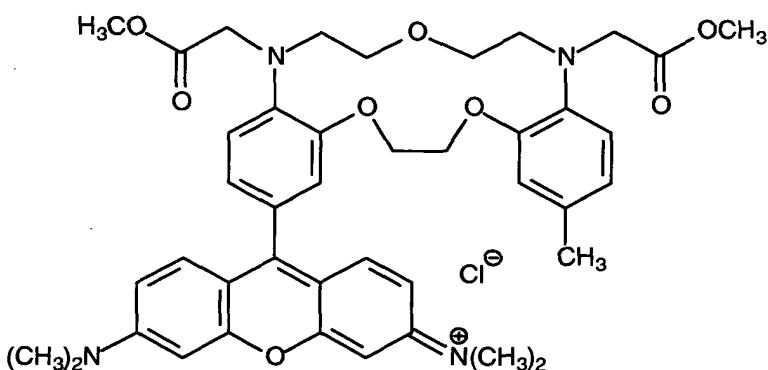


Compound 26

**Example 22. Preparation of Compound 27**

- 5 A mixture of Compound 26 (660 mg, 0.81 mmol) and chloranil (400 mg, 1.62 mmol) in MeOH (25 mL) and  $\text{CHCl}_3$  (25 mL) was stirred for 2 h, filtered from excess oxidizer, and evaporated. The residue was purified by chromatography on  $\text{SiO}_2$  using 9% MeOH and 1% AcOH in  $\text{CHCl}_3$  as eluant to give the crude product, which was again chromatographed on  $\text{SiO}_2$  using the same eluant to give Compound 27, 112 mg (19%) as crimson solid.

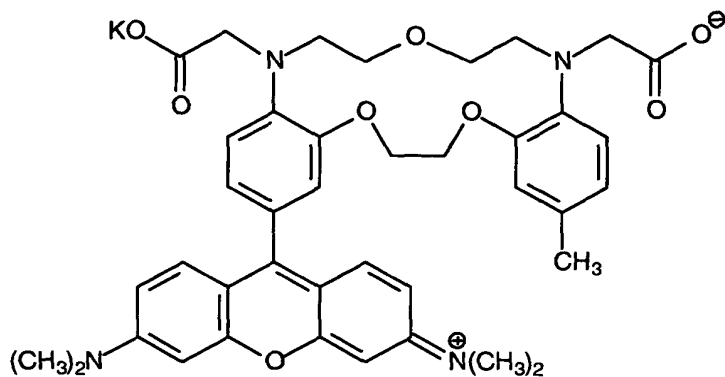
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Compound 27

**Example 23. Preparation of Compound 28**

- 15 A mixture of Compound 27 (40 mg, 0.05 mmol) in MeOH (2 mL) and 1 M KOH (1 mL, 1.0 mmol) was stirred for 16 h and then evaporated to dryness. The residue was purified by chromatography on a SEPHADEX LH-20 column using  $\text{H}_2\text{O}$  as eluant to give salt 28, 8 mg (21%) as red flakes after lyophilization.

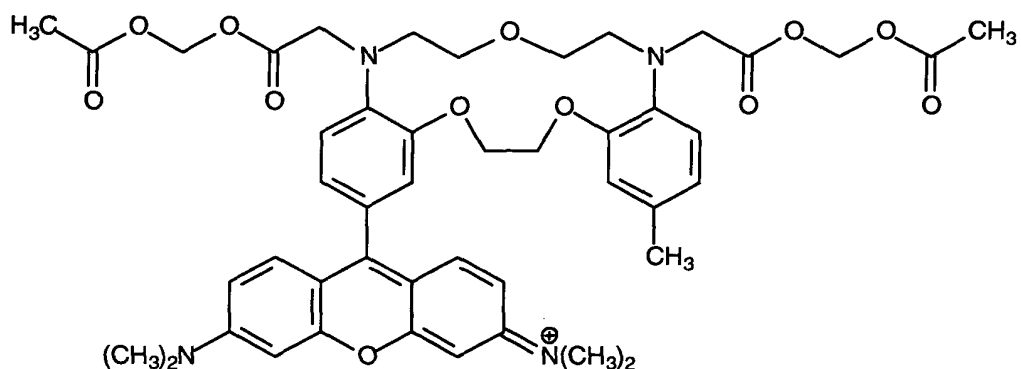


Compound **28**

**Example 24.** Preparation of Compound **29**

- 5 A mixture of Compound **28** (75 mg, 0.1 mmol), bromomethyl acetate (0.1 mL, 1.0 mmol), and DIPEA (0.35 mL, 2 mmol) in DMF (2 mL) was stirred for 2 h, then poured into 1% AcOH (200 mL). The suspension is extracted with CHCl<sub>3</sub>, washed with H<sub>2</sub>O, filtered and evaporated. The residue was purified by preparative TLC on two SiO<sub>2</sub> plates using 5% MeOH in CHCl<sub>3</sub> to give Compound **29** as a red solid.

10

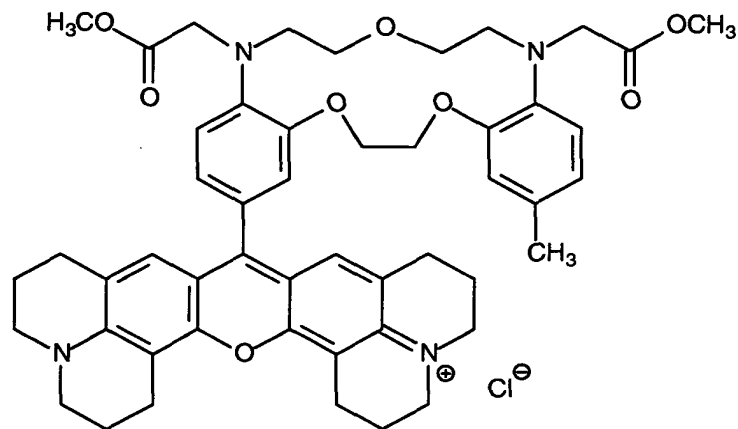


Compound **29**

**Example 25.** Preparation of Compound **30**

15

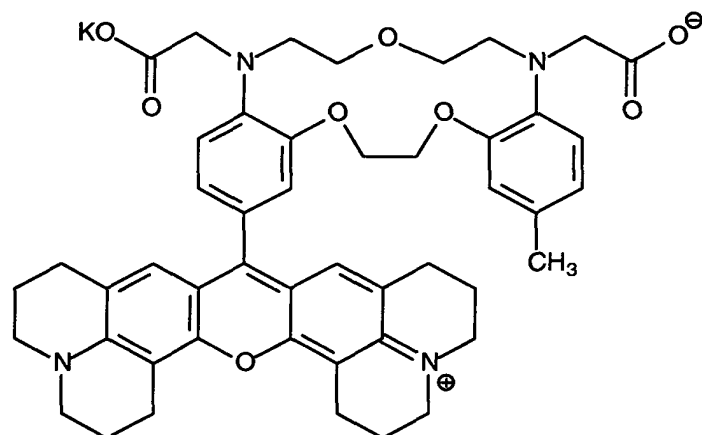
Compound **30** was prepared analogously to Compound **27**, using 8-hydroxyjulolidine rather than *m*-dimethylaminophenol.



Compound 30

**Example 26.** Preparation of Compound 31

- 5 Compound 31 was prepared from Compound 27 using the procedure for preparing Compound 28 from Compound 27.

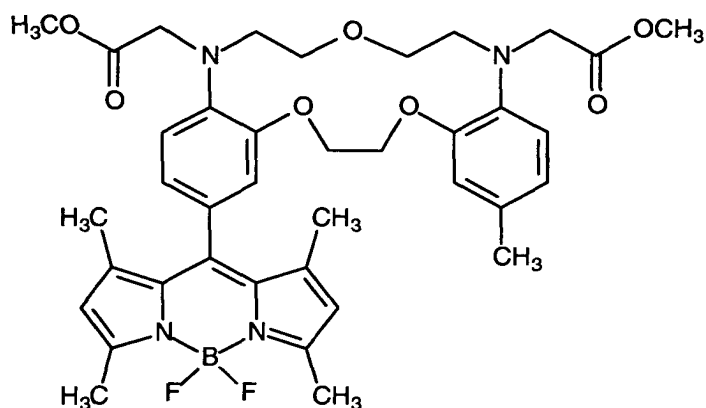


Compound 31

10 **Example 27.** Preparation of Compound 32

- To a solution of aldehyde 13 (250 mg, 0.5 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (20 mL) was added 2,4-dimethylpyrrole (0.125 mL, 1.2 mmol). The solution was stirred for 5 min, then TFA (0.046 mL, 0.6 mmol) is introduced. After 16 h, the mixture was diluted with CHCl<sub>3</sub> (150 mL),
- 15 washed with 2% tetrabutylammonium hydroxide (150 mL) then H<sub>2</sub>O then evaporated. The residue was dissolved in toluene (20 mL) and stirred for 3 h with chloranil (148 mg, 0.6 mmol). DIPEA (0.87 mL, 5 mmol) was introduced, followed by BF<sub>3</sub> etherate (0.52 mL, 4 mmol). The mixture was evaporated and the residue was purified by chromatography on

SiO<sub>2</sub> using a gradient of 0-2% MeOH in CHCl<sub>3</sub> to give compound **32**, 102 mg (28%) as a brown solid.



Compound **32**

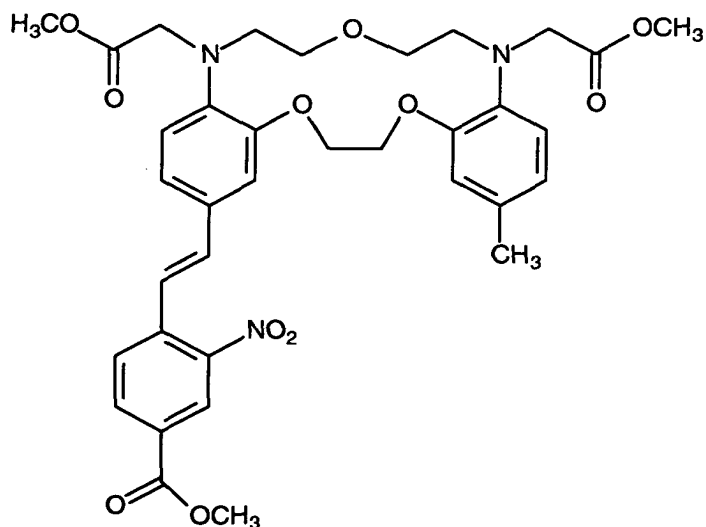
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**Example 28. Preparation of Compound 33**

A mixture of aldehyde **13** (1.00 g, 2.0 mmol), the Wittig base (4-methoxycarbonyl-2-nitrobenzyl)triphenylphosphonium bromide (1.34 g, 2.5 mmol), and K<sub>2</sub>CO<sub>3</sub> (1.38 g, 10 mmol) in DMF (20 mL) was stirred for 16 h at 95 °C. More of the Wittig base (0.500 g, 0.93 mmol) was added and the mixture was stirred for an additional 6 h, then cooled to room temperature and poured into H<sub>2</sub>O. The solution was acidified with 1 M HCl to pH 3, extracted with CHCl<sub>3</sub>, dried over MgSO<sub>4</sub>, and evaporated. The crude product was purified by chromatography on SiO<sub>2</sub> using CHCl<sub>3</sub> then 2% MeOH in CHCl<sub>3</sub> to give compound **33**, 0.915 g (68%) as an orange solid.

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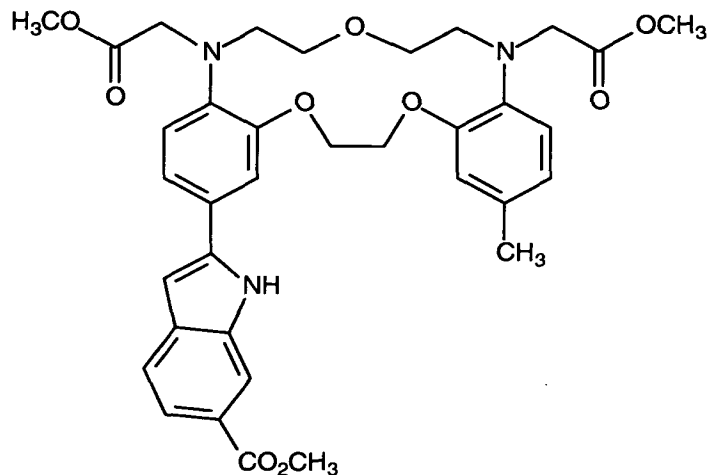
15



Compound **33**

**Example 29. Preparation of Compound 34**

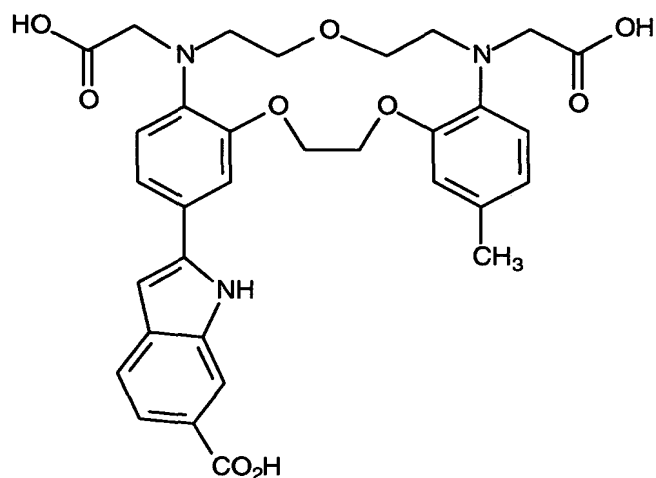
A mixture of the ethylene derivative **33** (70 mg, 0.1 mmol) and triethylphosphite (3 mL) was heated at 120 °C for 6 h, then evaporated and subsequently co-evaporated with DMF (3 x 10 mL). The residue was purified by preparative TLC on two SiO<sub>2</sub> plates using 7% MeOH in CHCl<sub>3</sub> as eluant to give compound **34**, 52 mg (80%) as a slightly yellowish solid.



**Compound 34**

**Example 30. Preparation of Compound 35**

A mixture of Compound **34** (300 mg, 0.46 mmol), and 1 M KOH (6 mL, 6 mmol) in MeOH (25 mL) was stirred for 16 h. More 1 M KOH (5 mL, 5 mmol) was added and the mixture was stirred for an additional 6 h, then evaporated to half its original volume, and acidified with 1 M HCl to pH 3. The suspension was extracted with CHCl<sub>3</sub> then with n-BuOH. The combined organic extract was filtered and was evaporated. Ether (20 mL) was added and the precipitated solid was filtered and washed with ether to give compound **35**, 157 mg (57%) as an off-white solid.

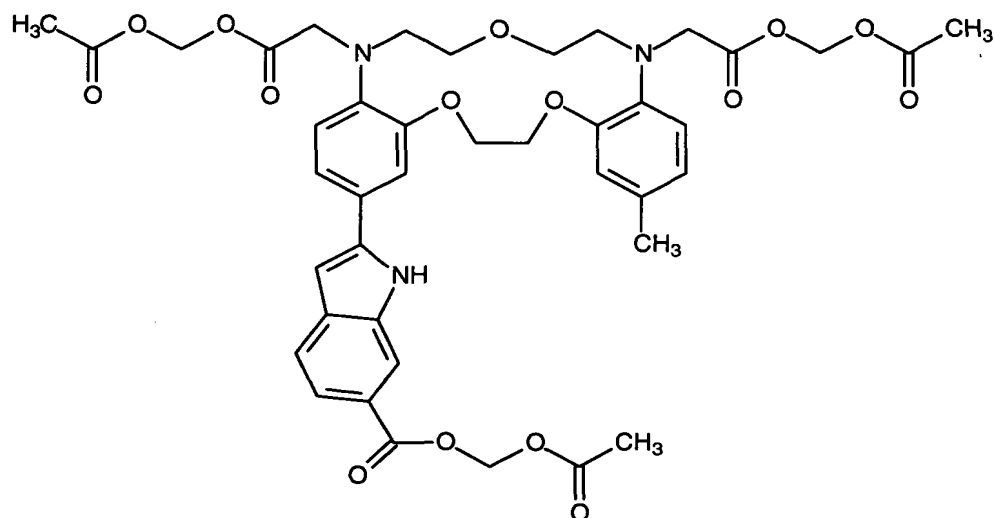


Compound **35**

**Example 31.** Preparation of Compound **36**

- 5 A mixture of Compound **35** (72 mg, 0.12 mmol), bromomethyl acetate (0.14 mL, 1.0 mmol), and DIPEA (0.35 mL, 2 mmol) in DMF (5 mL) was stirred for 16 h, poured into 1% AcOH (200 mL), extracted with  $\text{CHCl}_3$ , washed with  $\text{H}_2\text{O}$  then evaporated. The residue was purified by preparative TLC on two  $\text{SiO}_2$  plates using 5% MeOH in  $\text{CHCl}_3$  to give Compound **36**, 3 mg (3%) as a brown solid.

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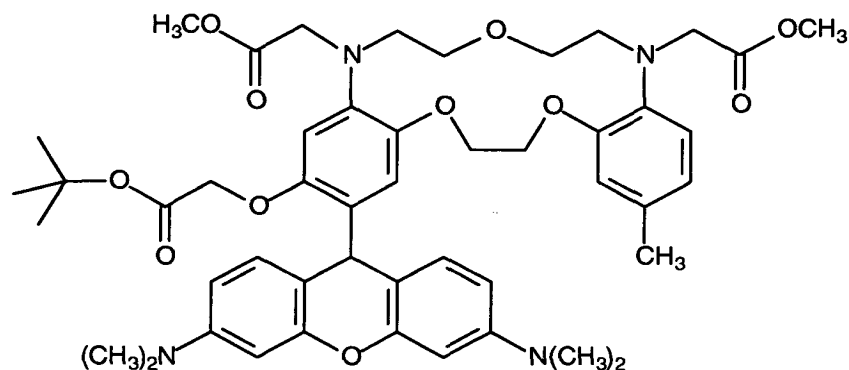
Compound **36**

**Example 32.** Preparation of Compound **37**

- 15 A mixture of aldehyde **16** (630 mg, 1.0 mmol), 3-dimethylaminophenol (330 mg, 2.4 mmol), and p-toluenesulphonic acid (20 mg, catalyst) in propionic acid (5 mL) was stirred for 16 h at

60 °C, then cooled to room temperature and poured into 3 M NaOAc (150 mL). The precipitated solid was filtered, washed with H<sub>2</sub>O and dried to give Compound **37**, 360 mg (97%) as a rose solid. Compound **37** is unstable to oxidation and was used in next step without additional purification.

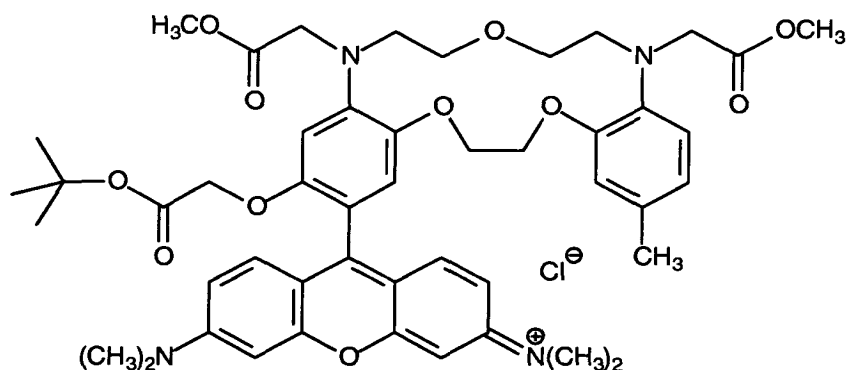
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Compound **37**

### Example 33. Preparation of Compound **38**

- 10 A mixture of Compound **37** (2.50 g, 2.9 mmol) and chloranil (1.42 g, 5.8 mmol) in MeOH (100 mL) and CHCl<sub>3</sub> (100 mL) was stirred for 4 h, filtered from excess oxidizer, and evaporated. The residue was purified by chromatography on SiO<sub>2</sub> using a gradient of 6-10% MeOH and 1% AcOH in CHCl<sub>3</sub> to give compound **38**, 1.13 g (45%) as a crimson solid.

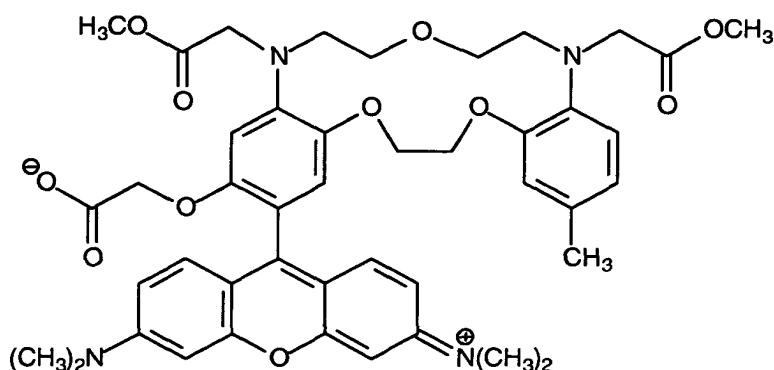


Compound **38**

### Example 34. Preparation of Compound **39**

- 20 A mixture of Compound **38** (500 mg, 0.58 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (20 mL) and TFA (20 mL) was stirred for 4 h, then evaporated and co-evaporated with CHCl<sub>3</sub>. Ether (25 mL) was added to the residue and the precipitated product was filtered and washed with ether to give Compound **39**, 447 mg (95%) as a violet-red solid.





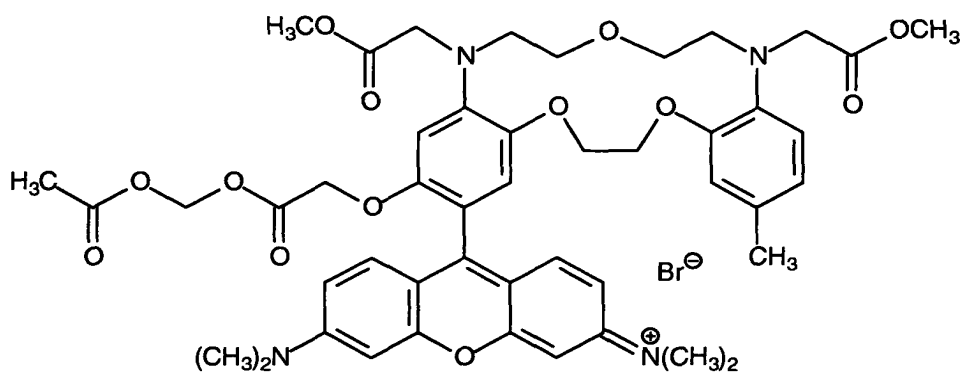
Compound **39**

**Example 35.** Preparation of Compound **40**

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A mixture of Compound **39** (20 mg, 0.1 mmol), bromomethyl acetate (0.025 mL, 0.25 mmol), and DIPEA (0.08 mL, 0.5 mmol) in DMF (1 mL) was stirred for 3 h. More bromomethyl acetate (0.025 mL, 0.25 mmol) and DIPEA (0.08 mL, 0.5 mmol) were added and the mixture was stirred for an additional 3 h, diluted with CHCl<sub>3</sub> (100 mL), washed with 1% AcOH then H<sub>2</sub>O then evaporated. The residue was purified by preparative TLC on two SiO<sub>2</sub> plates using 10% MeOH and 0.5% AcOH in CHCl<sub>3</sub> to give Compound **40**, 12 mg (54%) as a dark red solid.

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Compound **40**

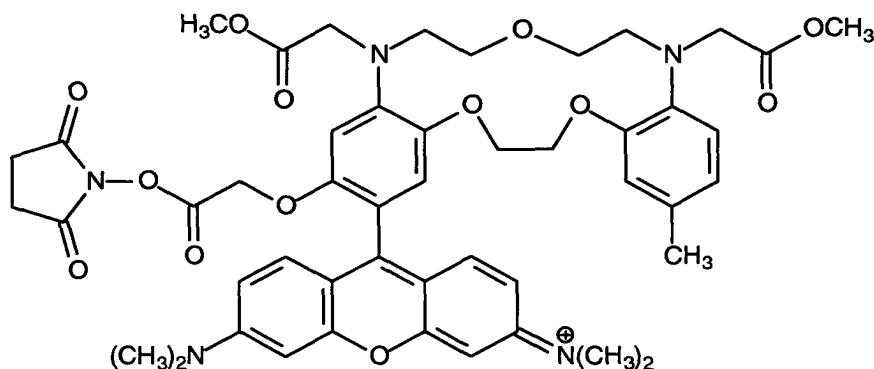
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**Example 36.** Preparation of Compound **41**

To a stirred solution of Compound **39** (162 mg, 0.2 mmol) and pyridine (0.08 mL, 1 mmol) in DMF (5 mL), was added dry N-trifluoroacetoxysuccinimide (90 mg, 0.4 mmol). After 7 h more N-trifluoroacetoxysuccinimide (90 mg, 0.4 mmol) and pyridine (0.08 mL, 1 mmol) were added, and the mixture was stirred for 16 h more. Analytical TLC confirmed the formation of the single product, while starting material **39** was consumed. Compound **41** was very

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reactive and unstable towards isolation attempts. It was used in further transformations upon preparation in DMF solution without isolation and purification.



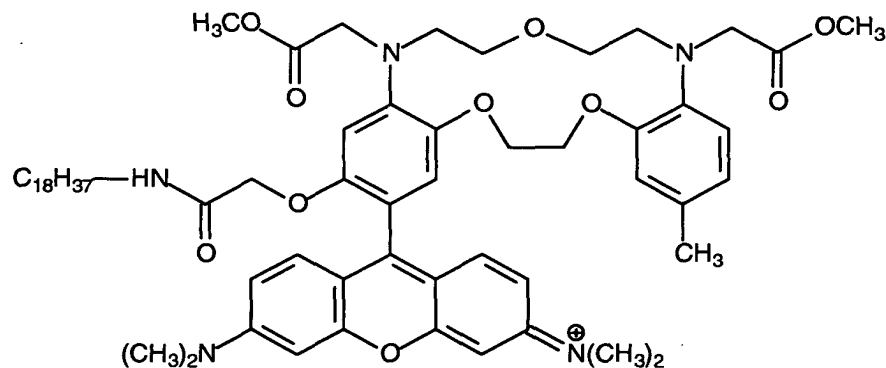
Compound 41

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### Example 37. Preparation of Compound 42

To a stirred solution of succinimidyl ester 41 prepared as in Example 36 from compound 39 (81 mg, 0.1 mmol) in DMF (2 mL) was added dry hexadecylamine (135 mg, 0.5 mmol). The mixture was stirred for 16 h, diluted with CHCl<sub>3</sub> (150 mL), washed with 1% AcOH the H<sub>2</sub>O, dried over MgSO<sub>4</sub>, and evaporated. The residue was purified by preparative TLC on two SiO<sub>2</sub> plates using CHCl<sub>3</sub> as eluant to give Compound 42, 57 mg (54%) as a red oil.

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Compound 42

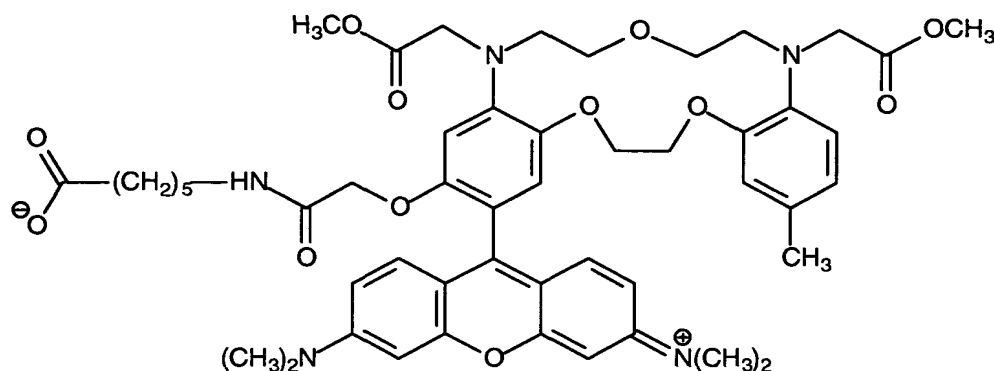
15

### Example 38. Preparation of Compound 43

To a stirred solution of succinimidyl ester 41 prepared as in Example 36 from Compound 39 (81 mg, 0.1 mmol) in DMF (2 mL) was quickly introduced a solution of 5-aminopentanoic acid (40 mg, 0.3 mmol) and 1 M methanolic tetrabutylammonium hydroxide (1 mL, 1 mmol) in H<sub>2</sub>O (5 mL). The mixture was stirred for 4 h, then diluted with H<sub>2</sub>O (100 mL), acidified with AcOH to pH 4, extracted with CHCl<sub>3</sub>, dried over MgSO<sub>4</sub> and evaporated. The residue was

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purified by preparative TLC on two SiO<sub>2</sub> plates using 10% MeOH and 2.5% AcOH in CHCl<sub>3</sub> as the eluant to give Compound **43**, 56 mg (61%) as a red solid.



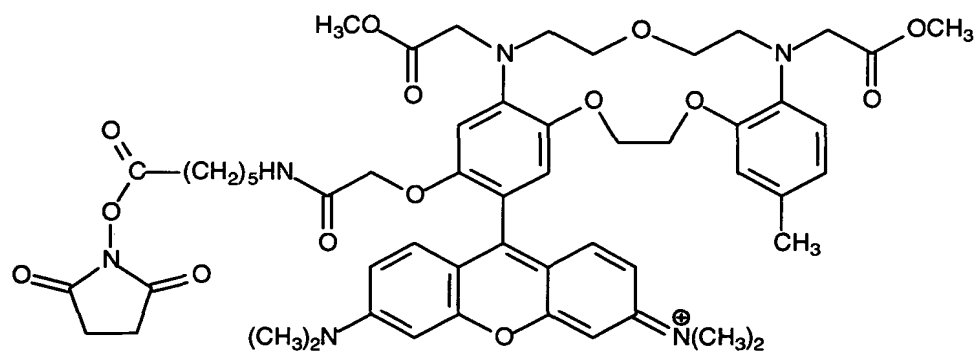
Compound **43**

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### Example 39. Preparation of Compound **44**

To a stirred solution of Compound **43** (37 mg, 0.04 mmol) and pyridine (0.08 mL, 1 mmol) in DMF (1 mL), was added dry N-trifluoroacetoxysuccinimide (36 mg, 0.16 mmol). After 4 h, more N-trifluoroacetoxysuccinimide (36 mg, 0.4 mmol) and pyridine (0.08 mL, 1 mmol) are added, and the mixture was stirred for an additional 16 h. The mixture was diluted with CHCl<sub>3</sub> (100 mL), washed with 1% AcOH then H<sub>2</sub>O then evaporated. Hexanes (5 mL) were added to the residue. The precipitated solid was filtered, washed with hexanes and dried to give Compound **44**, 15 mg (37%) as a dark red solid.

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Compound **44**

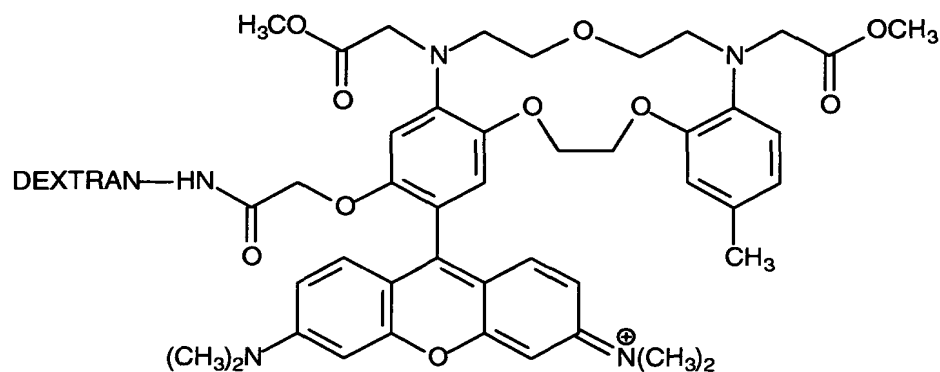
### Example 40. Preparation of Compound **45**

A solution of succinimidyl ester **41** (81 mg, 0.1 mmol) in DMF (2 mL) was added to 5 mL of aqueous aminodextran (100 mg, 0.037 eq.) and 1 M methanolic tetrabutylammonium hydroxide (1 mL, 1 mmol). The mixture was stirred for 16 h, poured into MeOH (400 mL),

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and the precipitated conjugate was filtered off and washed with MeOH. The crude product was dissolved in H<sub>2</sub>O (3 mL), filtered through a membrane filter and loaded onto a SEPHADEX G-15 resin column pre-equilibrated with H<sub>2</sub>O. The colored fraction was eluted in a void volume and lyophilized to give labeled dextran **45**, (41 mg) as a red solid.

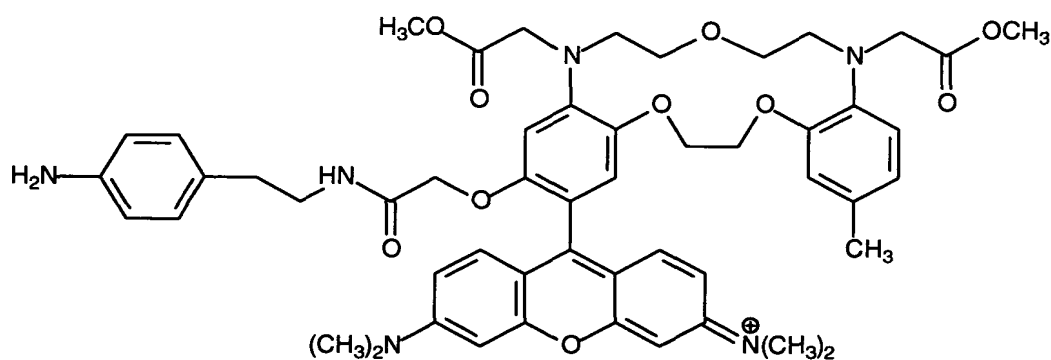
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Compound **45**

#### Example 41. Preparation of Compound **46**

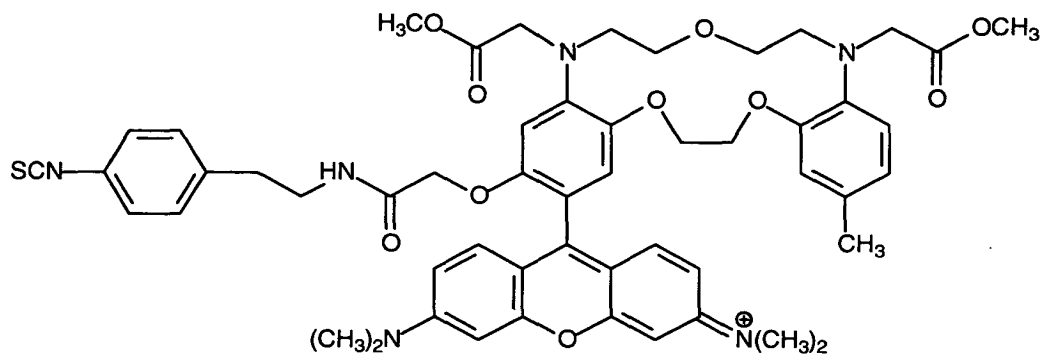
- 10 To a stirred solution of succinimidyl ester **41** (81 mg, 0.1 mmol) in DMF (2 mL) was added Et<sub>3</sub>N (0.14 mL, 1 mmol) followed by 2-(4-aminophenyl)ethylamine (0.05 mL, 0.5 mmol). The mixture was stirred for 2 h, diluted with CHCl<sub>3</sub> (150 mL), washed with 1% AcOH then H<sub>2</sub>O then evaporated. Ether (5 mL) was added to the residue, and the solid was filtered off, washed with ether and dried to give Compound **46** as a dark red solid, pure on TLC and
- 15 HPLC.



Compound **46**

- 20 **Example 42. Preparation of Compound 47**

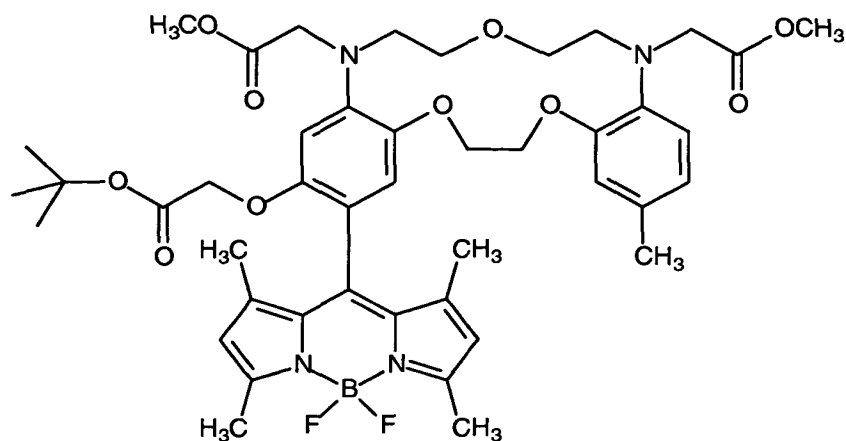
To a solution of Compound **46** (9 mg, 0.01 mmol) in AcOH (2 mL) was added a 1 M CSCI<sub>2</sub> solution in CHCl<sub>3</sub> (0.1 mL, 0.1 mmol). The mixture was stirred for 2 h then evaporated. Ether (5 mL) was added to the residue, and the precipitated solid was filtered off, washed with ether and dried to give Compound **47**, 9 mg (95%) as a dark red solid. Compound **47** reacts quickly with n-BuNH<sub>2</sub> in pre-column derivatization for HPLC analysis.



Compound **47**

#### 10 Example 43. Preparation of Compound **48**

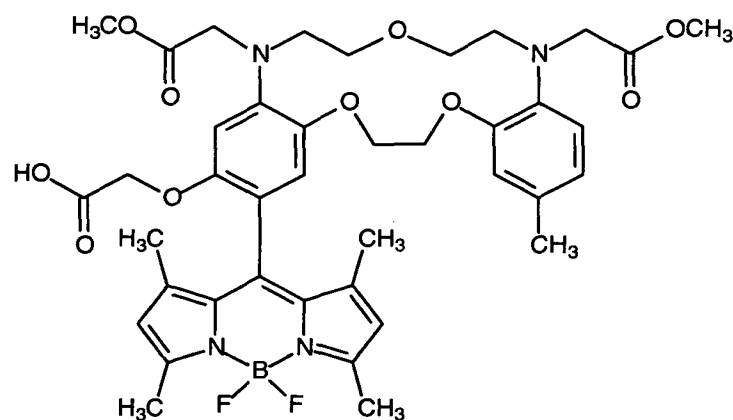
To a stirred solution of aldehyde **16** (630 mg, 1.0 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (40 mL), 2,4-dimethylpyrrole (0.25 mL, 2.4 mmol) was added. The solution was stirred for 5 min, then TFA (0.1 mL, 1.2 mmol) was introduced. The mixture was stirred for 16 h, and diluted with CHCl<sub>3</sub> (300 mL). The chloroform solution was washed with 2% tetrabutylammonium hydroxide then H<sub>2</sub>O, evaporated, and subsequently co-evaporated with toluene. The residue was dissolved in toluene (40 mL), stirred for 2 h with chloranil (296 mg, 1.2 mmol), then DIPEA (1.74 mL, 10 mmol) was introduced, followed by BF<sub>3</sub> etherate (1.04 mL, 8 mmol). The mixture was evaporated and the residue was purified by chromatography on SiO<sub>2</sub> using a gradient of 0-1% MeOH in CHCl<sub>3</sub> to give Compound **48**, 280 mg (35%) as a brown solid.



Compound 48

**Example 44. Preparation of Compound 49**

- 5 A mixture of Compound 48 (120 mg, 0.15 mmol) in  $\text{CHCl}_3$  (6 mL) and TFA (0.4 mL) was stirred for 16 h, then evaporated and co-evaporated with  $\text{CHCl}_3$ . The residue was purified by preparative TLC on two  $\text{SiO}_2$  plates using 10% MeOH in  $\text{CHCl}_3$  as eluant to give Compound 49, 82 mg (77%) as a brown solid.

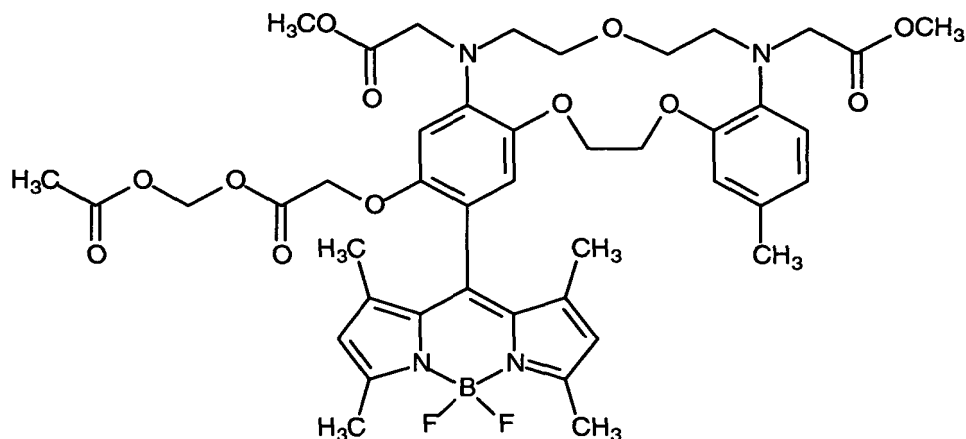


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Compound 49

**Example 45. Preparation of Compound 50**

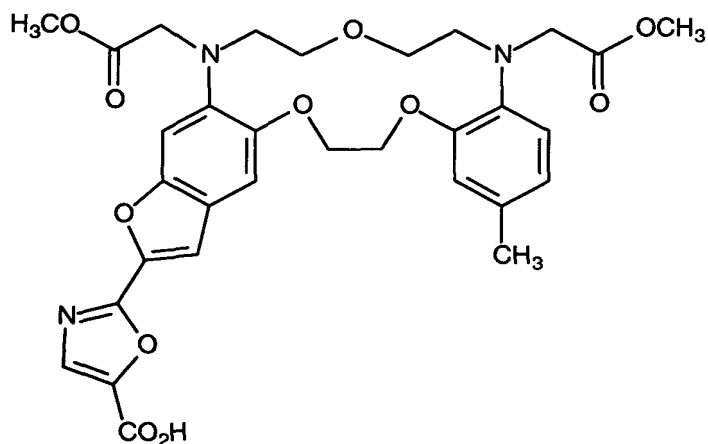
- 15 A mixture of Compound 49 (80 mg, 0.11 mmol), bromomethyl acetate (0.05 mL, 0.5 mmol), and DIPEA (0.17 mL, 1.0 mmol) in DMF (2 mL) was stirred for 2 h, poured into  $\text{H}_2\text{O}$  (150 mL), and extracted with  $\text{CHCl}_3$ . The extract was dried over  $\text{MgSO}_4$  then evaporated. The residue was purified by preparative TLC on two  $\text{SiO}_2$  plates using 3% MeOH and 0.5% AcOH in  $\text{CHCl}_3$  to give Compound 50, 46 mg (52%) as a brown solid.



### Compound 50

### Example 46. Preparation of Compound 51.

5 A mixture of aldehyde **15** (800 mg, 1.55 mmol), 5-carboxy-2-chloromethyloxazole (300 mg, 1.86 mmol), K<sub>2</sub>CO<sub>3</sub> (1.07 g, 7.75 mmol), and NaI (75 mg, 0.5 mmol; catalyst) in DMF (20 mL) was stirred at 135 °C for 4 h, then poured into H<sub>2</sub>O. The mixture was acidified with 1 M HCl to pH 3, extracted with CHCl<sub>3</sub>, dried over MgSO<sub>4</sub> then evaporated. The residue was purified by chromatography on SiO<sub>2</sub> using a gradient of 5–20% MeOH and 1-2% AcOH in CHCl<sub>3</sub> as  
10 eluant to give Compound **51**, 255 mg (26%) as a yellowish solid.



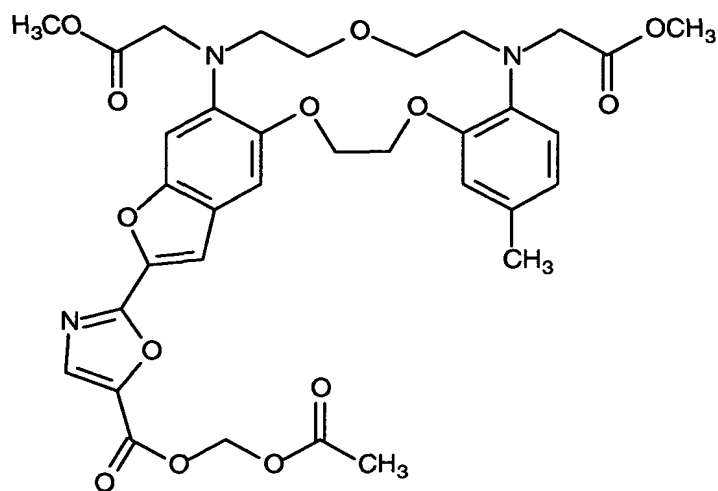
### Compound 51

### Example 47. Preparation of Compound 52

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A mixture of Compound **51** (62 mg, 0.1 mmol), bromomethyl acetate (0.05 mL, 0.5 mmol), and DIPEA (0.17 mL, 1.0 mmol) in DMF (1 mL) was stirred for 2 h, poured into 1% AcOH (100 mL), and extracted with CHCl<sub>3</sub>. Then 3 M NaOAc (100 mL) was added to the aqueous phase and the mixture was extracted with more CHCl<sub>3</sub>. The combined extracts were dried

over  $\text{MgSO}_4$  then evaporated. The residue was purified by preparative TLC on two  $\text{SiO}_2$  plates using 5% MeOH in  $\text{CHCl}_3$  to give Compound **52**, 29 mg (42%) as an orange solid.



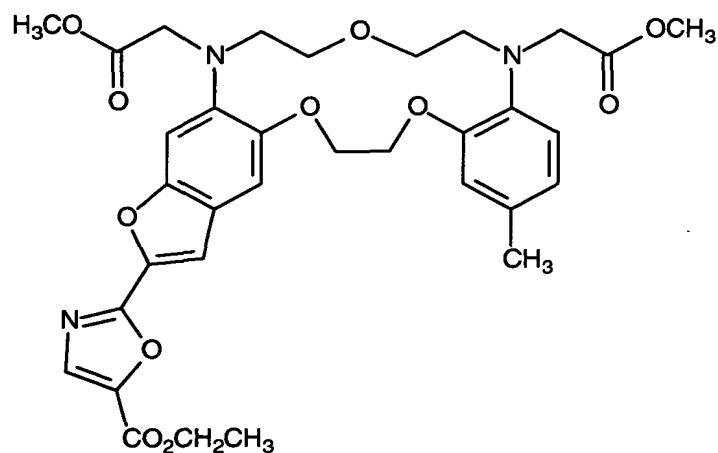
Compound **52**

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**Example 48. Preparation of Compound 53**

A mixture of aldehyde **15** (1.40 g, 2.71 mmol), 2-chloromethyl-5-ethoxycarbonyloxazole (0.564 g, 1.86 mmol),  $\text{K}_2\text{CO}_3$  (1.87 g, 13.60 mmol), and NaI (0.150 g, 1.00 mmol; catalyst) in DMF (30 mL) was stirred at 135 °C for 4 h and then poured into  $\text{H}_2\text{O}$ . The mixture was acidified with 1 M HCl to pH 3, extracted with  $\text{CHCl}_3$  and evaporated. The residue was purified by chromatography on  $\text{SiO}_2$  using 1% MeOH in  $\text{CHCl}_3$  as eluant to give Compound **53**, 1.260 g (71%) as a yellowish solid.

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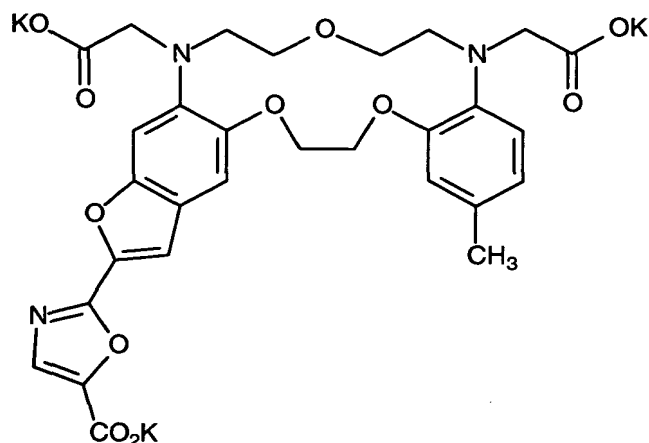
Compound **53**

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**Example 49. Preparation of Compound 54**



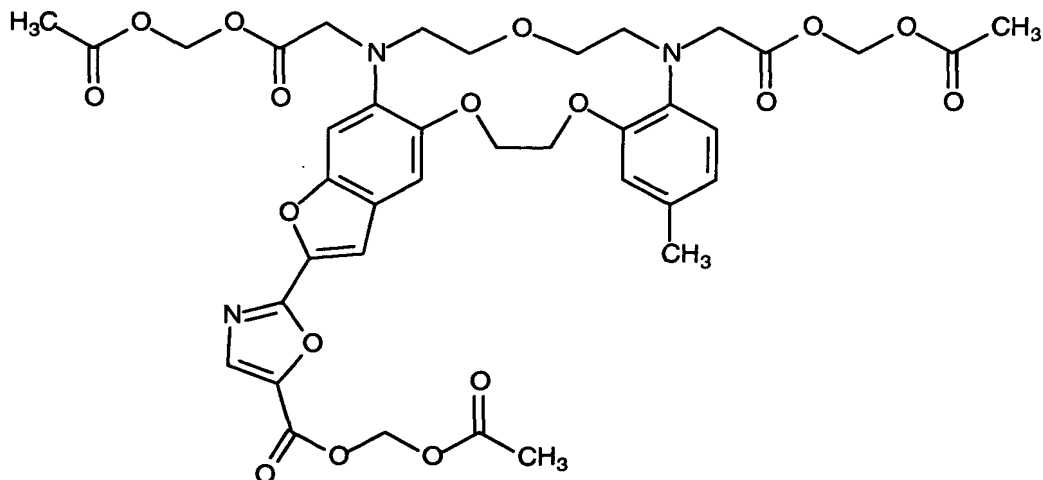
A mixture of Compound **53** (376 mg, 0.5 mmol) and 1 M KOH (3 mL, 3.0 mmol) in MeOH (5 mL) was stirred for 5 h, then evaporated to 1/5 volume, and acidified with 1 M HCl to pH 2.8. The precipitated solid was filtered, washed with H<sub>2</sub>O and dried on a filter. This crude product was suspended in H<sub>2</sub>O (2 mL) and then made basic with 0.1 M KOH to pH 9.5. The solution was loaded onto a SEPHADEX LH-20 resin column and chromatographed using H<sub>2</sub>O as eluant to give Compound **54**, 256 mg (72%) as a yellow-greenish solid.



Compound **54**

#### 10 Example 50. Preparation of Compound **55**

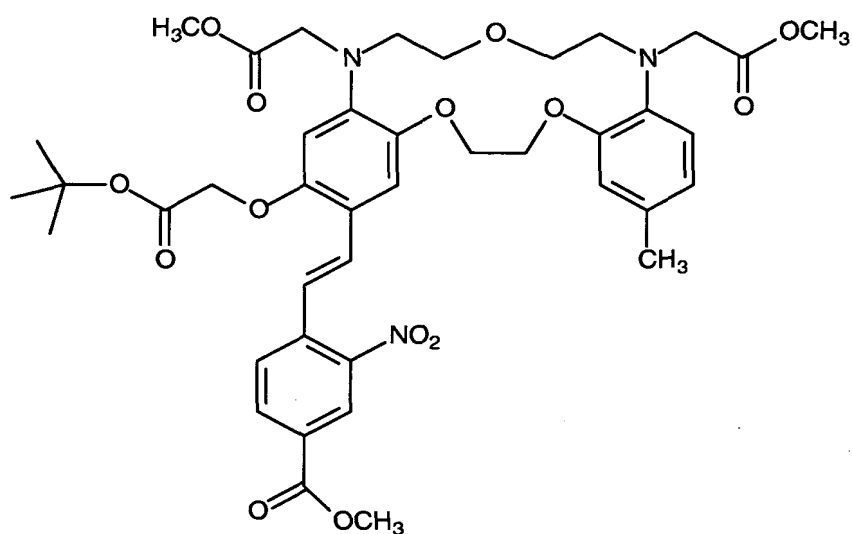
A mixture of Compound **54** (71 mg, 0.1 mmol), bromomethyl acetate (0.1 mL, 1.0 mmol), and DIPEA (0.35 mL, 2.0 mmol) in DMF (2 mL) was stirred for 4 h, diluted with CHCl<sub>3</sub>, then washed with 1% AcOH then H<sub>2</sub>O and evaporated. The residue was purified by preparative TLC on two SiO<sub>2</sub> plates using 5% MeOH and 0.5% AcOH in CHCl<sub>3</sub> to give compound **55**, 63 mg (77%) as an orange solid.



## Compound 55

### Example 51. Preparation of Compound 56

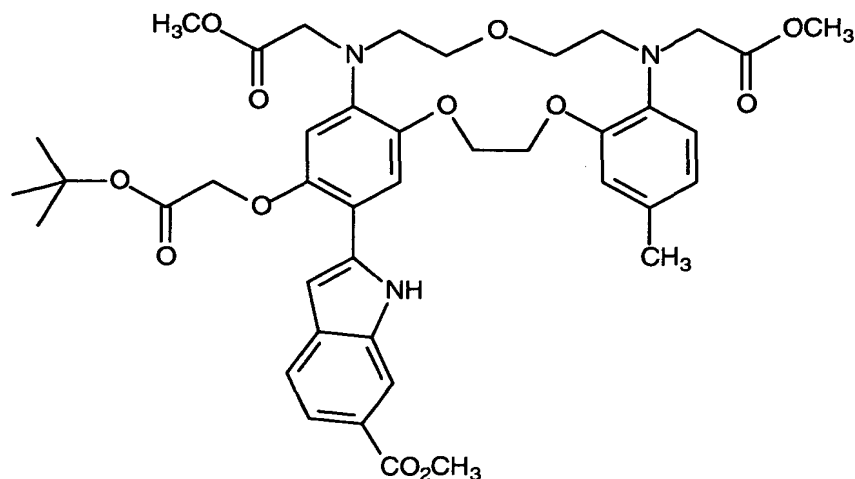
- 5 A mixture of aldehyde **16** (315 mg, 0.5 mmol), the Wittig base (4-methoxycarbonyl-2-nitrobenzyl)triphenylphosphonium bromide (375 mg, 0.7 mmol), and  $K_2CO_3$  (345 mg, 2.5 mmol) in DMF (3 mL) was stirred for 6 h at 95 °C then cooled to room temperature and poured into  $H_2O$ . The solution was acidified to pH 5 with 1 M HCl and extracted with  $CHCl_3$ . The extract was dried over  $MgSO_4$  and evaporated. The crude product was purified by
- 10 chromatography on  $SiO_2$  using  $CHCl_3$ , then 0.5% MeOH in  $CHCl_3$  to give Compound **56**, 339 mg (84%) as an orange solid.



Compound **56**

### 15 Example 52. Preparation of Compound 57

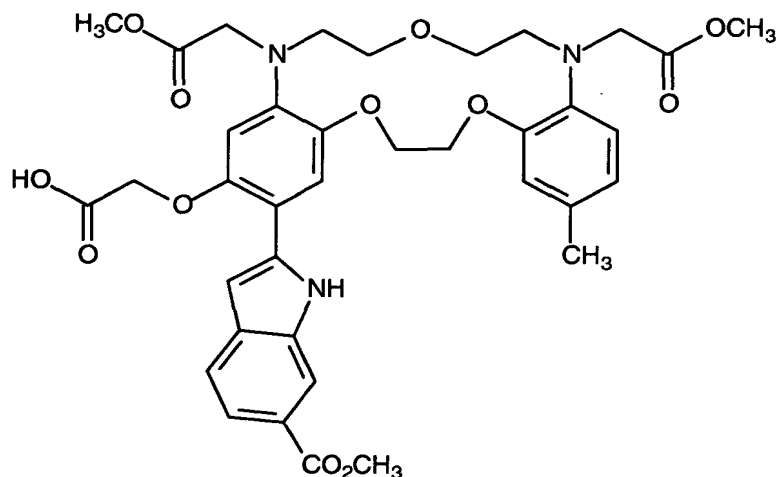
- A mixture of the ethylene derivative **56** (338 mg, 0.42 mmol) and triethylphosphite (8 mL) was heated at 130 °C for 7 h, then evaporated and subsequently co-evaporated with DMF. The residue was first purified by chromatography on  $SiO_2$  using  $CHCl_3$  followed by 1%
- 20 MeOH in  $CHCl_3$ , then by preparative TLC on two  $SiO_2$  plates using 50% EtOAc in hexanes as eluant to give Compound **57**, 112 mg (34%) as a slightly yellowish solid.



Compound 57

**Example 53. Preparation of Compound 58**

- 5 A mixture of Compound 57 (104 mg, 0.13 mmol) in  $\text{CH}_2\text{Cl}_2$  (2 mL) and TFA (2 mL) was stirred for 3 h, then evaporated and co-evaporated with  $\text{CHCl}_3$ . The residue was purified by preparative TLC on two  $\text{SiO}_2$  plates using 7% MeOH and 2% AcOH in  $\text{CHCl}_3$  to give Compound 58, 55 mg (57%) as a yellowish solid.

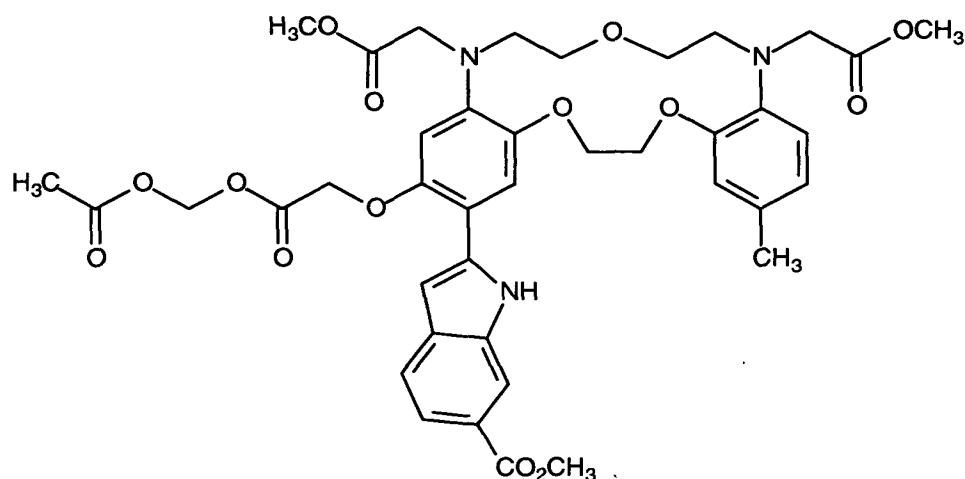


Compound 58

**Example 54. Preparation of Compound 59**

- 15 A mixture of Compound 58 (22 mg, 0.03 mmol), bromomethyl acetate (0.1 mL, 1.0 mmol), and DIPEA (0.26 mL, 1.5 mmol) in DMF (2 mL) was stirred for 16 h, then diluted with  $\text{CHCl}_3$  (100 mL), washed with 1% AcOH then  $\text{H}_2\text{O}$  then evaporated. The residue was purified by

preparative TLC on two SiO<sub>2</sub> plates using 5% MeOH in CHCl<sub>3</sub> to give compound **59**, 16 mg (67%) as an orange solid.



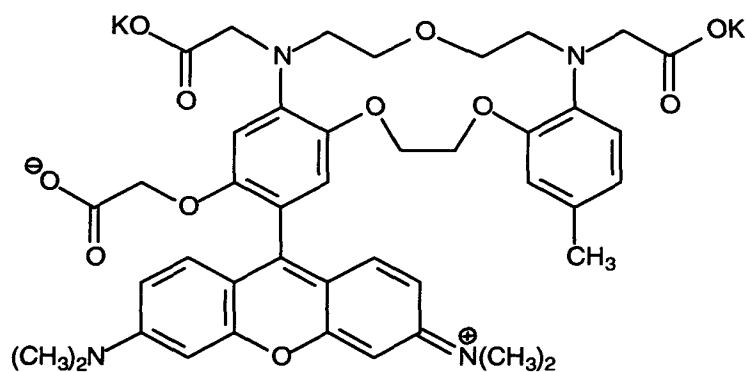
Compound **59**

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#### Example 55. Preparation of Compound **60**

A mixture of Compound **38** (87 mg, 0.1 mmol) and 1 M KOH (2 mL, 2 mmol) in MeOH (5 mL) was stirred for 16 h, and evaporated. The residue was purified on SEPHADEX LH-20 resin using H<sub>2</sub>O as the eluant. The fractions containing product were collected and lyophilized to give Compound **60**, 12 mg (14%) as a crimson solid.

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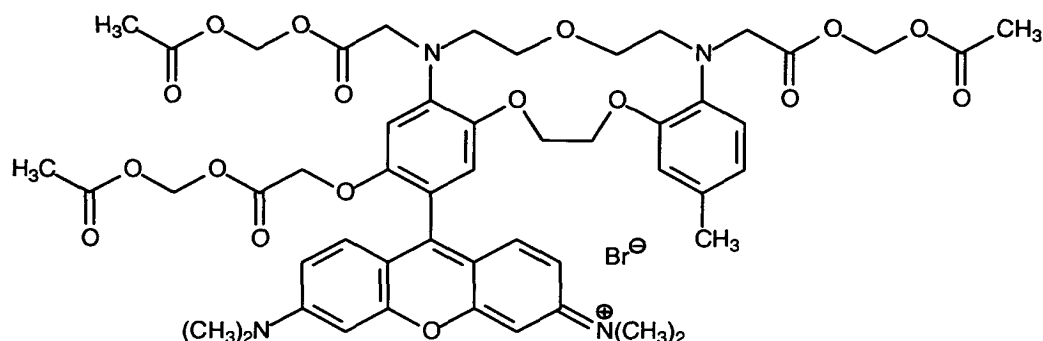
Compound **60**

#### Example 56. Preparation of Compound **61**

A mixture of Compound **60** (8 mg, 0.01 mmol), bromomethyl acetate (0.025 mL, 0.25 mmol), and DIPEA (0.08 mL, 0.5 mmol) in DMF (1 mL) was stirred for 16 h, and diluted with CHCl<sub>3</sub> (100 mL). The solution was washed with 1% AcOH then H<sub>2</sub>O then evaporated. Ether (5 mL) was added to the residue. The precipitated solid was filtered off and washed with ether to

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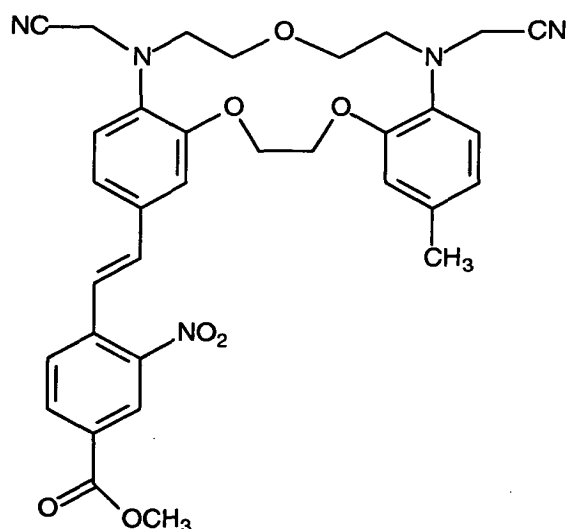
give Compound **61**, 5 mg (50%) as a dark red solid. Compound **61** was pure on TLC and HPLC.



## 5 Compound 61

### Example 57. Preparation of Compound 62

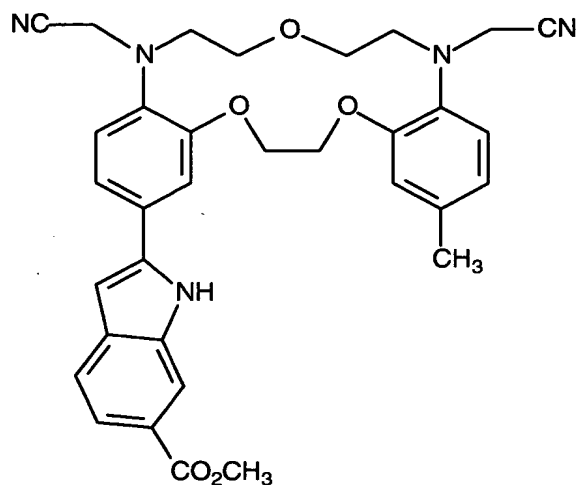
A mixture of aldehyde **18** (217 mg, 0.5 mmol), the Wittig base (4-methoxycarbonyl-2-nitrobenzyl)triphenylphosphonium bromide (536 mg, 0.7 mmol), and  $K_2CO_3$  (345 mg, 2.5 mmol) in DMF (3 mL) was stirred for 16 h at 95 °C, then cooled to room temperature and poured into  $H_2O$ . The solution was acidified with 1 M HCl to pH 3, extracted with  $CHCl_3$ , dried over  $MgSO_4$  and evaporated. The crude product was purified by chromatography on  $SiO_2$  using  $CHCl_3$ , as eluant to give Compound **62**, 171 mg (56%) as an orange solid.



Compound **62**

### Example 58. Preparation of Compound 63

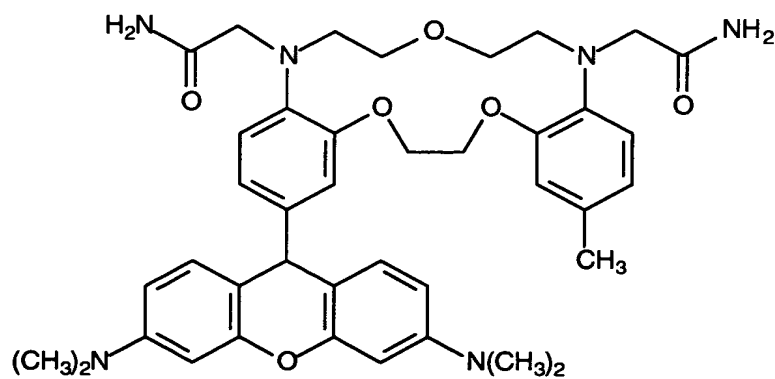
A mixture of the ethylene derivative **62** (170 mg, 0.28 mmol) and triethylphosphite (4 mL) was heated at 120 °C for 16 h, then evaporated and subsequently co-evaporated with DMF. The residue was purified by preparative TLC on two SiO<sub>2</sub> plates using 3% MeOH in CHCl<sub>3</sub> as eluant, then on two SiO<sub>2</sub> plates using 40% EtOAc in hexanes as eluant to give Compound **63**, 26 mg (16%) as a yellowish solid.



Compound **63**

#### Example 59. Preparation of Compound **64**

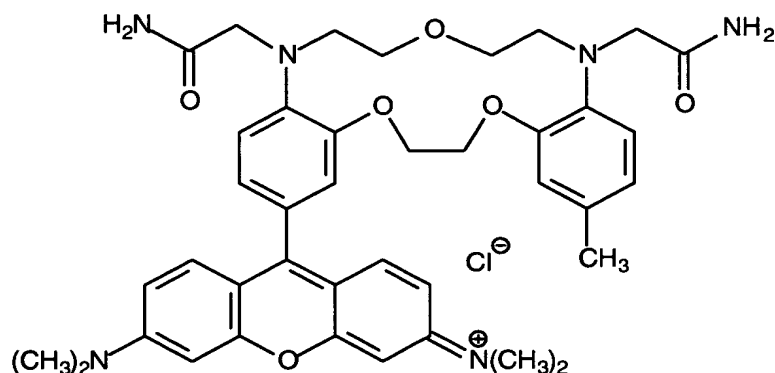
A mixture of aldehyde **20** (100 mg, 0.21 mmol), 3-dimethylaminophenol (55 mg, 0.40 mmol), and p-toluenesulphonic acid (5 mg, catalyst) in propionic acid (2 mL) was stirred overnight at 60 °C, then cooled to room temperature and poured into 3 M NaOAc (40 mL). Precipitated solid was filtered, washed with H<sub>2</sub>O and dried to give Compound **64**, 140 mg (93%) as a rose-colored solid. Compound **64** was unstable to oxidation and was used in next step without additional purification.



Compound **64**

#### Example 60. Preparation of Compound **65**

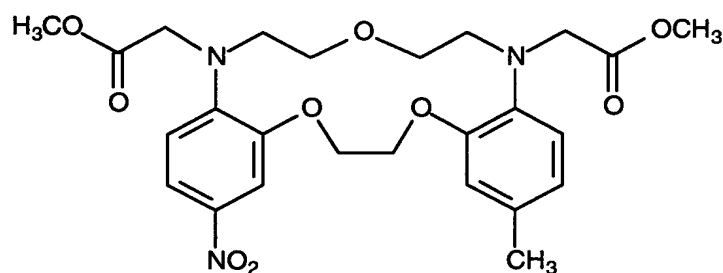
A mixture of Compound **64** (138 mg, 0.20 mmol) and chloranil (84 mg, 0.34 mmol) in MeOH (5 mL) and CHCl<sub>3</sub> (5 mL) was stirred for 4 h, and evaporated. The residue was purified by chromatography on SiO<sub>2</sub> using a gradient of 7-9% MeOH and 0.5-1% AcOH in CHCl<sub>3</sub> as eluant to give Compound **65**, 118 mg (86%) as a dark red solid.



Compound **65**

#### 10 Example 61. Preparation of Compound **66**

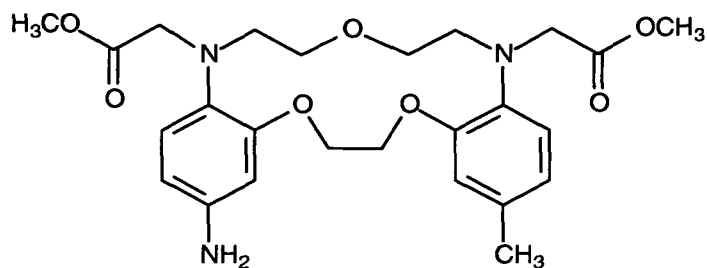
To a 0.5 M solution of arene **12** in sulfuric acid is added one equivalent of potassium nitrate. The resulting solution is stirred until TLC analysis of reaction aliquots, treated with water and ether, shows full conversion to nitroarene **66**. The reaction mixture is poured into excess aqueous sodium acetate, followed by extraction with ether. The extract is washed with aqueous sodium bicarbonate and brine, dried over magnesium sulfate and then concentrated to a reddish-brown oil, which is purified by trituration with ether-hexanes to give Compound **66** as a yellow powder.



Compound **66**

#### Exempl 62. Preparation of Compound **67**

A 0.5 M solution of Compound **66** in methanol is treated with 10% palladium on charcaol, at 10 wt% of Compound **66**. The resulting mixture is shaken under 40 psi hydrogen until TLC analysis shows full conversion to aniline **67**. The reaction mixture is filtered through diatomaceous earth, followed by concentration in vacuo. The residue is triturated with ether-  
5 hexanes to give pure Compound **67** as an off-white powder.

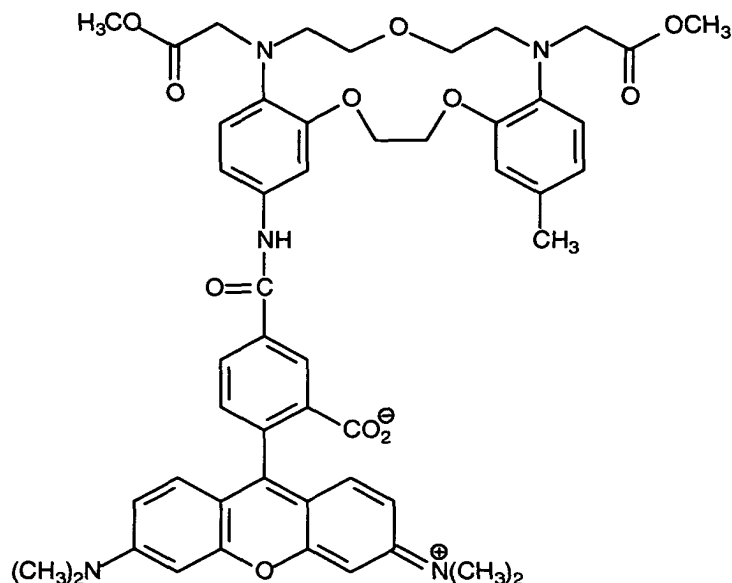


Compound **67**

**Example 63. Preparation of Compound 68**

A mixed anhydride of 6-carboxytetramethylrhodamine is prepared according to the procedure given in U.S. Patent No. 5,453,517 to Kuhn et al. To a solution of the mixed anhydride and one equivalent of diisopropylethylamine in anhydrous THF under nitrogen is added slowly a 0.3 M solution of Compound **67**. The resulting mixture is stirred until TLC  
15 analysis indicates consumption of Compound **67**. The reaction mixture is concentrated in vacuo, and the residue partitioned between chloroform and water. The chloroform layer is washed with brine and dried over magnesium sulfate, then concentrated in vacuo. The residue is purified by flash chromatography on silica gel using increasing amounts of methanol in chloroform as eluant to give pure Compound **68** as a red powder.

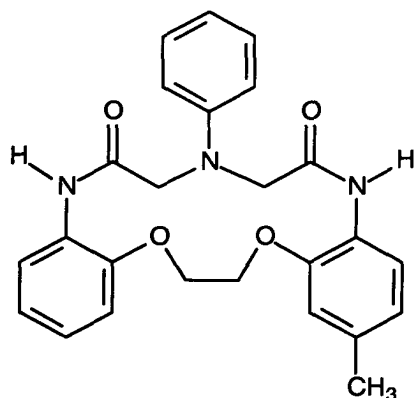




Compound **68**

**Example 64.** Preparation of a tris-aza crown ether (**69**).

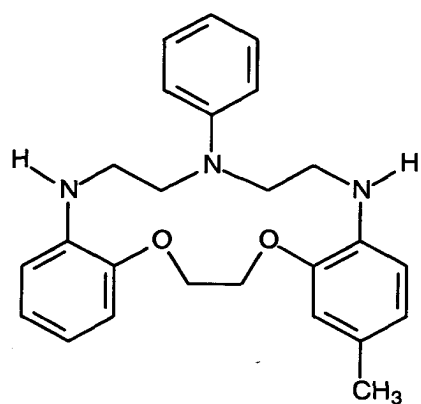
- 5 To a dilute solution (0.1 M) of diamine **9** and two equivalents of DIPEA in anhydrous THF is slowly added at room temperature a dilute solution (0.1 M) of N,N-bis(chlorocarbonylmethyl)-aniline. After stirring overnight, the volatiles are removed by evaporation, and the residue is partitioned between 5% hydrochloric acid and ethyl acetate. The organic layer is washed with brine, dried, and concentrated. The resulting residue is purified by column
- 10 chromatography on silica gel using increasing amounts of methanol in chloroform to give pure Compound **69**.



Compound **69**

- 15 **Example 65.** Preparation of Compound **70**

The bisamide **69** is reduced to bisaniline **70** as described for Compound **11**.

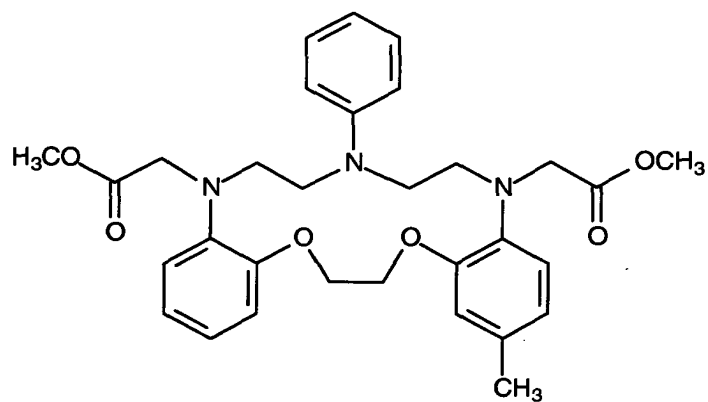


Compound **70**

**Example 66.** Preparation of Compound **71**

5

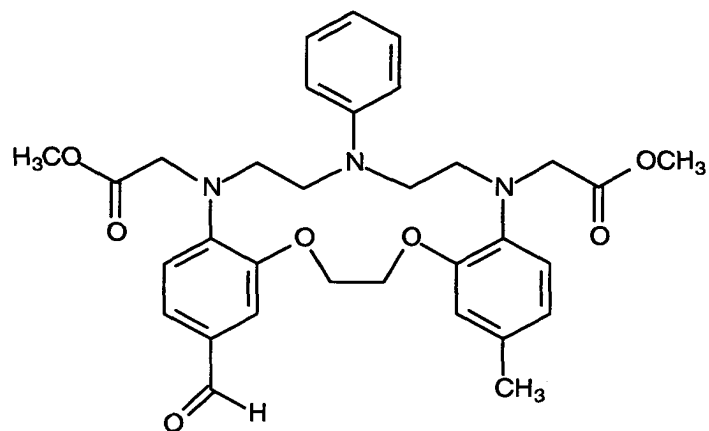
The bisaniline **70** is alkylated to give Compound **71**, as described for Compound **12**.



Compound **71**

10 **Example 67.** Preparation of Compound **72**

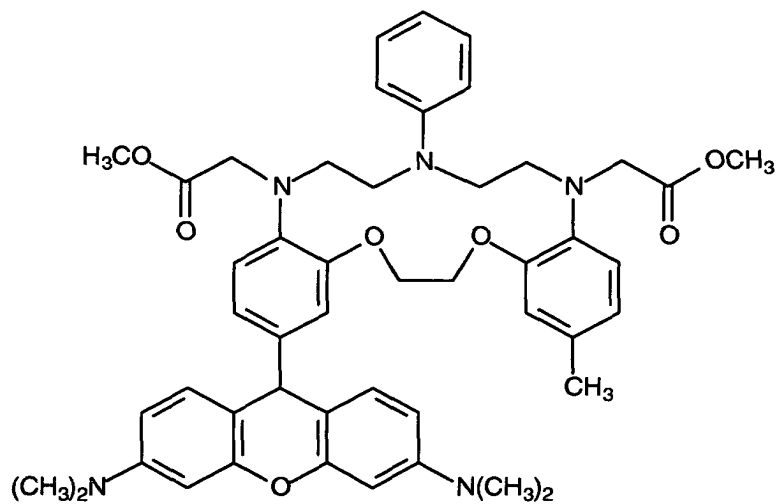
The bisester **71** is formylated to give the aldehyde Compound **72**, as described for Compound **13**.



Compound 72

**Example 68.** Preparation of Compound 73

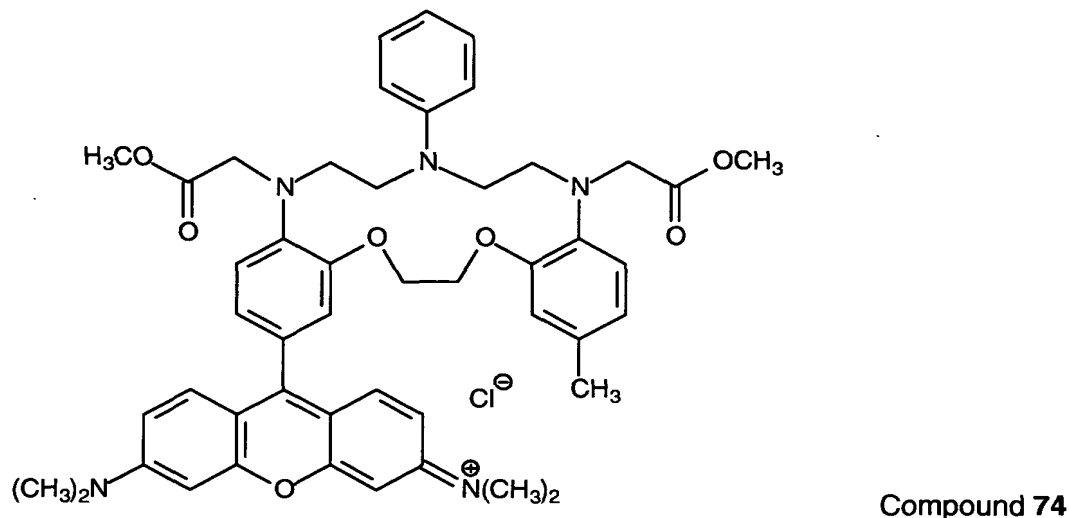
- 5 The aldehyde 72 is condensed with two equivalent of 3-(dimethylamino)phenol as described for Compound 26 to give Compound 73.



Compound 73

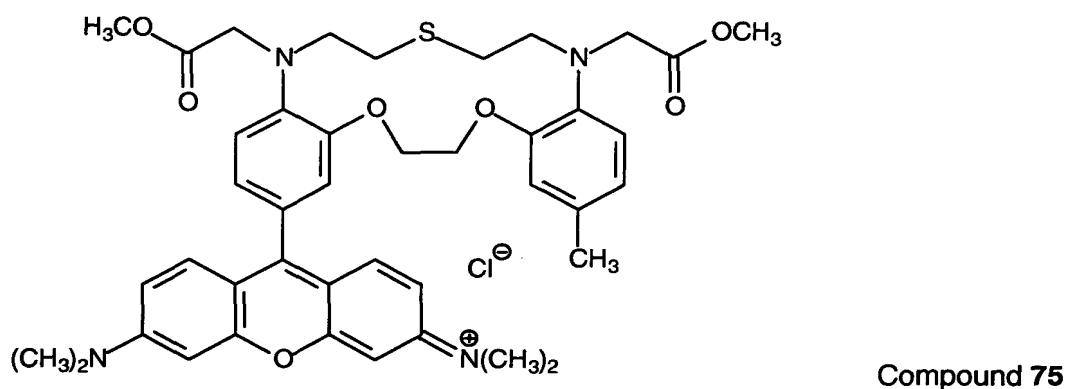
10 **Example 69.** Preparation of Compound 74

Compound 73 is oxidized with chloranil to give the chloride salt 74 as a red powder, as described for Compound 27.



**Example 70.** Preparation of a thia-substituted crown ether (**75**).

- 5 A bis-aza crown ether that incorporates a sulfur atom in the crown is prepared using the procedures of Examples 64-69, except that in place of N,N-bis(chlorocarbonylmethyl)-aniline, the bis-acid chloride of thiodiglycolic acid is used ( $\text{Cl}(\text{C}=\text{O})\text{CH}_2\text{-S-CH}_2(\text{C}=\text{O})\text{Cl}$ ). After preparation of the crown ether itself, condensation with 3-(N,N-dimethylamino)phenol, and oxidation, the chloride salt of the thia-crown ether, Compound **75**, is isolated as a red powder.
- 10



**Example 71.** Determination of  $\text{Na}^+$ ,  $\text{K}^+$ ,  $\text{Li}^+$  and  $\text{Tb}^+$  indicator fluorescence response.

- 15 The fluorescence response of a selected compound of the invention as a function of ion concentration was determined by dissolving a sample of the pure compound in 3 mL of each of two solutions: solution 1 ("high") consists of 200 mM NaCl (or other appropriate ion) and 10 mM MOPS buffer at pH 7.05; solution 2 ("zero ") consists of 10 mM MOPS buffer at pH 7.05 in

deionized water. A series of curves are generated by cross dilution between the two solutions to arrive at intermediate concentrations of Na<sup>+</sup>, K<sup>+</sup>, Li<sup>+</sup> and Tb<sup>3+</sup>. The emission of a selected indicator in solution 2 was scanned and then was repeated to cover the entire range from zero to 200 mM Na<sup>+</sup>. For example, the fluorescence emission of the chosen indicator was scanned while the sample was excited at that indicator's absorption maximum wavelength, and then 1/100 of the sample was removed and replaced with 1/100 of solution 1 to arrive at a Na<sup>+</sup> concentration of 2 mM. This dilution was repeated to cover the entire range from zero to 1 M Na<sup>+</sup> and the resulting emission intensities were plotted versus the ion concentrations. A least-squares fit was used to arrive at the concentration where the selected indicator was maximally sensitive to changes in Na<sup>+</sup> concentration. This corresponded to the dissociation constant of that indicator for Na<sup>+</sup> and was expressed as a concentration.

This methodology was repeated for K<sup>+</sup>, Li<sup>+</sup> and Tb<sup>3+</sup> to obtain emission spectra and to calculate the K<sub>d</sub> values. See, Table 2 and Figure 4

For visible wavelength probes such as Compound **22** (Example 17), Compound **27** (Example 22) and Compound **86** (Example 79) the indicator's fluorescence emission was typically scanned from 450-650 nm while the sample was excited at the absorption maximum wavelength. For UV-excitable ratiometric probes such as Compound **51** (Example 46), the excitation wavelengths of the dye in solution 2 were scanned from 260 to 450 nm while monitoring constant fluorescence emission at 510 nm (as in Figure 1).

The binding affinity of a selected indicator for sodium ions, in the presence of potassium ions, is determined using the process above, except in the presence of 100 mM K<sup>+</sup> concentrations.

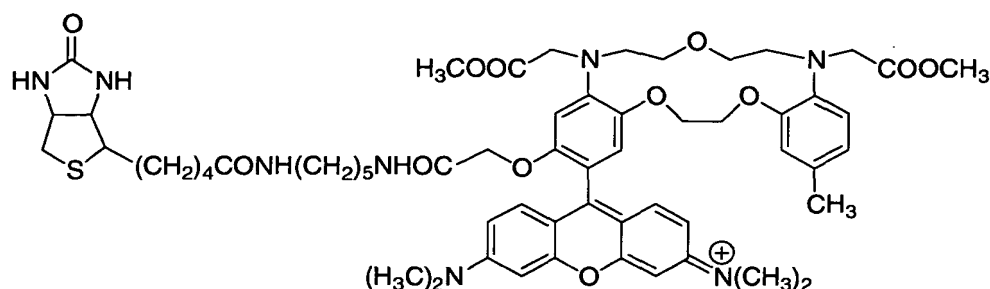
**Example 72:** Calibration of sodium indicator fluorescence response in cells.

Jurkat cells were loaded with a 5 μM solution of either Compound **25** (Example 20), or commercially available SODIUM GREEN tetraacetate indicator (Molecular Probes, Inc., Eugene, OR) and 10 μM for Compound **87** (Example 80) for 30 minutes at 37 °C. The use of PLURONIC dispersing agent helped dissolve the selected indicator. Intracellular sodium concentrations were then established by varying the sodium concentration of the extracellular buffer in the presence of 2 μM gramicidin (a sodium-pore forming antibiotic). The extracellular buffer was set at 0 mM, 10 mM, 20 mM, 50 mM, 100 mM, and 145 mM, respectively. Intracellular fluorescence response of the selected indicators was measured using a FACSCAN flow cytometer and associated software, with fluorescence excitation at 488 nm. Compound **25** exhibits substantially brighter intracellular fluorescence intensity

than the SODIUM GREEN indicator at comparable  $\text{Na}^+$  concentrations, as shown in Figure 3. Similarly, as intracellular  $\text{Na}^+$  concentration was increased, Compound **25** exhibits a consistent increase in fluorescence intensity. Compound **87** demonstrated an increased fluorescent signal with increasing concentrations of sodium ions.

5

**Example 73.** Preparation of a biotinylated indicator (Compound **76**)



Compound **76**

- 10 To a stirred solution of succinimidyl ester **41** prepared as in Example 36 from compound **39** (81 mg, 0.1 mmol) in DMF (2 mL) was added  $\text{Et}_3\text{N}$  (0.27 mL, 2 mmol) followed by biotin-cadaverin trifluoroacetate (66 mg, 0.2 mmol). The mixture was stirred for 16 h, diluted with  $\text{CHCl}_3$  (100 mL), washed 1% AcOH (3 x 50 mL),  $\text{H}_2\text{O}$  (100 mL) dried over  $\text{MgSO}_4$  and evaporated. The residue was purified by preparative TLC on four C18 reverse-phase plates
- 15 using 50% aqueous 2-PrOH with 0.2% TFA as eluant to give Compound **76**, 25 mg (22%) as an orange solid.

**Example 74.** Avidin-labeling with a biotinylated fluorescent crown ether.

- 20 A 5 mg/mL solution of streptavidin in phosphate-buffered saline (PBS, pH 7.0) is treated with a 1 mg/mL solution of biotinylated indicator (e.g., Compound **212**, Example 73) in 2% DMSO/PBS at a molar ratio such that two equivalents of biotinylated indicator are present for every streptavidin molecule. The resulting solution is incubated at  $37^\circ\text{C}$  for 4 hours and then centrifuged. The supernatant is applied to a Sephadex G-25 gel filtration column (2 mL bed volume/mg protein) and eluted with PBS. The streptavidin-indicator complex elutes first.
- 25 Fractions are analyzed by TLC to ensure that no free indicator is present in the complex. Pure product fractions are pooled and lyophilized. The complex is useful as a bridging method to apply the indicator to any biotinylated surface or substance.

**Example 75: Preparation of 1-Aza-benzo-15-crown-5 ether (82)**

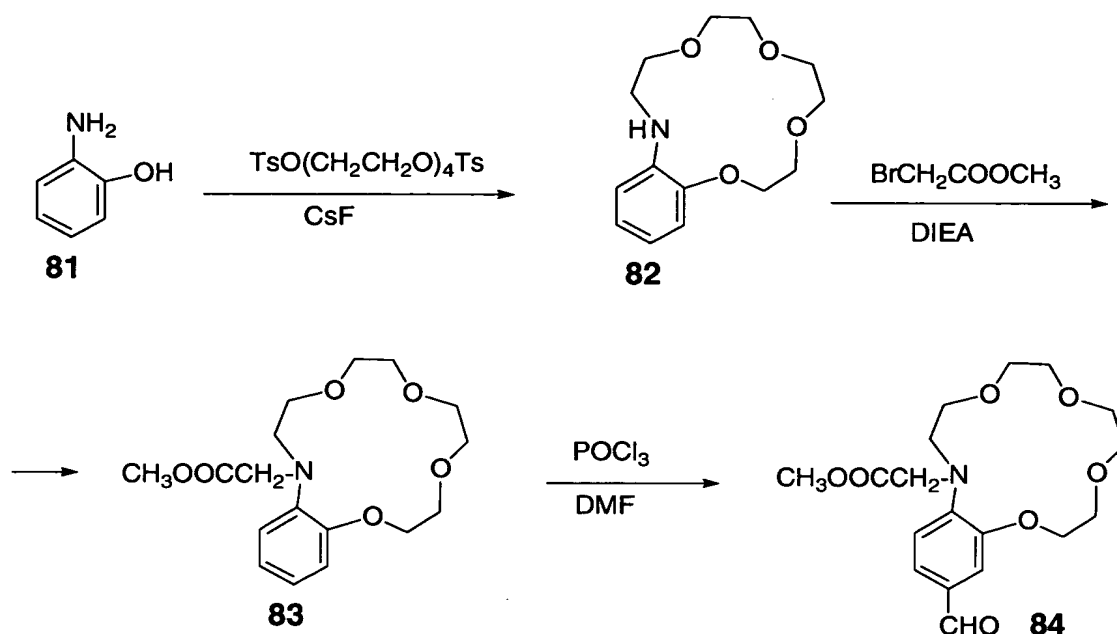
To a solution of 2-aminophenol **81** (1.322 g, 12 mmol) in MeCN (1 L), powdered CsF (7.300 g, 48 mmol) is added. The mixture is stirred vigorously for 1 h and then tetraethyleneglycole  
5 ditosylate (6.100 g, 12 mmol) in MeCN (50 mL) is introduced. The mixture is refluxed under N<sub>2</sub> atmosphere for 24h and evaporated. The residue is dissolved in CHCl<sub>3</sub> (800 mL), washed with H<sub>2</sub>O, sat. NaHCO<sub>3</sub>, H<sub>2</sub>O, sat. NaCl (200 mL each), filtered through paper, and evaporated. The crude product is purified by column chromatography on silica gel (12x60  
10 cm bed column, made in CHCl<sub>3</sub>) using CHCl<sub>3</sub> as eluant to give compound **82**, 1.995 g (62% yield) as a low melting solid.

**Example 76: Preparation of 1-Methoxycarbonylmethyl-1-aza-benzo-15-crown-5-ether (83).**

The mixture of compound **82** (1.380 g, 5.17 mmol), DIEA (2.4 mL, 25.84 mmol), methyl  
15 bromoacetate (1.8 mL, 10.34 mmol), and NaI (0.750 g, 5.00 mmol) in MeCN (100 mL) is refluxed under N<sub>2</sub> atmosphere for 24 h, then cooled and evaporated. The residue is re-dissolved in CHCl<sub>3</sub> (400 mL), washed with 1% AcOH (2x200 mL), H<sub>2</sub>O (200 mL). The chloroform solution is dried over MgSO<sub>4</sub>, filtered through paper, and evaporated. The crude  
20 product is purified by column chromatography on silica gel (3x30 cm bed column, made in CHCl<sub>3</sub>) using 0-2.5% MeOH gradient in CHCl<sub>3</sub> as eluant to give compound **83**, 1.102 g (63% yield) as a yellow oil.

**Example 77: Preparation of 15-Formyl-1-methoxycarbonylmethyl-1-aza-benzo-15-crown-5-ether (84).**

To a solution of the Vilsmeier reagent prepared from POCl<sub>3</sub> (1.0 mL, 11.0 mmol) in 5 mL  
DMF, compound **83** (0.750 g, 2.20 mmol) in DMF (2 mL) is introduced. The mixture is stirred under a N<sub>2</sub> atmosphere for 16 h, then poured into ice (20 g)/ sat. K<sub>2</sub>CO<sub>3</sub> (50 mL)  
25 mixture. The mixture is extracted with CHCl<sub>3</sub> (50 + 5x10 mL), and the extract is dried over MagSO<sub>4</sub> and evaporated. The crude product is purified by column chromatography on silica  
30 gel (1.5x30 cm bed column, made in CHCl<sub>3</sub>) using CHCl<sub>3</sub> as eluant to give aldehyde **84**, 0.632 g (78% yield) as an off-white low-melting solid.



**Example 78:** Preparation of compound **85** with a fluorinated xanthene as a DYE

5

A mixture of the aldehyde **84** (0.212 g, 0.58 mmol) and 4-fluororesorcinol (0.163 g, 1.27 mmol) in methanesulfonic acid (7 mL) is stirred for 24 h, then poured into 3N NaOAc (100 mL). The precipitated solid is filtered off, washed generously with water, and dried under vacuum to give compound **85** (0.229 g, 67%) as an off-white solid. Crude compound **85** is

10

**Example 79:** Preparation of a cell-impermeable metal chelating compound (Compound **86**) with a fluorinated xanthene as a DYE

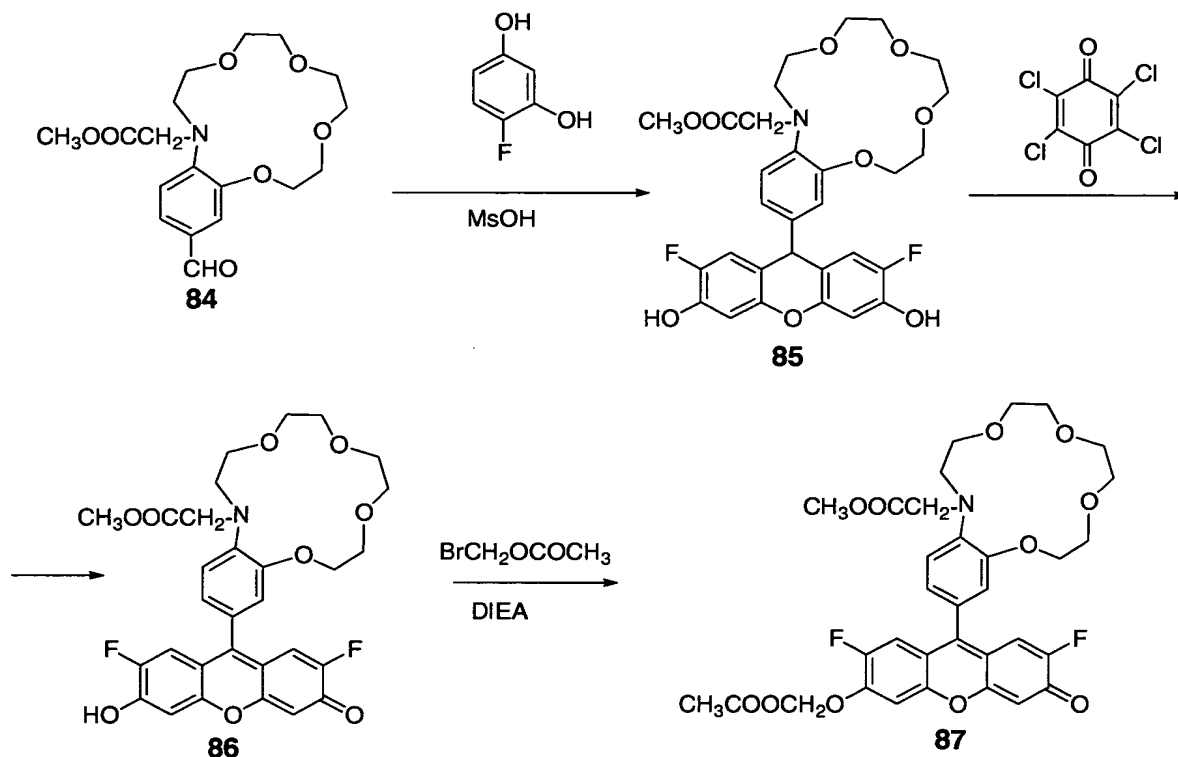
15 A mixture of the dihydro compound **85** (0.117 g, 0.20 mmol) and freshly powdered chloranil (0.246 g, 1.00 mmol) in  $\text{CHCl}_3/\text{MeOH}$  1:1 (10 mL) is vigorously stirred upon reflux for 3 h, then cooled down, filtered from the excess oxidizer and evaporated. The residue is purified by preparative TLC on silica gel, using  $\text{MeOH}/\text{AcOH}$  5%:2% in  $\text{CHCl}_3$  as eluant to give the compound **86** (0.046 g, 39%) as an orange solid.

20

**Example 80:** Preparation of a cell-permeable metal chelating compound (Compound **87**) with a fluorinated xanthene as a DYE.



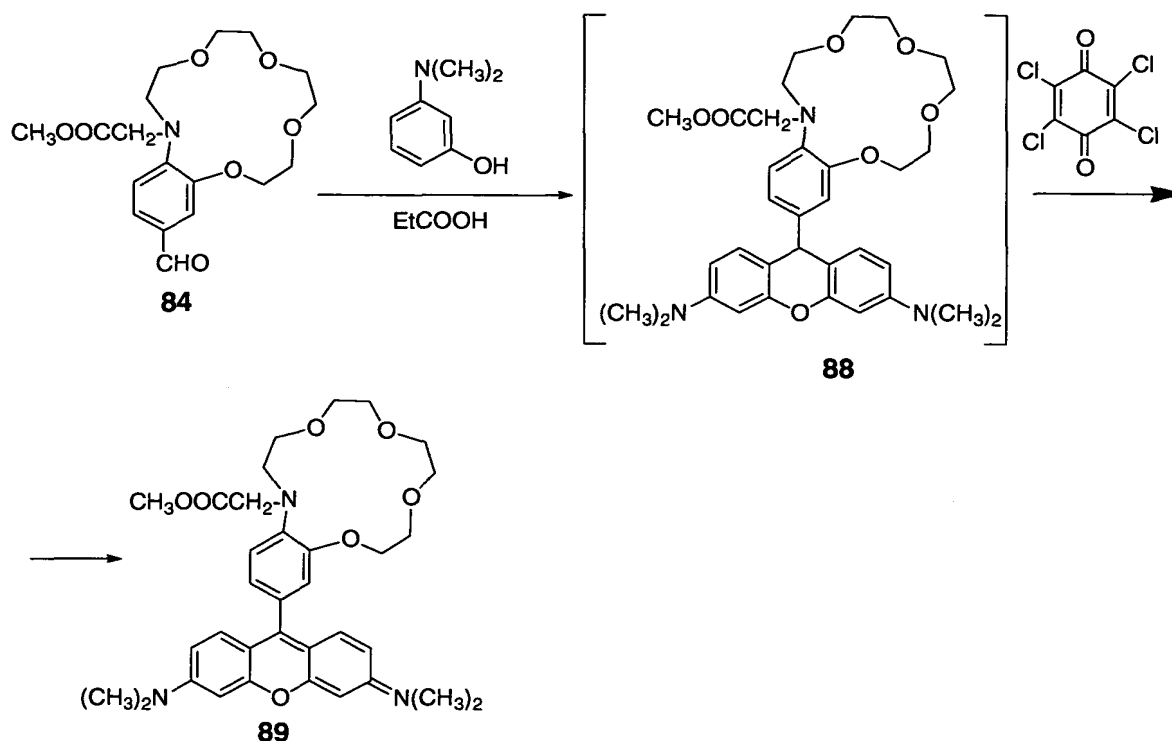
To a solution of compound **86** (12 mg, 0.02 mmol) and DIEA (75  $\mu$ L, 0.4 mmol) in DMF (1 mL), bromomethyl acetate (20  $\mu$ L, 0.2 mmol) is added. The mixture is stirred for 3 h and evaporated at 1 mm Hg vacuum. The residue is purified by preparative TLC on silica gel, using MeOH/AcOH 7%:1% in  $\text{CHCl}_3$  as eluant to give the compound **87** (12 mg, 96%) as an orange solid.



### Example 81: Preparation of metal chelating compound (Compound **89**) with Tetramethyrosamine as DYE

A mixture of aldehyde **844** (0.410 g, 1.11 mmol), and 3-(N,N-dimethylamino)phenol (0.337 g, 2.46 mmol), and TsOH (22 mg) in  $\text{EtCOOH}$  (11 mL) is stirred for 16 h at 65  $^{\circ}\text{C}$ . The mixture is cooled down and poured into 3N  $\text{NaOAc}$  (100 mL). The resulting suspension is extracted with  $\text{CHCl}_3$  (100 + 7x20 mL). The chloroform extract is filtered through paper and evaporated. The crude dihydro compound **88** is re-dissolved in MeOH/ $\text{CHCl}_3$  1:1 mixture (10 mL) and treated with chloranil (0.246 g, 1.00 mmol). The oxidation is continued upon vigorous stirring for 2 h, then the solvents are evaporated, the residue is re-dissolved in  $\text{CHCl}_3$ , and purified by column chromatography on silica gel (4 x 60 cm bed column, made with 5% MeOH + 1% AcOH in  $\text{CHCl}_3$ ) using (5-12.5%) MeOH / (1.0-1.3%) AcOH gradient as eluant to give the crude product, which is purified further by preparative TLC on silica gel,

using 12% MeOH/ 3% AcOH in CHCl<sub>3</sub> as eluant to give compound **89** (0.038 g, 6%) as a dark red solid.

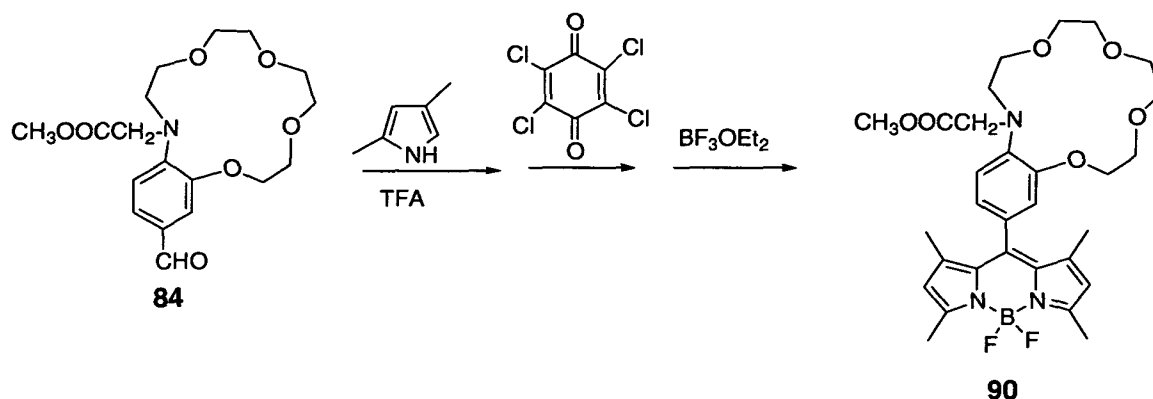


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**Example 82:** Preparation of a metal chelating compound (Compound **90**) with a borapolyazaindacene as a DYE

- 10 To a stirred solution of aldehyde **84** (0.367 g, 1.0 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (40 ml) 2,4-dimethylpyrrole (0.25 mL, 2.4 mmol) is added, followed by TFA (0.09 mL, 1.2 mmol). The mixture is stirred for 20 h, diluted with CHCl<sub>3</sub> (200 mL) and washed with 2% Me<sub>4</sub>NOH (2 x 200 mL), H<sub>2</sub>O (200 mL). The chloroform layer is separated, filtered through paper filter and evaporated. The residue is co-evaporated with toluene (50 mL), re-dissolved in toluene (40
- 15 mL) and stirred 2h with chloranil (0.296 g, 1.2 mmol). DIEA (1.7 mL, 10 mmol) is added, followed by BF<sub>3</sub>·OEt<sub>2</sub> (1.04 mL, 8 mmol) and the mixture is stirred for 20 h, filtered through Cellite, and evaporated. The residue is purified by column chromatography on silica gel (4x40 cm bed, made in 2% MeOH + 1% AcOH in CHCl<sub>3</sub>), using the same mixture as eluant to give compound **90** (0.169 g, 30%) as an orange solid.

20



**Example 83: Synthesis of an  $\alpha$ -nitrostilbene (Compound 92)**

5

A mixture of aldehyde **84** (0.367 g, 1.0 mmol), Wittig salt **91** (0.804 g, 1.5 mmol), and  $K_2CO_3$  (0.690 g, 5.0 mmol) in DMF (6 mL) is stirred for 7 h at 90 °C, then left overnight at rt. The mixture is poured into  $H_2O$  (200 mL), and the resulting suspension is extracted with  $CHCl_3$  (200 + 10x20 mL). The extract is evaporated to dryness at 3 mm Hg and the residue is

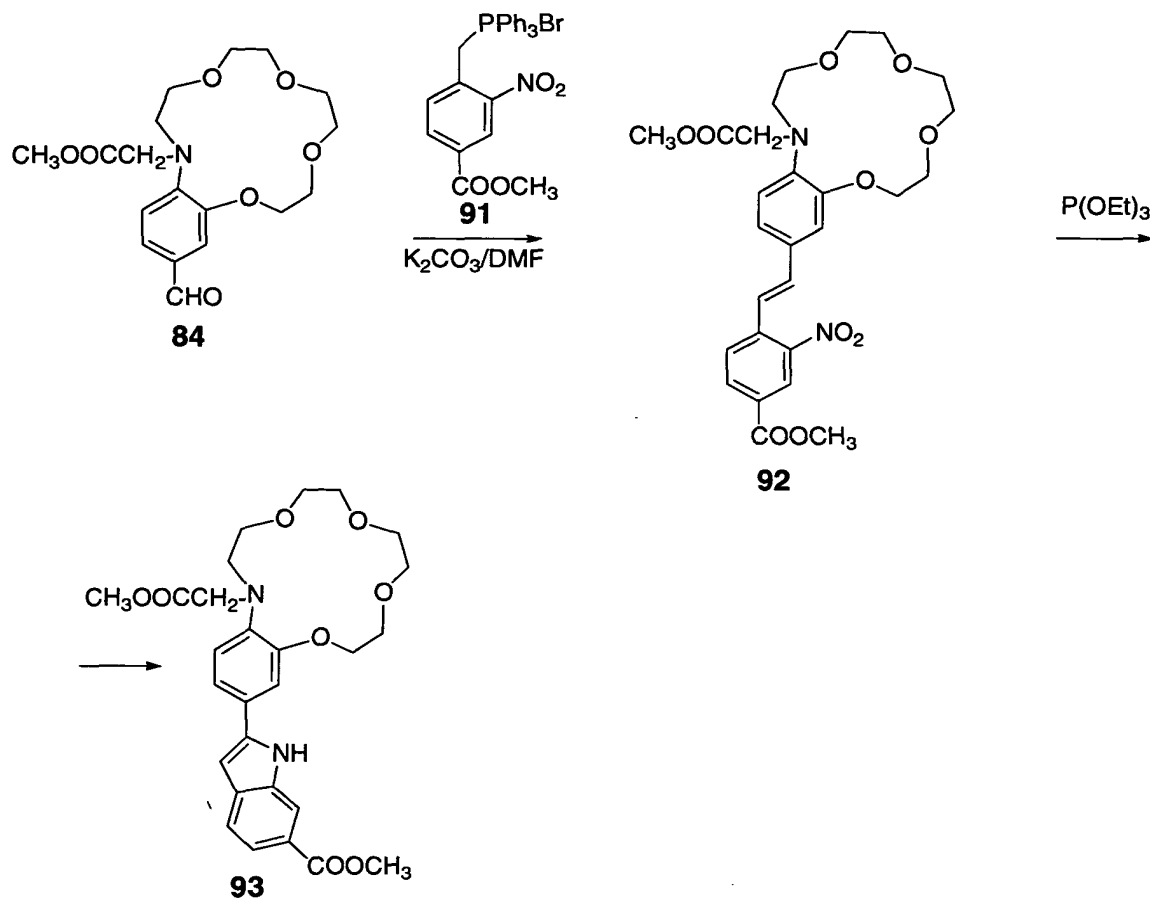
10 purified by column chromatography on silica gel (4 x 60 cm bed column, prepared in 4% MeOH in  $CHCl_3$ ) using the same mixture of solvents as eluant to give stilbene **92** (0.530g, 97%) as a dark red low-melting solid.

**Example 84: Preparation of a metal chelating compound (Compound 93) with indole as a**

15 DYE.

A solution of stilbene **92** (0.530 g, 0.97 mmol) in  $P(OEt)_3$  (5 mL) is heated at 125 °C for 4 h, then evaporated. The residue is purified by column chromatography on silica gel (4 x 50 cm bed column, prepared in 5% MeOH in  $CHCl_3$ ) using the same mixture of solvents as eluant

20 to give compound **93** (0.171 g, 34%).



**Example 85:** Preparation of 15-Nitro-1-methoxycarbonylmethyl-1-aza-benzo-15-crown-5-ether (Compound **94**).

5

To a stirred solution of crown ether **83** (0.420 g, 1.24 mmol) in  $Ac_2O$  (10 mL), 65%  $HNO_3$  (0.10 mL, 1.5 mmol) is introduced at  $0^\circ C$ . The mixture is stirred for 2 h at  $0^\circ C$  and then poured into sat.  $K_2CO_3$  (100 mL), and stirred for 1 h. The resulting solution is extracted with  $CHCl_3$  (100 + 7x20 mL), the extract is filtered through paper and evaporated. The residue is

10

**Example 86:** Preparation of Azo-dye metal chelating compound (Compound **96**)

15

A solution of crown ether **83** (0.339 g, 1.0 mmol) in dioxane (1 mL) and  $AcOH$  (1 mL) is treated with a solution of diazo compound **95**, prepared from sulfanilic acid (0.190 g, 1.1 mmol) in 6 mL  $H_2O$ . The mixture is stirred for 3 h, poured into  $H_2O$  (100 mL) and extracted with  $CHCl_3$  (100 + 25 x 20 mL). The chloroform extract is evaporated, and the residue is

purified by column chromatography on Sephadex LH-20 (3 x 35 cm bed column, prepared in H<sub>2</sub>O) using H<sub>2</sub>O as eluant to give azo compound **96** (0.029g, 7%) as a dark red solid.

**Example 87:** Preparation of 15-Amino-1-methoxycarbonylmethyl-1-aza-benzo-15-crown-5-ether (Compound **97**).

A) A sample of nitro compound **94** (0.220 g, 0.57 mmol) is hydrogenated at 50 psi in DMF (25 mL) over 10% Pc/C (40 mg, catalyst) for 7 h. The mixture is filtered, and evaporated at 1 mm Hg. The residue is purified by preparative TLC on silica gel, using 20% MeOH in CHCl<sub>3</sub> as eluant to give the amine **97** (128 mg, 63%).

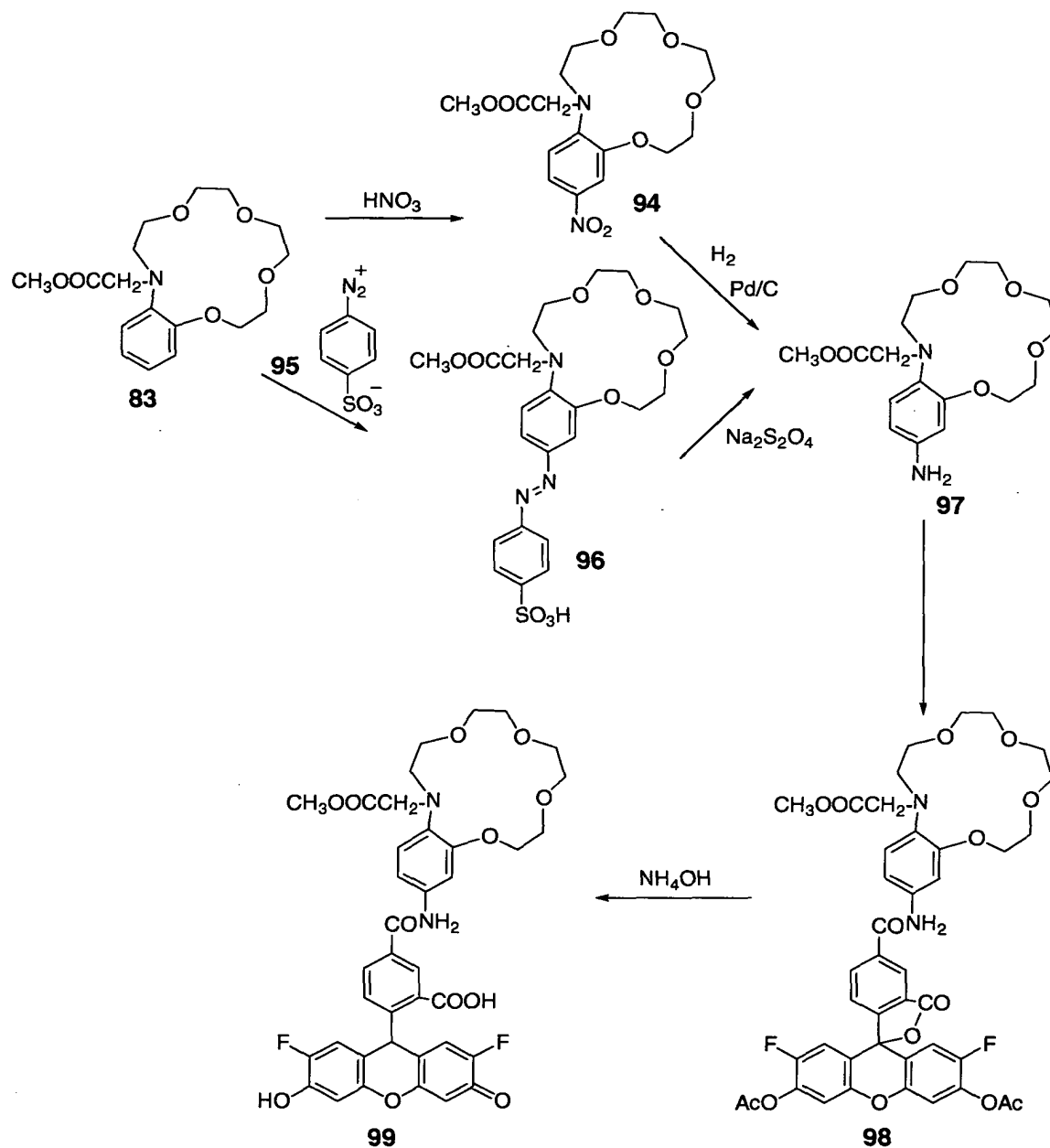
B) A sample of azo compound **96** (4 mg, 0.01 mmol) in 0.5 ml ethanol is treated with sodium dithionite (10 mg, 0.06 mmol) in 0.2 ml water. The solvents are evaporated and the residue is washed with water to give amine **97** (3mg, 85%), identical to that prepared in the above example.

**Example 88:** Preparation of cell-permeable Compound **98**.

To a stirred solution of the amine **97** (25 mg, 0.07 mmol) and DIEA (0.17 ml, 1 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (3 mL) an acid chloride, prepared from fluorinated xanthene acid (55 mg, 0.11 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (2 mL) is added dropwise. The mixture is stirred for 2 h and evaporated. The residue is purified by preparative TLC on silica gel, using 5% MeOH and 1% AcOH in CHCl<sub>3</sub> as eluant to give the compound **98** (32 mg, 63%) as an off-white solid.

**Example 89:** Preparation of cell-impermeable Compound **99**.

To a solution of the compound **98** (10 mg, 0.012 mmol) in MeOH (2 mL) an aqueous solution of NH<sub>3</sub> (0.5 mL) is added. The mixture is stirred for 30 min, evaporated to dryness, and the residue is suspended in water (2 mL), and centrifuged. The supernatant is discarded, and the solid is purified with preparative TLC on reverse-phase C18 plates, using 50% 2-PrOH and 0.2% TFA in H<sub>2</sub>O as eluant to give compound **99** (4 mg, 48%), as an orange solid.



**Example 90: Preparation of 1-Aza-benzo-14-crown-4 ether (Compound 100).**

- 5 To a solution of 2-aminophenol **81** (1.322 g, 12 mmol) in MeCN (1 L), powdered CsF (7.300 g, 48 mmol) is added. The mixture is stirred vigorously for 1 h and then triethyleneglycol ditosylate (5.503 g, 12 mmol) in MeCN (50 mL) is introduced. The mixture is refluxed under  $\text{N}_2$  atmosphere for 16 h and then evaporated. The residue is dissolved in  $\text{CHCl}_3$  (500 mL), washed with  $\text{H}_2\text{O}$ , sat.  $\text{NaHCO}_3$ ,  $\text{H}_2\text{O}$ , sat.  $\text{NaCl}$  (200 mL each), filtered through paper, and
- 10 evaporated. The residue is purified by column chromatography on silica gel (12x60 cm bed column, made in  $\text{CHCl}_3$ ) using  $\text{CHCl}_3$  as eluant to give crude product, which is purified by

column chromatography on silica gel (12x60 cm bed column, made in 25% EtOAc in hexanes) using the same mixture of solvents as eluant to give compound **100**, 0.370 g (14% yield) as an off-white solid.

5 **Example 91:** Preparation of 1-Methoxycarbonylmethyl-1-aza-benzo-12-crown-4-ether (Compound **101**).

10 The mixture of compound **100** (0.360 g, 1.61 mmol), DIEA (0.56 mL, 3.22 mmol), methyl bromoacetate (0.76 mL, 8.05 mmol), and NaI (0.241 g, 1.61 mmol; catalyst) in MeCN (30 mL) is refluxed under N<sub>2</sub> atmosphere for 16 h, then cooled down and evaporated. The residue is redissolved in CHCl<sub>3</sub> (200 mL), washed with 1% AcOH (2x200 mL), H<sub>2</sub>O (200 mL). The chloroform solution is dried over MgSO<sub>4</sub>, filtered through paper, and evaporated. The residue (TLC-pure crude product) is used in the next step without purification, considering it to be a 1.6 mmol quantity.

15

**Example 92:** Preparation of 13-Formyl-1-methoxycarbonylmethyl-1-aza-benzo-12-crown-4-ether (Compound **102**).

20 To a solution of the Vilsmeier reagent prepared from POCl<sub>3</sub> (0.75 mL, 8.0 mmol) in 8 mL DMF compound **101** (1.60 mmol) in DMF (2 mL) is introduced. The mixture is stirred under a N<sub>2</sub> atmosphere for 16 h, then poured into ice (20 g)/ sat. K<sub>2</sub>CO<sub>3</sub> (50 mL) mixture. The mixture is extracted with CHCl<sub>3</sub> (100 + 6x10 mL); the extract is dried over MgSO<sub>4</sub>, and evaporated. The crude product is purified by column chromatography on silica gel (1.5x30 cm bed column, made in CHCl<sub>3</sub>) using CHCl<sub>3</sub> as eluant to give aldehyde **102**, 0.285 g (55% yield on two steps) as an off-white solid.

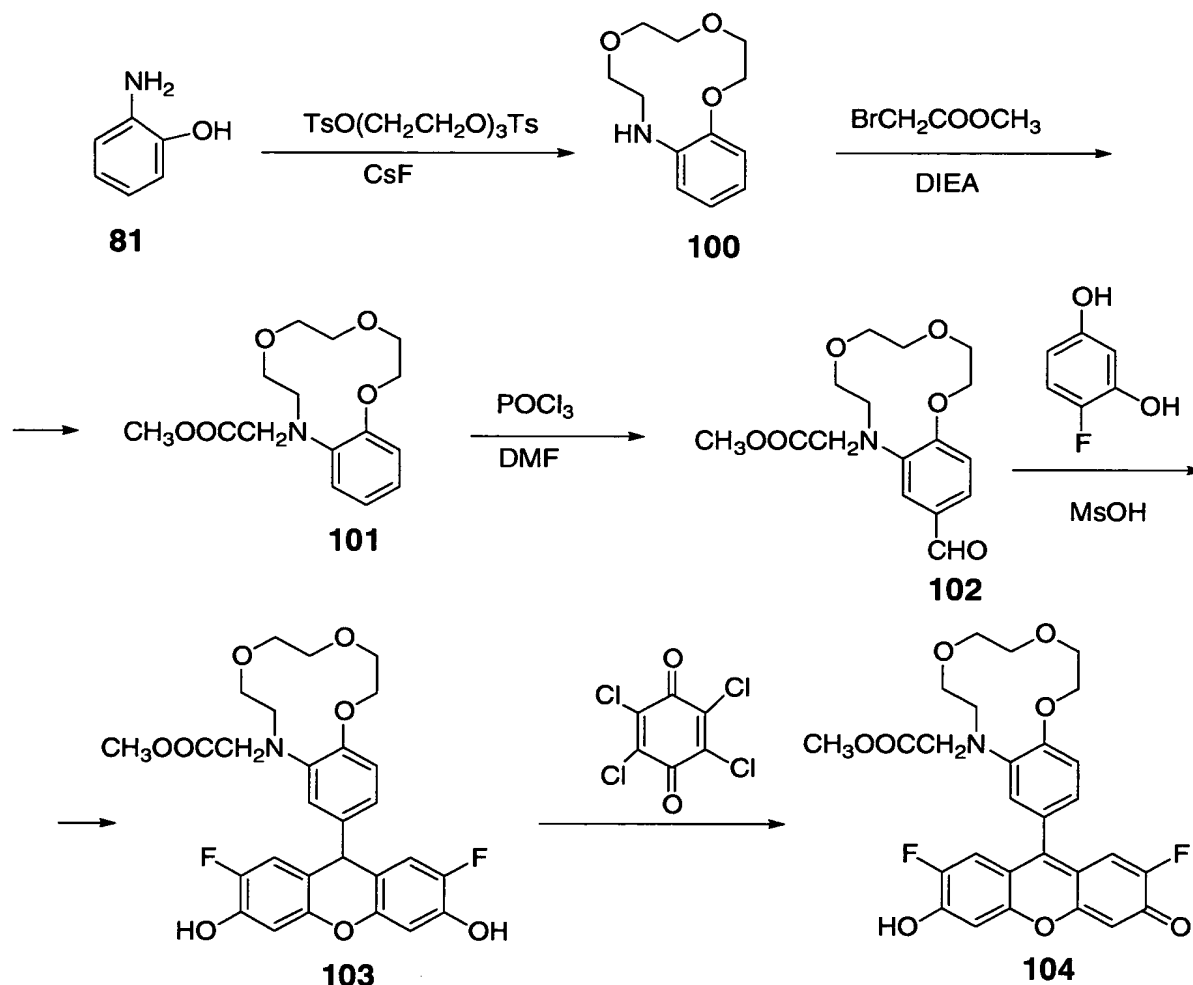
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**Example 93:** Preparation of compound **103**.

30 A mixture of the aldehyde **102** (0.272 g, 0.88 mmol) and 4-fluoro resorcinol (0.248 g, 1.94 mmol) in MeOH (13 mL) is stirred for 24 h and then poured into 3N NaOAc (100 mL). The mixture is centrifuged, and the supernatant is discarded. The solid is washed with water, and dried in vacuum to give compound **103** (0.087 g, 18%) as an off-white solid. Crude compound **104** is used in the next step without purification.

35 **Example 94:** Preparation of cell-impermeable compound **104**.

A mixture of the dihydro compound **85** (0.085 g, 0.16 mmol), and freshly powdered chloranil (0.197 g, 0.80 mmol) in CHCl<sub>3</sub>/MeOH 1:1 mixture (6 mL) is vigorously stirred during reflux for 4h, then cooled down, filtered from the excess oxidizer and evaporated. The residue is purified by preparative TLC on silica gel, using MeOH/AcOH 5%:2% in CHCl<sub>3</sub> as eluant to give the compound **104** (0.031 g, 36%) as an orange solid.



**Example 95: Preparation of 1-Aza-benzo-18-crown-6 ether (Compound **105**).**

To a solution of 2-aminophenol **81** (1.090 g, 10 mmol) in MeCN (1 L), powdered CsF (6.080 g, 40 mmol) is added. The mixture is stirred vigorously for 1 h and then pentaethyleneglycol ditosylate (5.500 g, 11 mmol) in MeCN (50 mL) is introduced. The mixture is refluxed under N<sub>2</sub> atmosphere for 70 h and then evaporated. The residue is dissolved in CHCl<sub>3</sub> (600 mL), washed with H<sub>2</sub>O, sat. NaHCO<sub>3</sub>, H<sub>2</sub>O, and sat. NaCl (200 mL each). The chloroform solution is dried over MgSO<sub>4</sub>, filtered through paper, and evaporated. The residue is purified by column chromatography on silica gel (8x45 cm bed column, made in CHCl<sub>3</sub>) using 0-1.2%



MeOH gradient in  $\text{CHCl}_3$  as eluant to give compound **105**, 0.681 g (22% yield) as a off-white solid.

**Example 96:** Preparation of 1-Methoxycarbonylmethyl-1-aza-benzo-18-crown-6-ether (Compound **126**).

A mixture of compound **105** (0.670 g, 2.15 mmol), DIEA (0.72 mL, 4.15 mmol), methyl bromoacetate (1.01 mL, 10.75 mmol), and NaI (0.322 g, 2.15 mmol; catalyst) in MeCN (40 mL) is refluxed under a  $\text{N}_2$  atmosphere for 16 h, then cooled down and evaporated. The residue is redissolved in  $\text{CHCl}_3$  (200 mL), washed with 1% AcOH (2x200 mL),  $\text{H}_2\text{O}$  (200 mL). Chloroform solution is dried over  $\text{MgSO}_4$ , filtered through paper, and evaporated. The residue is purified by column chromatography on silica gel (3x30 cm bed column, made in  $\text{CHCl}_3$ ) using 2.5 - 4% MeOH gradient in  $\text{CHCl}_3$  as eluant to give compound **106**, 0.430 g (52% yield) as an off-white solid.

**Example 97:** Preparation of 18-Formyl-1-methoxycarbonylmethyl-1-aza-benzo-18-crown-6-ether (Compound **107**).

To a solution of the Vilsmeier reagent prepared from  $\text{POCl}_3$  (1.23 mL, 13.2 mmol) in 6 mL DMF compound **106** (0.410g, 1.32 mmol) in DMF (2 mL) is introduced. The mixture is stirred under a  $\text{N}_2$  atmosphere for 16 h, then poured into ice (40 g)/ sat.  $\text{K}_2\text{CO}_3$  (100 mL) mixture. The mixture is extracted with  $\text{CHCl}_3$  (5x100 mL), the extract is dried over  $\text{MgSO}_4$ , and evaporated. The crude product is treated with cold ether (10 mL), and the precipitated solid is collected to give aldehyde **107**, 0.246 g (45% yield) as a white solid.

**Example 98:** Preparation of compound **108**.

A mixture of the aldehyde **107** (0.236 g, 0.57 mmol) and 4-fluoro resorcinol (0.186 g, 1.45 mmol) in MsOH (8 mL) is stirred for 16 h, then poured into 3N NaOAc (100 mL). The precipitate is filtered, washed with water, and dried in vacuum to give compound **108** (0.350 g, 97%) as an off-white solid. Crude compound **108** is used in the next step without purification.

**Example 99:** Preparation of Cell-impermeable compound **109**.

A mixture of the dihydro compound **108** (0.350 g, 0.56 mmol), and freshly powdered chloranil (0.701 g, 2.85 mmol) in  $\text{CHCl}_3$ /MeOH 1:1 mixture (25 mL) is vigorously stirred upon

reflux for 4 h, then cooled down, filtered from the excess oxidizer and evaporated. The residue is purified by column chromatography on silica gel, (4 x 40 cm bed column, made in 10% MeOH + 1% AcOH in CHCl<sub>3</sub>) using the same mixture of solvents as eluant to give the compound **109** (0.088 g, 25%) as an orange solid.

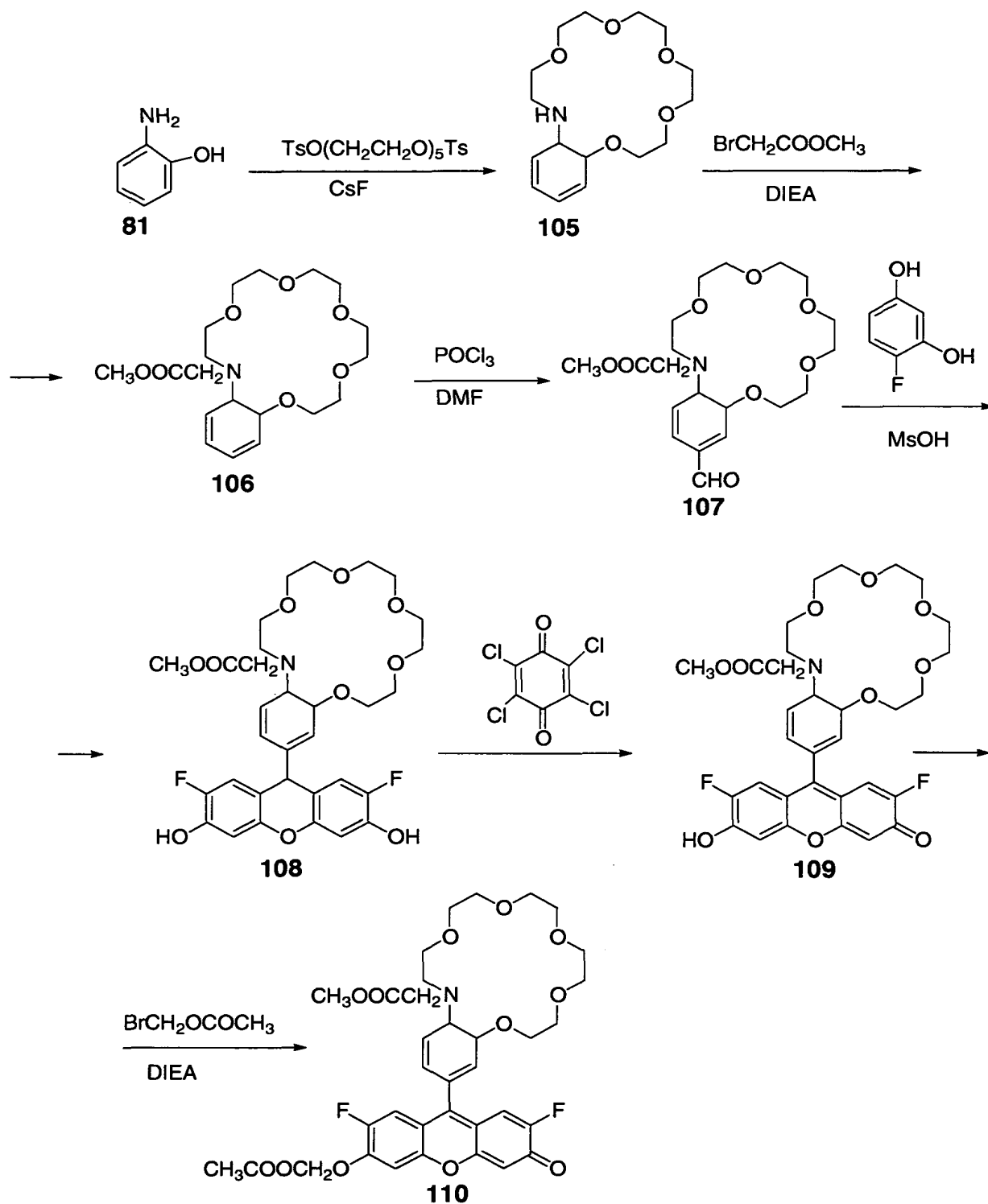
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**Example 100: Preparation of Cell-permeable Compound 110.**

To a solution of compound **109** (32 mg, 0.05 mmol) and DIEA (0.17 mL, 1.0 mmol) in DMF (2 mL), bromomethyl acetate (50  $\mu$ L, 0.5 mmol) is added. The mixture is stirred for 3 h and  
10 evaporated at 1 mm Hg vacuum. The residue is purified by preparative TLC on silica gel, using MeOH/AcOH 15%:1% in CHCl<sub>3</sub> as eluant to give the compound **110** (20 mg, 57%) as an orange solid.

15

20



**Example 101:** Preparation of 1-(2'-Methoxyacetyl)-1-aza-benzo-15-crown-5 ether  
5 (Compound 111).

To a stirred solution of the crown ether **82** (6.20 g, 23 mmol) and Et<sub>3</sub>N (9.3 mL, 70 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (250 mL) a methoxyacetyl chloride (3.2 mL, 35 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (15 mL) is introduced within 15 min. The mixture is stirred for 3 h, diluted with CHCl<sub>3</sub> (150 mL), and washed with  
5 H<sub>2</sub>O, 1% AcOH, H<sub>2</sub>O, sat. NaHCO<sub>3</sub>, H<sub>2</sub>O (200 mL each), filtered through paper filter and evaporated. The residue is purified by column chromatography on silica gel (6 x 45 cm bed column, made in CHCl<sub>3</sub>), using 2.5% MeOH in CHCl<sub>3</sub> as eluant to give compound **111** (4.25 g, 55%) as a yellow oil.

10 **Example 102:** Preparation of 1-(2'-Methoxyethyl)-1-aza-benzo-15-crown-5 ether (Compound **112**).

A) To a stirred solution of compound **111** (4.25 g, 12.5 mmol) in dry THF (50 mL) a 1N diborane in THF (50 mL, 50 mmol) is added. The mixture is stirred under reflux for 20 h,  
15 decomposed by dropwise addition of MeOH (50 mL), and evaporated. The residue is co-evaporated with MeOH (5 x 100 mL) to remove boron complexes, and dried in vacuo to give compound **112** (4.05 g, 99%) as a colorless oil.

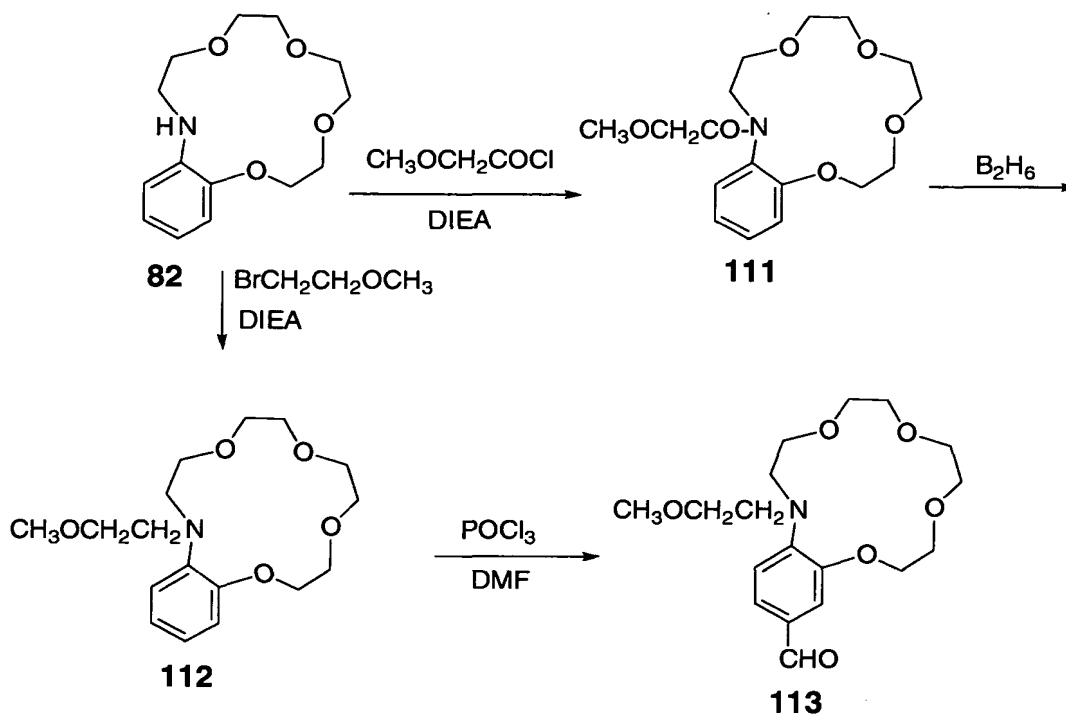
B) A mixture of compound **82** (1.20g, 4.50 mmol), DEIA (1.6 mL, 9.0 mmol), 2-  
20 bromoethylmethyl ether (4.2 mL, 45 mmol), NaI (0.68 g, 4.50 mmol, catalyst) in MeCN (200 mL) is stirred under reflux for 70h. More DIEA (1.6 mL, 9.0 mmol) and 2-bromoethylmethyl ether (4.2 mL, 45 mmol) are added and heating continued for another 40 h. The reaction mixture is evaporated, the residue is re-dissolved in CHCl<sub>3</sub> (300 mL), washed with 1% AcOH (3x200 mL), H<sub>2</sub>O (200 mL), filtered through paper filter and evaporated. The residue is  
25 purified by column chromatography on silica gel (3 x 40 cm bed column, made in 2.5% MeOH in CHCl<sub>3</sub>), using the same mixture of solvents as eluant to give compound **112** (0.080 g, 5%), identical to that prepared section A of this example.

30 **Example 103:** Preparation of 1-(2'-Methoxyethyl)-15-formyl-1-aza-benzo-15-crown-5 ether (Compound **113**).

To a solution of the Vilsmeier reagent prepared from POCl<sub>3</sub> (12 mL, 125 mmol) in 70 mL DMF compound **112** (4.05 g, 12.5 mmol) in DMF (10 mL) is introduced. The mixture is stirred under N<sub>2</sub> atmosphere for 16 h at 40 °C, then more Vilsmeier reagent from POCl<sub>3</sub> (12  
35 mL, 125 mmol) in 70 mL DMF is added and the stirring continued for another 24 h. The mixture is poured into ice (400 g)/ sat. K<sub>2</sub>CO<sub>3</sub> (400 mL) mixture. The mixture is extracted with CHCl<sub>3</sub> (300 + 6x50 mL), the extract is dried over MgSO<sub>4</sub>, and evaporated. The crude

product is purified by column chromatography on silica gel (6x50 cm bed column, made in  $\text{CHCl}_3$ ) using 1.5% MeOH in  $\text{CHCl}_3$  as eluant to give aldehyde **113**, 1.613 g (36% or, 58% based on recovery), then eluant is changed to 10% to recover starting material (1.49 g, 37%).

5



**Example 104: Preparation of compound 114.**

- 10 A mixture of the aldehyde **113** (0.700 g, 1.98 mmol) and 4-fluororesorcinol (0.558 g, 4.36 mmol) in methanesulfonic acid (30 mL) is stirred for 24 h and poured into 3N NaOAc (300 mL). The mixture is extracted with n-BuOH (7 x 50 mL). The extract is evaporated to give the crude dihydro derivative **114**. Crude compound **114** is used in the next step without purification, considering it as 1.8 mmol.

15

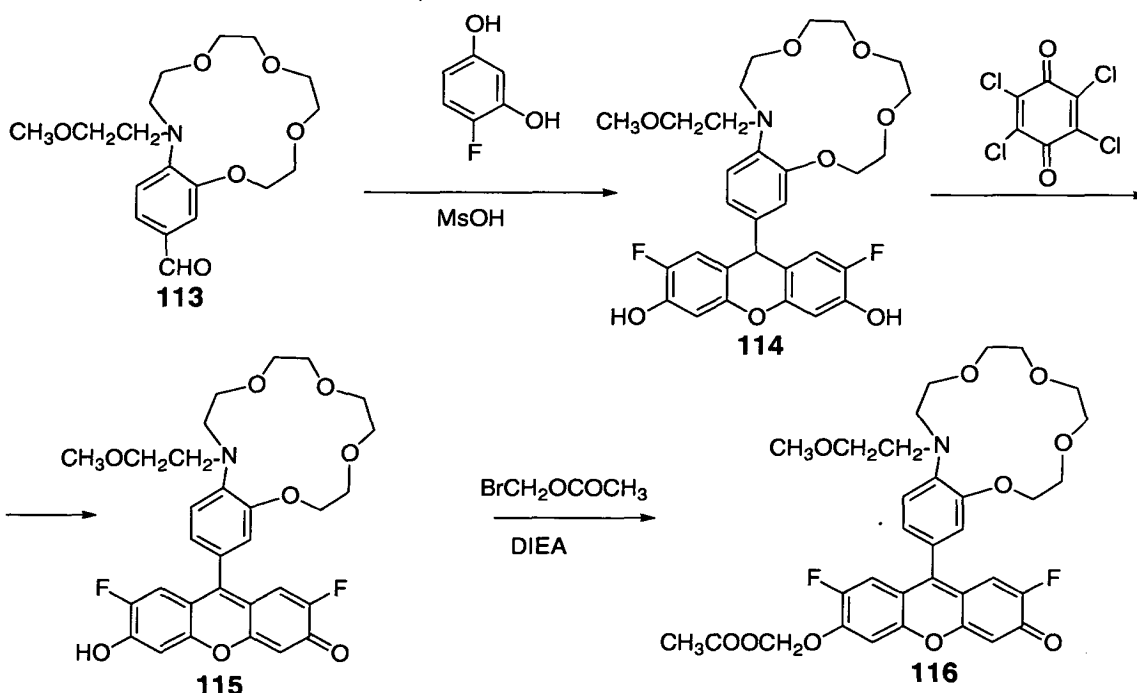
**Example 105: Preparation of Cell-impermeable Compound 115.**

- A mixture of the dihydro compound **114** (1.8 mmol), and freshly powdered chloranil (2.210 g, 9.00 mmol) in  $\text{CHCl}_3$ /MeOH 1:1 mixture (100 mL) is vigorously stirred upon reflux for 4 h, then cooled down, filtered from the excess oxidizer and evaporated. The residue is purified by column chromatography on silica gel (6 x 55 cm bed column, made in 10% MeOH + 1%
- 20

AcOH in CHCl<sub>3</sub>), using 10-15% MeOH gradient in CHCl<sub>3</sub> + 1% AcOH as eluant to give the compound **115** (0.189 g, 18% for two steps) as a brown solid.

**Example 106: Preparation of Cell-permeable Compound 116.**

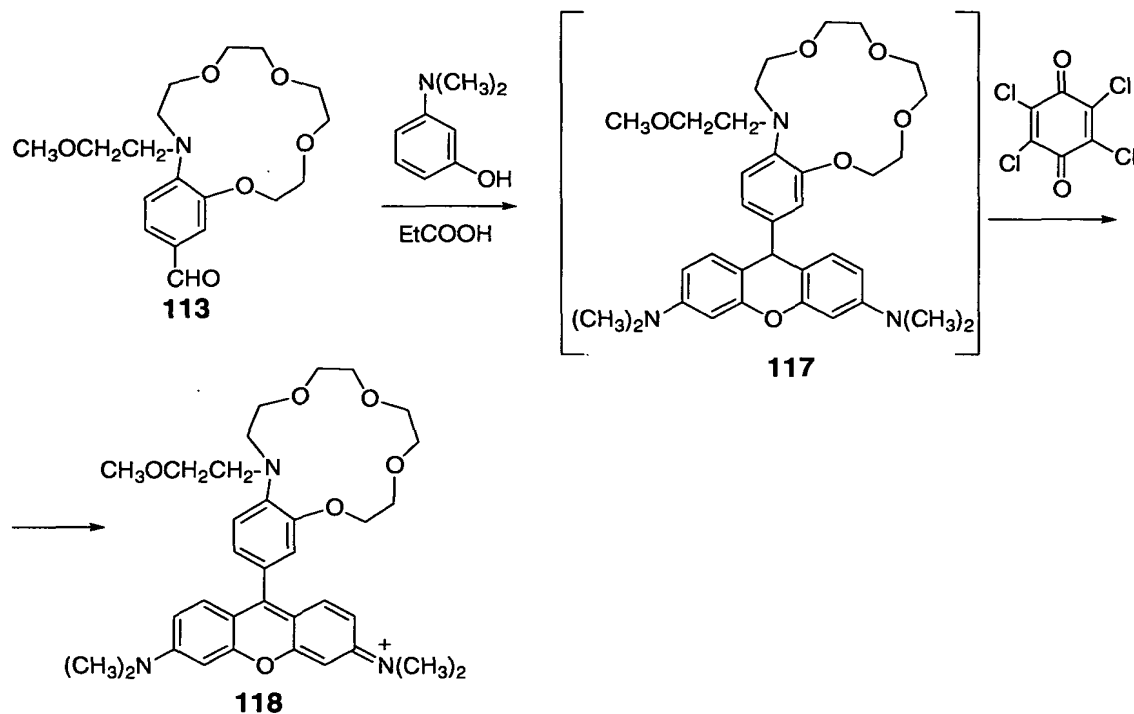
To a solution of compound **115** (28 mg, 0.05 mmol) and DIEA (0.17, 1.0 mmol) in DMF (2 mL), bromomethyl acetate (50  $\mu$ L, 0.5 mmol) is added. The mixture is stirred for 1 h and evaporated at 1 mm Hg vacuum. The residue is purified by preparative TLC on silica gel, using MeOH/AcOH 10%:2% in CHCl<sub>3</sub> as eluant to give the compound **116** (4 mg, 12%) as an orange solid.



**Example 107: Preparation of Tetramethyrosamine compound 118.**

A mixture of aldehyde **113** (0.900 g, 2.55 mmol), and 3-(N,N-dimethylamino)phenol (0.769 g, 5.61 mmol), and TsOH (50 mg, catalyst) in EtCOOH (25 mL) is stirred for 16 h at 65 °C. The mixture is cooled down and poured into 3N NaOAc (500 mL). The resulting suspension is extracted with CHCl<sub>3</sub> (200 + 7x30 mL). The extract is filtered through paper and evaporated. The resulting crude dihydro compound **117** is re-dissolved in MeOH/CHCl<sub>3</sub> 1:1 mixture (100 mL) and treated with chloranil (0.541 g, 2.20 mmol). The oxidation is continued upon vigorous stirring for 2 h, then the solvents are evaporated, the residue is re-dissolved in

$\text{CHCl}_3$ , and purified by column chromatography on silica gel (8 x 50 cm bed column, made with 5% MeOH + 1% AcOH in  $\text{CHCl}_3$ ) using 5-20% MeOH gradient in  $\text{CHCl}_3$ +1.0% AcOH as eluant to give compound **118** (0.236 g, 16% on two steps) as a dark red solid.



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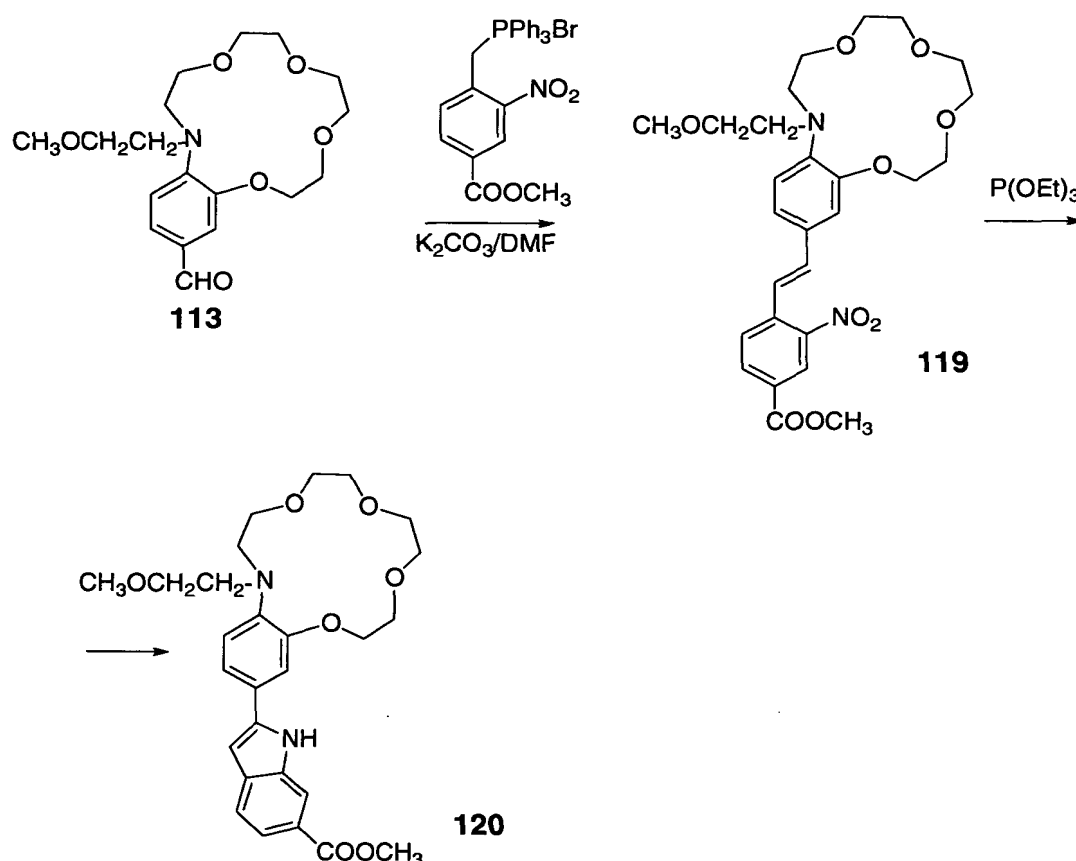
#### Example 108: Synthesis of an $\alpha$ -nitrostilbene (Compound **119**)

- 10 A mixture of aldehyde **113** (0.172 g, 0.49 mmol), Wittig salt **91** (0.392 g, 0.73 mmol), and  $\text{K}_2\text{CO}_3$  (0.338 g, 2.45 mmol) in DMF (3 mL) is stirred for 6 h at 90 °C, then left overnight at rt. The mixture is poured into  $\text{H}_2\text{O}$  (200 mL), and the resulting suspension is extracted with  $\text{CHCl}_3$  (50 + 7x20 mL). The extract is evaporated to dryness at 3 mm Hg and the residue is purified by column chromatography on silica gel (2.5 x 30 cm bed column, prepared in 3% MeOH in  $\text{CHCl}_3$ ) using the same mixture of solvents as eluant to give stilbene **119** (0.195g, 75%) as a dark red low-melting solid.
- 15

#### Example 109: Preparation of Compound **120**.

- 20 A solution of stilbene **119** (0.190 g, 0.36 mmol) in  $\text{P}(\text{OEt})_3$  (3 mL) is heated at 125 °C for 4 h, then evaporated. The residue is purified by column chromatography on silica gel (2.5 x 50

cm bed column, prepared in 5% MeOH and 1% AcOH in  $\text{CHCl}_3$ ) using the same mixture of solvents as eluant to give compound **120** (0.099 g, 55%) as an off-white solid.



5

**Example 110:** Preparation of 1-Ethoxyoxalyl-1-aza-benzo-15-crown-5-ether (Compound **121**).

- 10 To a mixture of compound **82** (1.345 g, 5.00 mmol) and Et<sub>3</sub>N (1.4 mL, 10.00 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (30 mL), ethyl oxalylchloride (0.84 mL, 7.50 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (5 mL) is added dropwise. The mixture is stirred for 1 h, diluted with CHCl<sub>3</sub> (200 mL), washed with 1% AcOH (3 x 100 mL), H<sub>2</sub>O (100 mL), sat. NaHCO<sub>3</sub> (2 x 100 mL), sat. NaCl (200 mL). The chloroform fraction is dried over MgSO<sub>4</sub> and evaporated. The residue is purified by column chromatography on silica gel (4 x 30 cm bed, made in 30% EtOAc in hexanes), using 30-60% EtOAc gradient in hexanes as eluant to give compound **121** (1.170 g, 65% yield) as a yellow oil.
- 15

**Example 111:** Preparation of 1-(N,N-Dimethylaminooxalyl)-1-aza-benzo-15-crown-5-ether (Compound **122**).



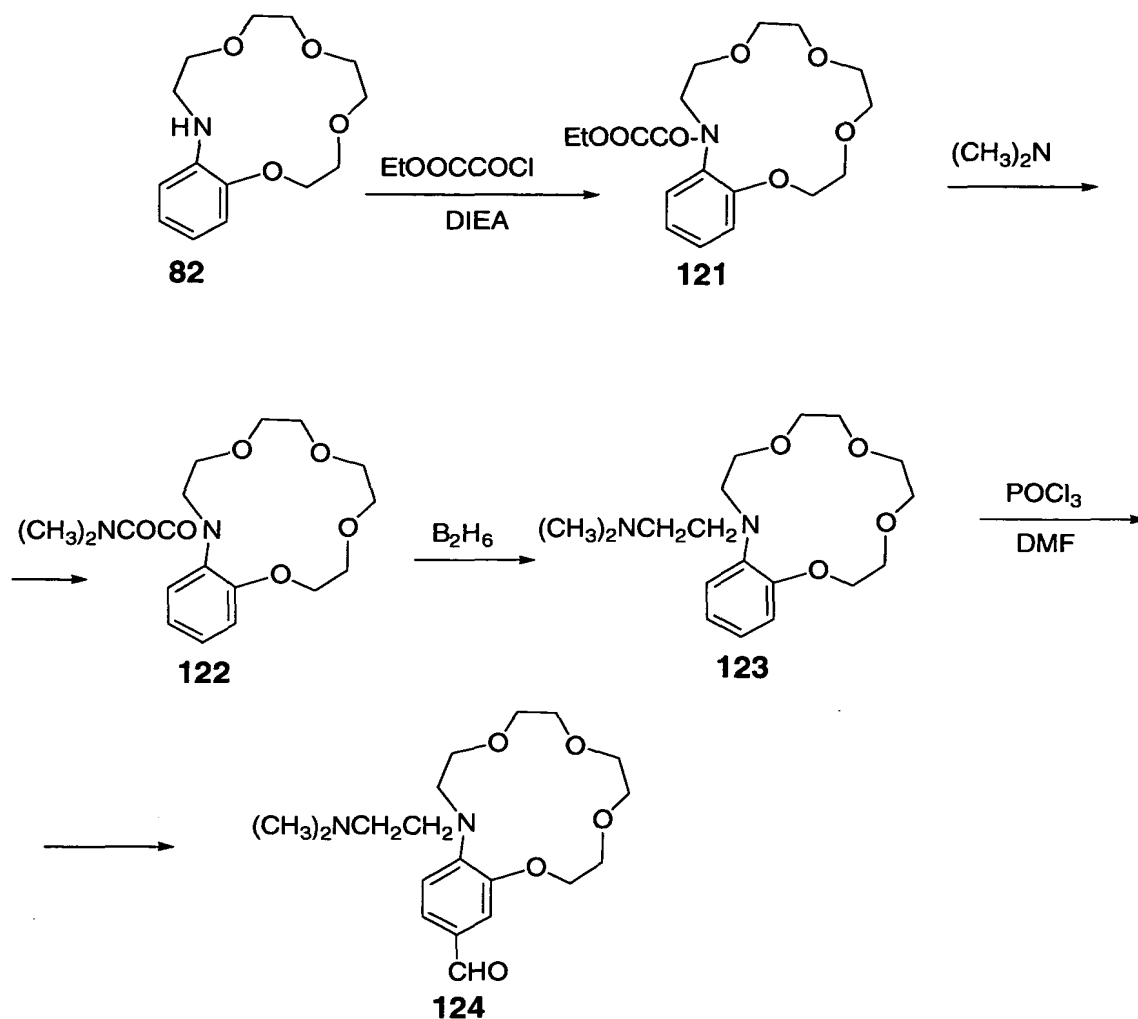
To a solution of the compound **121** (0.734 g, 2.00 mmol) in MeOH (15 mL), a solution of dimethylamine (2 mL, 40 mmol) in cold MeOH (15 mL) is added upon cooling to 0 °C. The mixture is heated to 30 °C for 16 h, then cooled to 0 °C and a new portion of dimethylamine (2 mL, 40 mmol) is introduced. The mixture is stirred for 3 h at 30 °C and then evaporated and the residue is treated with cold ether. The precipitated solid is filtered off, and washed with ether to give compound **122** (0.542 g, 74% yield) as a white solid.

**Example 112:** Preparation of 1-(2'-N,N-Dimethylaminoethyl)-1-aza-benzo-15-crown-5-ether (Compound **123**).

To the solution of compound **122** (0.440 g, 1.20 mmol) in THF (10 ml) a solution of 1N diborane in THF (12 mL, 12 mmol) is added. The mixture is refluxed for 16 h, then decomposed by careful addition of MeOH (30 mL), evaporated, and co-evaporated with MeOH (5 x 50 mL) to remove boron complexes. The crystalline residue of the compound **123** (0.405 g, 100% yield) is used in next step without purification.

**Example 113:** Preparation of 1-(2'-N,N-Dimethylaminoethyl)-15-formyl-1-aza-benzo-15-crown-5-ether (Compound **124**).

To a solution of the Vilsmeier reagent prepared from POCl<sub>3</sub> (1.5 mL, 16.0 mmol) in 5 mL DMF, compound **123** (0.550 g, 1.62 mmol) in DMF (2 mL) is introduced. The mixture is stirred under N<sub>2</sub> atmosphere for 16 h, then poured into ice (20 g)/ sat. K<sub>2</sub>CO<sub>3</sub> (50 mL) mixture. The mixture is extracted with CHCl<sub>3</sub> (100 + 7x25 mL), the extract is dried over MgSO<sub>4</sub>, and evaporated. The crude product is purified by column chromatography on silica gel (3x35 cm bed column, made with 3% MeOH in CHCl<sub>3</sub>) using the same mixture of solvents as eluant to give aldehyde **124**, (0.175 g 30% yield) as a yellow low-melting solid.



**Example 114:** Synthesis of  $\alpha$ -nitrostilbene (Compound **125**).

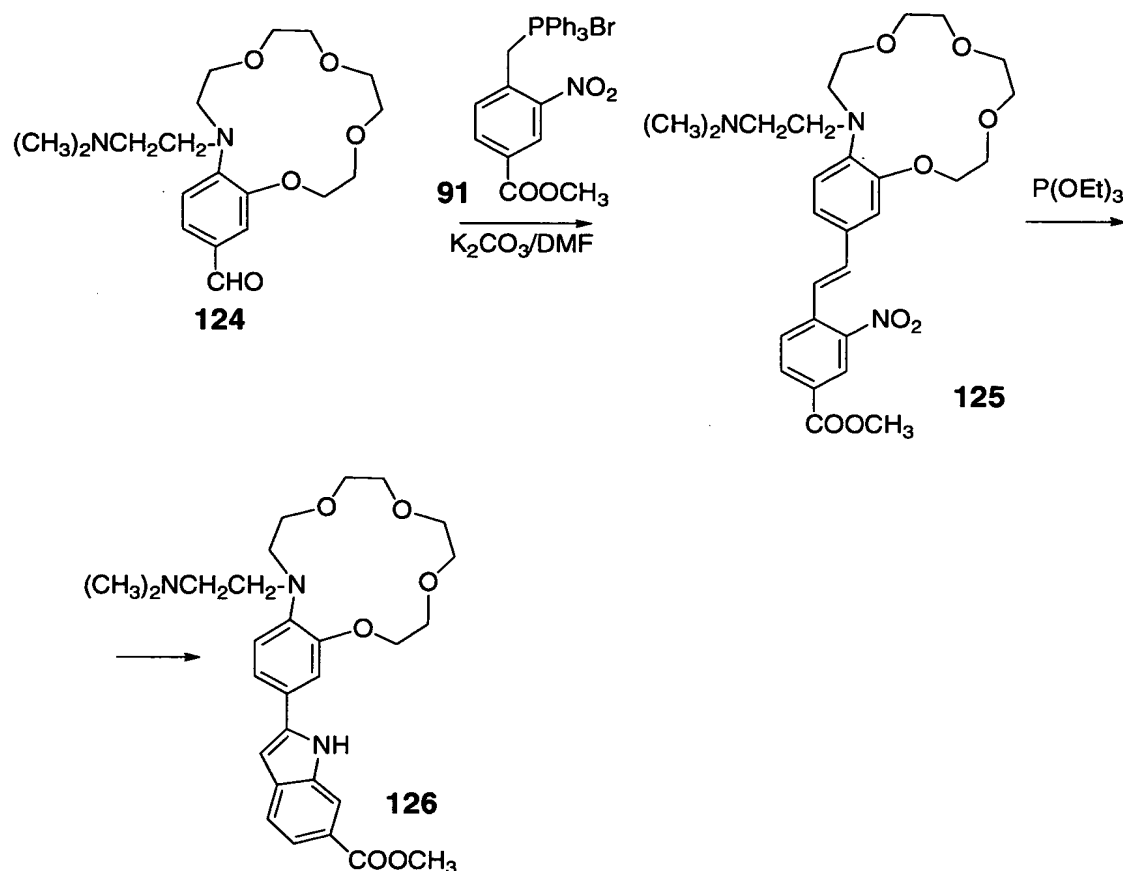
- 5 A mixture of aldehyde **124** (0.070 g, 0.19 mmol), Wittig salt **91** (0.154 g, 0.29 mmol), and  $\text{K}_2\text{CO}_3$  (0.131 g, 0.95 mmol) in  $\text{DMF}$  (2 mL) is stirred for 16 h at  $90^\circ\text{C}$ . The mixture is diluted with  $\text{CHCl}_3$  (10 mL), filtered from inorganic materials and evaporated to dryness at 3 mm Hg. The residue is purified by preparative TLC on silica gel using 20% MeOH and 5% AcOH in  $\text{CHCl}_3$  as eluant to give stilbene **125** (0.060 g, 58%) as a dark red solid.

10

**Example 115:** Preparation of Compound **126**

A solution of stilbene **119** (0.055 g, 0.01 mmol) in  $\text{P}(\text{OEt})_3$  (2 mL) is heated at  $125^\circ\text{C}$  for 14 h, then evaporated. The residue is purified by preparative TLC on silica gel using 25%

MeOH and 5% AcOH in  $\text{CHCl}_3$  as eluant to give compound **126** (0.029 g, 57%) as an off-white solid.

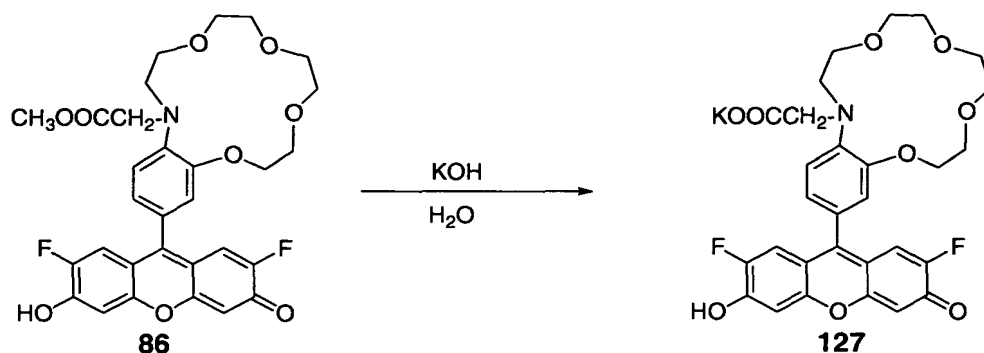


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#### Example 116: Hydrolysis of compound **86**.

To a solution of the compound **86** (0.032 g, 0.055 mmol) in 50% MeOH (2 mL), 1N KOH (0.16 mL, 16 mmol) is added. The mixture is stirred for 16 h, diluted with  $\text{H}_2\text{O}$ , and 0.2 N HCl is added to achieve pH=8.0. The resulting solution is filtered through a nylon membrane filter and evaporated. The crude product is purified by column chromatography on Sephadex LH-20 (2.6 x 50 cm bed column, made with  $\text{H}_2\text{O}$ ), using  $\text{H}_2\text{O}$  as eluant to give compound **127** (0.021 mg, 59 % yield) as a brown solid.

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### Example 117: Preparation of Compound 138

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To a solution of compound **128** in dry THF (0.5M) and two equivalents of DIEA is added dropwise benzoyl chloride (1.0 eq). After 1h at room temperature, the volatiles are removed. The residue is partitioned between 0.1M HCl and ethyl acetate. The organic layer is washed with brine, dried over sodium sulfate, and concentrated to give amide **129**.

10

To a solution of **129** in THF (0.5M) is added a solution of borane-THF (1.0M, 2 eq) at room temperature. The resulting solution is stirred at room temperature over night and then methanol (2 eq) is carefully added. The volatiles are removed in vacuo, and the residue partitioned between water and ethyl acetate. The organic layer is dried over sodium sulfate and concentrated to give aniline **130**, which is purified further if necessary by flash chromatography on silica gel using ethyl acetate in hexanes.

15

To a solution of **130** (0.05M) in dry THF with 2 eq DIEA is slowly added a 0.1M solution of diglycolyl chloride in dry THF with rapid stirring. After 2h, the reaction solution is concentrated in vacuo. The residue is partitioned between 0.1M HCl and ethyl acetate. The organic layer is washed with brine, dried over sodium sulfate, and concentrated to give bisamide **131**.

20

To a 0.1M solution of **131** in dry THF is slowly added borane-THF (1.0M, 5 eq). The resulting solution is stirred at room temperature overnight and then methanol (5 eq) is added. The resulting mixture is concentrated in vacuo, and the residue partitioned between water and ethyl acetate. The organic phase is washed with brine and dried over sodium sulfate, and concentrated in vacuo to give the diol **132**, which is purified further if necessary by flash chromatography on silica gel using methanol in chloroform.

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30

A solution of **132** is dissolved in dry acetonitrile to 0.5M. The resulting solution is treated with methanesulfonyl chloride (1.2 eq) and cesium carbonate (2 eq). The resulting mixture is stirred at reflux for 6 hours, then cooled and filtered and concentrated in vacuo. The desired azacrown ether **133** is formed as a minor product, isolated from the residue by flash chromatography on silica gel using ethyl acetate in hexanes.

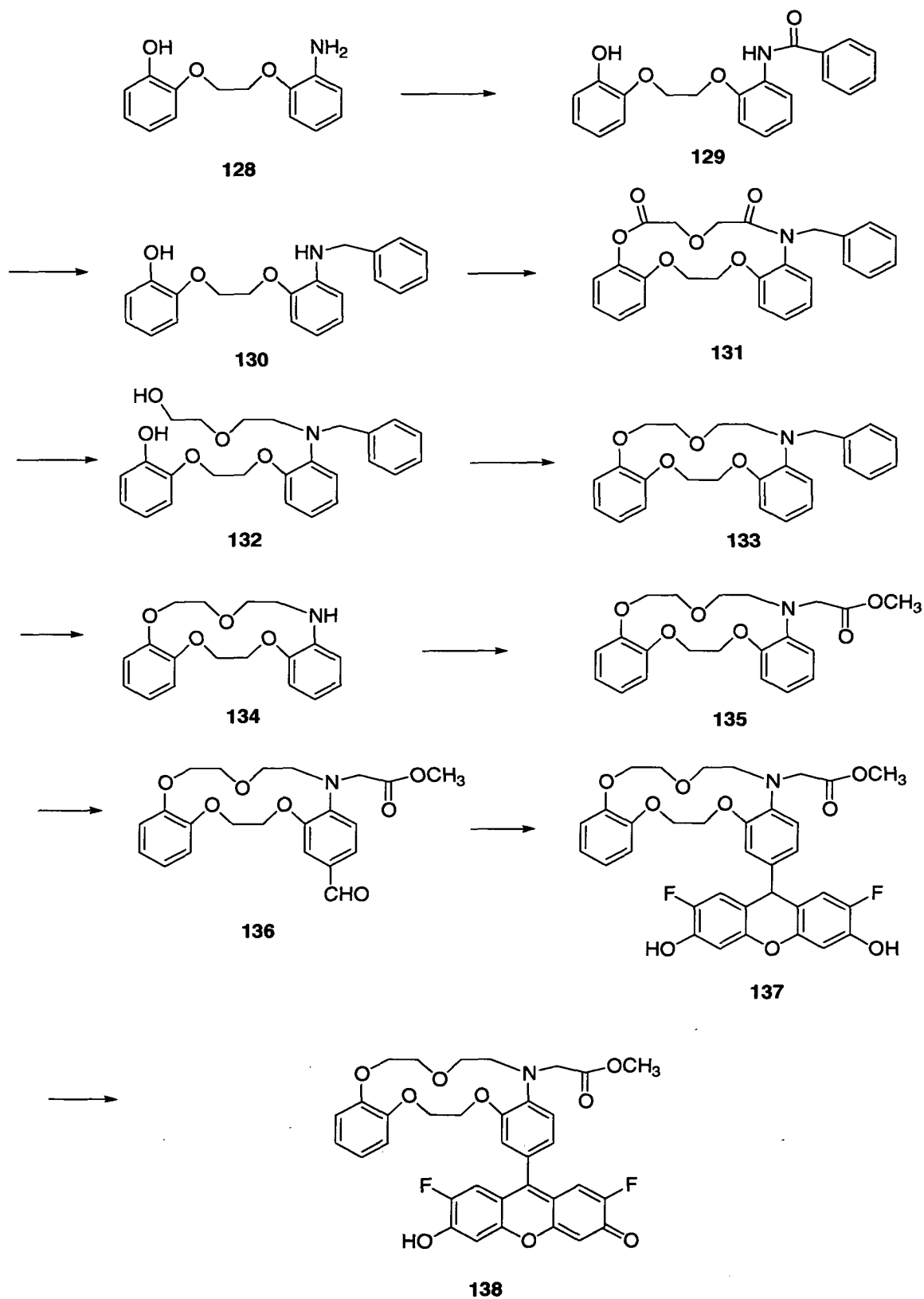
The crown **133** is stirred in ethyl acetate (0.5M) with 10 wt% Pd/C (cat.) under 30 psi hydrogen gas for 6 hours. After filtration, the reaction solution is concentrated in vacuo to give aniline **134**.

To a solution of **134** in DMF (0.5M) is added methyl bromoacetate (5 eq) and DIEA (3 eq). The resulting solution is stirred at 90°C overnight, then cooled and concentrated. The residue is partitioned between 0.1M HCl and ethyl acetate. The organic layer is washed with brine and dried over sodium sulfate, then concentrated in vacuo to give crown **135**, which is purified further if necessary by flash chromatography on silica gel using ethyl acetate in hexanes.

To a solution of Vilsmeier reagent made from 2 eq phosphorous oxychloride in DMF is added a solution of **135** in DMF. The resulting solution is stirred at room temperature overnight, then poured into aq. sodium bicarbonate. The resulting mixture is extracted with ethyl acetate. The extract is washed with water and brine, dried over sodium sulfate, and concentrated in vacuo to give aldehyde **136**, which is purified further if necessary by crystallization from methanol.

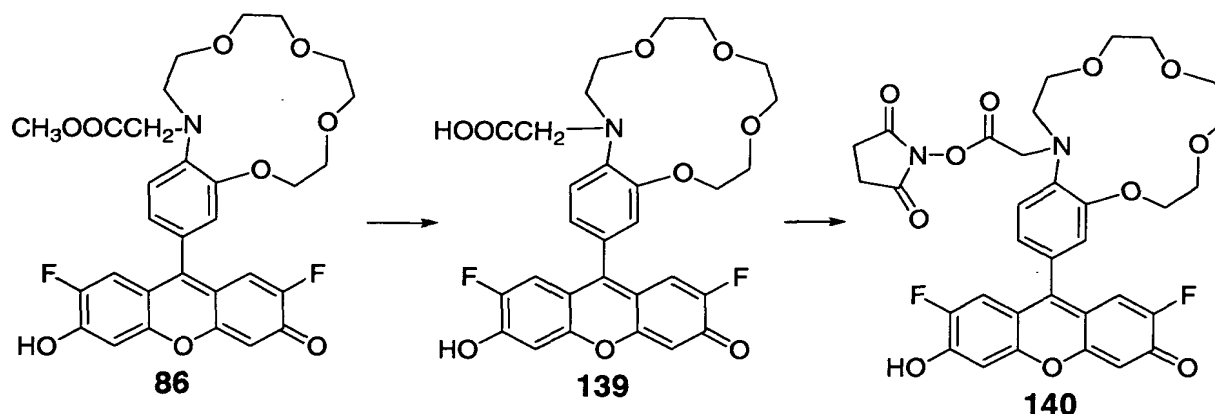
To a solution of **136** in methanesulfonic acid (0.5M) is added 4-fluororesorcinol (2 eq). The resulting solution is stirred at room temperature 15 minutes, then poured carefully into 3M aqueous NaOAc. The resulting precipitate is collected by suction filtration, rinsed with water, and dried in vacuo to give dihydro compound **137** as a pale brown powder.

A solution of **137** in methanol/chloroform (1:1, 0.5M) is treated with 2 eq of para-chloranil. The resulting mixture is stirred at room temperature overnight, then filtered and concentrated. The residue is purified by chromatography on silica gel using methanol in chloroform to give indicator **138** as a dark orange powder.



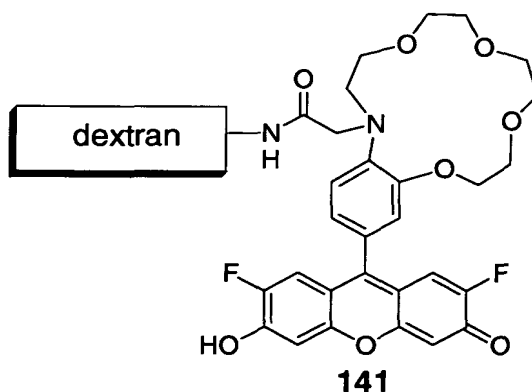
**Example 118: Synthesis of Compound 140**

A 0.5M solution of compound **86** in 1:1 methanol/dioxane is treated at room temperature with 5 equivalents of tetrabutylammonium hydroxide. After 2 hours the volatiles are removed in vacuo and the residue is dissolved in water. The pH is lowered to 2.0 by dropwise addition of aqueous HCl. The resulting precipitate is collected by filtration and dried in vacuo to give carboxylic acid **139**. A 0.5M solution of compound **139** in dry DMF is treated with 1.5 equivalents of O-(N-succinimidyl)-N,N,N',N'-tetramethyluronium tetrafluoroborate. After 2 hours the reaction solution is diluted 10x with diethyl ether. The resulting precipitate is collected by filtration to give reactive succinimidyl ester compound **140**.



#### Example 119: Synthesis of Compound **141**

A 0.1M solution of compound **140** in DMSO is added to a 0.1M solution of an equivalent mass of aminodextran (average MW 10000) in 0.3M sodium bicarbonate solution. The resulting solution is stirred at rt overnight, then slowly diluted with 10x methanol. The resulting precipitate is collected by centrifugation to give conjugated substance **141**, which can be purified further if necessary by gel filtration chromatography using water and Sephadex G-15.



All publications, patents and patent applications referred to within this document are incorporated by reference to the same extent as if each individual publication, patent or patent application was specifically and individually indicated to be incorporated by reference.

- 5 The reagents employed in the examples are commercially available or can be prepared using commercially available instrumentation, methods, or reagents known in the art. The foregoing examples illustrate various aspects of the invention and practice of the methods of the invention. The examples are not intended to provide an exhaustive description of the many different embodiments of the invention. Thus, although the foregoing invention has
- 10 been described in some detail by way of illustration and example for purposes of clarity of understanding, those of ordinary skill in the art will realize readily that many changes and modifications can be made thereto without departing from the spirit or scope of the appended claims.